



HAL
open science

Novel Ge–Ga–Te–CsBr Glass System with Ultrahigh Resolvability of Halide

Ci Cheng, Xunsi Wang, Tiefeng Xu, Qingde Zhu, Lihong Sun, Zhanghao Pan, Qiuhua Nie, Peiqing Zhang, Yuehao Wu, Shixun Dai, et al.

► **To cite this version:**

Ci Cheng, Xunsi Wang, Tiefeng Xu, Qingde Zhu, Lihong Sun, et al.. Novel Ge–Ga–Te–CsBr Glass System with Ultrahigh Resolvability of Halide. *Spectrochimica Acta Part A: Molecular and Biomolecular Spectroscopy* [1994-..], 2015, 150, pp.737-741. 10.1016/j.saa.2015.06.029 . hal-01165429

HAL Id: hal-01165429

<https://hal-univ-rennes1.archives-ouvertes.fr/hal-01165429>

Submitted on 15 Sep 2015

HAL is a multi-disciplinary open access archive for the deposit and dissemination of scientific research documents, whether they are published or not. The documents may come from teaching and research institutions in France or abroad, or from public or private research centers.

L'archive ouverte pluridisciplinaire **HAL**, est destinée au dépôt et à la diffusion de documents scientifiques de niveau recherche, publiés ou non, émanant des établissements d'enseignement et de recherche français ou étrangers, des laboratoires publics ou privés.

Novel Ge–Ga–Te–CsBr Glass System with Ultrahigh Resolvability of Halide

Ci Cheng¹, Xunsi Wang^{1*}, Tiefeng Xu¹, Qingde Zhu¹, Lihong Sun¹, Zhanghao Pan¹,
Qihua Nie¹, Peiqing Zhang¹, Yuehao Wu¹, Shixun Dai¹, Xiang Shen¹, Xianghua Zhang²

⁽¹⁾Laboratory of Infrared Material and Devices, The Research Institute of Advanced Technologies, Ningbo University, Ningbo, Zhejiang 315211, China)

⁽²⁾Laboratory of Glasses and Ceramics, UMR 6226 CNRS-University of Rennes 1, Rennes Cedex 135042, France)

Abstract: CO₂ molecule, one of the main molecules to create new life, should be probed accurately to detect the existence of life in exoplanets. The primary signature of CO₂ molecule is approximately 15 μm, and traditional S- and Se-based glass fibers are unsuitable. Thus, Te-based glass is the only ideal candidate glass for far-infrared detection. However, Te-based glass crystallizes easily, and its optical band gap is usually less than 0.6 eV. In this study, a new kind of Te-based chalcogenide glass system was discovered with relatively stable and large optical band gap. A traditional melt-quenching method was adopted to prepare a series of (Ge₁₅Ga₁₀Te₇₅)_{100-x} (CsBr)_x chalcogenide glass samples. The glass-forming ability and optical properties of these samples were analyzed by X-ray diffraction, differential scanning calorimetry, Fourier transform infrared spectroscopy, visible near-infrared spectroscopy, and Raman spectroscopy. Experiment results indicate that the glass-forming ability and thermal properties of glass samples were improved when CsBr was added in the host of Ge–Ga–Te glass. Ge–Ga–Te glass could remarkably dissolve CsBr content as much as 85 at.%, which is the highest halide content in all reports for Te-based chalcogenide glasses. Moreover, ΔT values of these glass samples were all above 100 °C. The glass sample (Ge₁₅Ga₁₀Te₇₅)₆₅ (CsBr)₃₅ with ΔT of 119 °C was the largest, which was 7 °C larger than that of Ge₁₅Ga₁₀Te₇₅ host glass. The infrared transmission spectra of these glasses show that the far-infrared cut-off wavelengths of (Ge₁₅Ga₁₀Te₇₅)_{100-x} (CsBr)_x chalcogenide glasses were all beyond 25 μm. In conclusion, (Ge₁₅Ga₁₀Te₇₅)_{100-x} (CsBr)_x chalcogenide glasses are potential materials for far-infrared optical application.

Keywords: Te-based chalcogenide glass; infrared transmission spectra; glass-forming ability; optical properties

1. Introduction

In recent years, chalcogenide glasses have attracted increasing attention in infrared technology fields, particularly in high nonlinearity, infrared fingerprints, and life detection [1–3]. Among these projects, the Darwin Mission and the Terrestrial Planet Finder conducted by the European Space Agency and the National Aeronautics and Space Administration, respectively, are the most impressive [3,4]. These projects aim to demonstrate whether life exists in exoplanet. H₂O, CO₂, and O₃ are the three main molecules to create new life. Detecting these molecules needs IR optical devices with excellent infrared transparency. The infrared cut-off wavelengths of S- and Se-based glass fibers are only up to 8 and 10 μm, respectively. However, the main signature of CO₂ molecule is approximately 15 μm [5]. Hence, current infrared transparencies of these glasses

* Corresponding author. Fax: +86 574 87600947.

E-mail address: xunsiwang@siom.ac.cn (X.Wang).

are not sufficiently wide to match with CO₂ molecular signature at 15 μm. Thus, a glass with wider infrared transparency is necessary. According to phonon vibration spectrum analysis [6], heavy atoms are needed for long cut-off wavelength. Te-based glasses have wider infrared transparency and longer cut-off wavelength than S- and Se-based glasses because the atom mass of Te-based glasses is larger than those of the other two glasses. In addition, the cut-off wavelength of Te-based glass in the long wavelength area is above 20 μm. Given the wide optical windows and low loss in 6–20 μm range, Te-based chalcogenide glasses are suitable material for far-infrared application, particularly in some gas detections, such as the greenhouse gas CO₂ with absorption peaks at 4.3 and 15 μm [7,8] and poisonous gas benzene with absorption peak at 14.8 μm [9]. However, these glasses possess many disadvantages. First, Te cannot form stable glass alone because of its strong metallic property. Second, traditional Te–X(F, Cl, Br, I) glasses have low transition temperature; most of these glasses cannot reach more than 100 °C. Its low optical band gap also limits its transmission in near IR (<2 μm), which is currently well developed. Moreover, the interstructure of traditional Te-based glasses possesses numerous inherent glass network defects, and these defects can result in high intrinsic loss. To solve these limitations, many scientists have exerted effort to develop suitable Te-based chalcogenide glasses. Adon et al. [10] reported that Ge–As–Te glasses have stable structure and good formability. However, As is poisonous to the environment and has a low band gap. Wilhelm et al. [3] demonstrated that Ge–Te–I glasses possess a wide infrared transparency window and its cut-off wavelength is more than 25 μm. However, this glass is prone to volatilization because I element is contained. Danto et al. [5] proposed that the GeTe₄ binary glass system doping Ga can open the GeTe₄ tetrahedral network structure and split the Te–Te chains; Ge₁₅Ga₁₀Te₇₅ is also superior among these glasses. Wang et al. discovered that doping halide to Te-based glass can decrease optical loss of glass [11,12]. The highest content of halide in the Te-based glasses is limited to 30 at.%.

In this study, an extensive investigation was developed to study the effect of alkali-halide on Ge–Ga–Te host glass. High content of CsBr was added to the Ge–Ga–Te glass system to enlarge the formation of Te-based glass. The optical and thermal properties of this glass were then analyzed by Fourier transform infrared spectroscopy (FTIR) and spectrophotometry. To date, no reports exist about the influence of CsBr on the structure and chemical or physical properties of Ge–Ga–Te–CsBr glasses.

2. Materials and Methods

(Ge₁₅Ga₁₀Te₇₅)_{100-x}(CsBr)_x glass samples were prepared by conventional melt-quenching methods. The specific glass compositions are listed in Table 1. High purities of 99.999% Ge, 99.999% Ga, 99.999% Te, and 99.999% CsBr were selected as raw materials. The weighed raw materials were mixed into a quartz tube, which was prewashed with deionized water. The tubes were then sealed with oxy-acetylene flame under a pressure of 1×10^{-3} Pa and heated in rocking furnaces with specific heating curve. Afterward, the quartz tubes were quenched in ice water and then annealed at 10 °C below T_g in a prepared furnace. The glass rods were removed and cut into discs. A polishing procedure on the discs is indispensable to test the properties of these glasses precisely.

The Archimedes' principle was used to measure the density of glass samples (with an accuracy of ± 0.001 g/cm³). XRD diffraction patterns of glass samples were tested through German Bruker D2 X-ray diffraction apparatus. Vis-near-infrared spectra of glass samples were acquired

using a Perkin-Elmer Lambda 950 spectrophotometer in the range of 400–2500 nm. The infrared transmission spectra were obtained with Nicolet 380 FTIR in a spectral range of 400–4000 cm^{-1} . Glass transition temperature (T_g) and onset crystallization temperature (T_x) were determined with differential scanning calorimetry (DSC) measurements at 50 and 350 $^{\circ}\text{C}$ with a heating rate of 10 $^{\circ}\text{C}/\text{min}$ via a TAQ2000 thermal analyzer. Raman spectra of glass samples were gathered in the range of 80 cm^{-1} –800 cm^{-1} by using a Renishaw Raman microscope with an Ar⁺ ion laser that operates at a wavelength of 488 nm. The spectral resolution was set to 1 cm^{-1} . All of these optical tests were performed at room temperature.

3 Results and Discussions

3.1 Physical properties

A series of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{100-x}(\text{CsBr})_x$ ($x = 0, 5, 10, 15, 20, 25, 35, 45, 55, 65, 75, 85$) glasses were prepared. These glasses are completely opaque and black in the visible region, and the physical parameters of these glasses are listed in Table 1. The densities of these glasses gradually decreased with increasing CsBr content because the glass density is generally determined by the elemental relative atomic mass. The relative molecular mass of CsBr was smaller than those of GeTe_4 and GaTe_3 . The recruitment and increase of CsBr resulted in reduced GeTe_4 and GaTe_3 contents. Consequently, the densities of these glasses gradually decreased, and the average molar volume increased. The average molar volume of the glass sample can be achieved using Formula (1).

$$V_m = \frac{\sum M_i}{\rho} \quad (1)$$

where M_i is the molar mass of glass sample ($M_i = A_i B_i$, A_i , molar concentration; B_i , molecular weight of glass composition; ρ , density of glass samples).

Table 1 Physical and thermal parameters of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{100-x}(\text{CsBr})_x$ glass samples

x (at.%)	ρ ($\text{g}\cdot\text{cm}^{-3}$)	V_m ($\text{m}^3\cdot\text{mol}^{-1}$)	T_g ($^{\circ}\text{C}$)	T_x ($^{\circ}\text{C}$)	ΔT ($^{\circ}\text{C}$)	Direct- E_{opt} / (eV)	Indirect- E_{opt} / (eV)	Cut-off λ (nm)
0	5.735	19.869	172	284	112	0.654	0.656	1789
5	5.667	20.982	171	285	114	0.659	0.659	1847
15	5.658	22.766	168	284	116	0.673	0.673	1816
25	5.555	24.971	173	286	113	0.702	0.699	1796
35	5.549	26.783	172	291	119	0.695	0.695	1718
45	5.459	29.039	169	283	114	0.677	0.676	1739
55	5.437	30.978	175	281	106	0.658	0.671	1744
65	5.412	32.951	175	280	105	0.671	0.669	1766
75	5.399	34.865	177	278	101	0.678	0.675	1830
85	5.331	37.168	179	279	100	0.639	0.617	1860

3.2 XRD analysis

XRD was conducted to verify the amorphous state of the samples. The XRD patterns of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{100-x}(\text{CsBr})_x$ glass samples are presented in Fig. 1. The peak shapes in these diffraction curves exhibited dispersion and width. The XRD results indicated that these glass samples possessed an amorphous state. $\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75}$ glass could remarkably dissolve as much as

85 at.% content of alkali-halide CsBr. Thus, its glass-forming ability is strong.

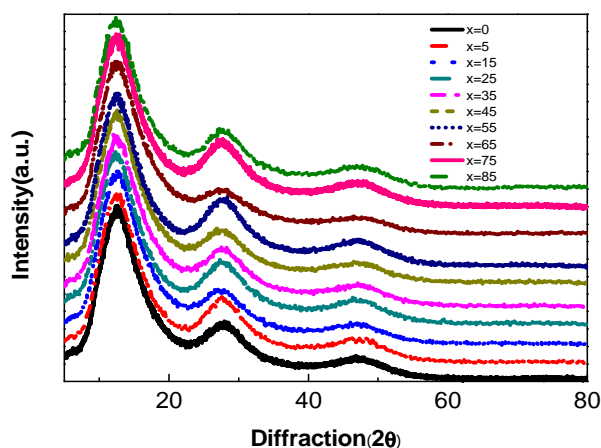


Fig. 1 XRD patterns of powdered glass samples

3.3 Thermal properties

Thermal properties of glass samples are an important aspect in evaluating a glass sample. The thermal characteristics of glass samples were tested with DSC instrument. DSC curves are all shown in Fig. 2. Two important parameters of glasses, namely, T_g and T_x , can be obtained with the curves. T_g is the glass transition temperature, and T_x is the glass onset crystallization temperature. Parameter ΔT (the difference between T_x and T_g) can be also obtained and used to estimate glass-forming ability and thermal stability. Stronger glass-forming ability and thermal stability indicates better anti-crystallization. The specific numerical temperatures T_g and T_x are also presented in Table 1. The lowest ΔT is 100 °C, and the highest ΔT is 119 °C, which corresponded to the $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{65}(\text{CsBr})_{35}$ glass sample. In contrast to the ΔT of $\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75}$ glass sample, the ΔT of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{65}(\text{CsBr})_{35}$ glass was increased by 7 °C. This finding proves that the proper content of alkali-halide CsBr added to the $\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75}$ glass could improve glass stability when the content was lower than 35 at.%. This result is attributed to the glass network structure, which was changed and improved by introducing halogen Br atoms. When the glass network was broken, the bromine atom could trap electrons of Te and form covalent bonds. Thus, a decreasing tendency of Te microcrystal formation occurred. Consequently, the stability of this glass against crystallization is ameliorated. However, increased CsBr content (>35 at.%) could destroy the glass network because of excess terminal halogen.

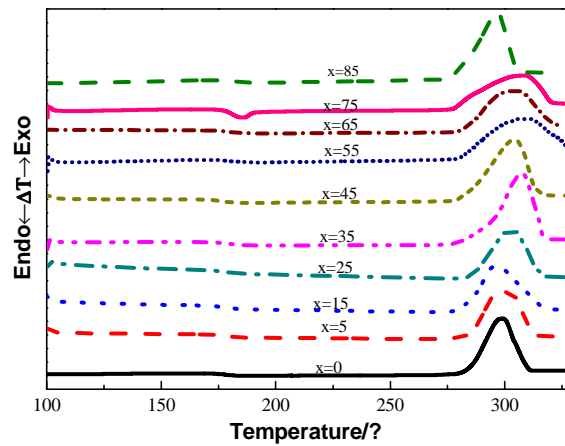


Fig. 2 DSC curves of Ge-Ga-Te-CsBr glass samples

3.4 Raman spectra analysis

The Raman spectra of the glasses are shown in Fig. 3, where four main vibration bands can be found. Three strong bands can be separately observed at 63, 126, and 159 cm^{-1} , and a very weak band was located at 217 cm^{-1} . The first and second bands could be attributed to the vibrations of Ge-Te bonds [13–15]. The peak at 159 cm^{-1} was ascribed to Te-Te bonds [11]. Moreover, the weak band at 217 cm^{-1} was caused by Ge-Ge bonds [16]. Some subtle changes occurred at the Te-Te peak, which was located at 159 cm^{-1} . The increase of CsBr content resulted in gradual withdrawal of vibration intensity of Te-Te bonds from $x = 5$ to $x = 35$. However, when x was above 35, the vibration intensity of Te-Te bonds gradually enhanced. This phenomenon occurred probably because the glass network structure was partly destroyed by CsBr. In general, with the high electronegativity of Br^- and low polarizability of Cs^+ doped into the Ge-Ga-Te glass, GaTe_3 triangle and GeTe_4 tetrahedra, which are the main structural units of glass network, were partly destroyed, particularly Te-Te chains. Bromine atom, as a network terminator, is considered to split the Te-Te chains. Thus, the Te-Te bonds reduced, and the terminal bonds Ge-Te-Br and Ga-Te-Br eventually formed. Therefore, the vibration band at 159 cm^{-1} gradually weakened. When CsBr content exceeded 45 at.%, many dissociative Te atoms or Te atoms, which were under full valence state, may appear in the glass network. These Te atoms eventually combined together because of their strong metallic property, and the peak intensity at 159 cm^{-1} gradually strengthened.

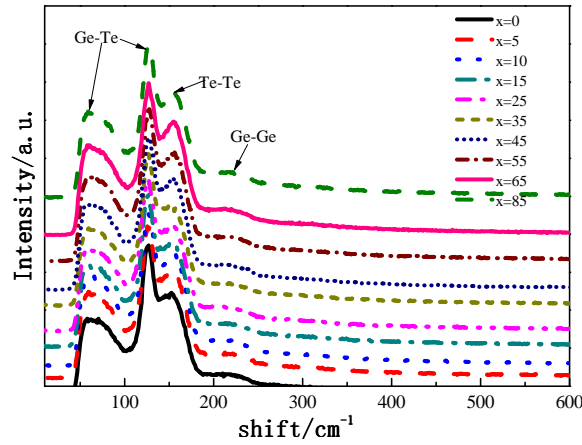


Fig. 3 Raman spectra of glass samples

3.5 Near-infrared absorption spectra and optical band gap

Near-infrared absorption spectra are shown in Fig. 4. When CsBr content increased from 5 at.% to 35 at.%, the absorption cut-off edge moved to short wavelength, which means that blue shift occurred. However, the absorption cut-off edge turned to long wavelength when x was above 35, which indicates that red shift occurred. Lastly, the shortest near-infrared cut-off wavelength was approximately 1718 nm. This phenomenon existed as a result of the structure of $\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75}$ glass, which was destroyed by the amount of CsBr content codoped. As shown in Fig. 3, the Te–Te bond intensity weakened when x was between 5 and 35. Thus, glass formality and stability were improved. This phenomenon may explain the blue shift of absorption cut-off edge at short wavelength. The red shift of absorption cut-off edge may be ascribed to the increase of the Te–Te bond intensity.

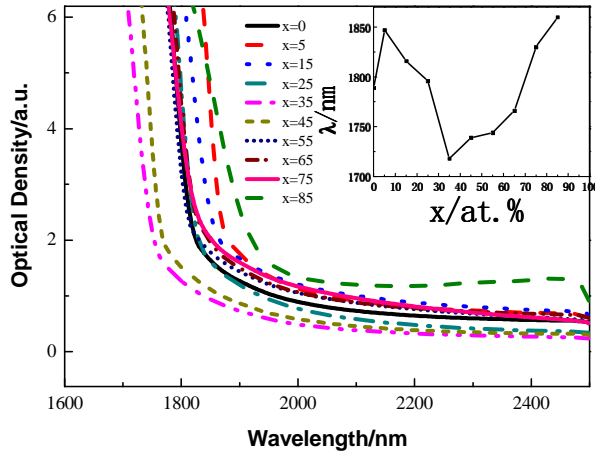


Fig. 4 Vis-IR absorption spectra of glass samples (relationships between cut-off edge wavelength and CsBr content are shown in the inset image)

The function relationship between absorption coefficient, $\alpha(\omega)$, and photon energy, $\hbar\omega$, of chalcogenide glass is given by Tauc equation [17]:

$$\alpha(\omega) \cdot \hbar\omega = B(\hbar\omega - E_{opt})^m \quad (2)$$

In the preceding equation, α is the absorption coefficient, which is determined as $\alpha = 2.303 A/d$ (A

is the optical density of the testing glass sample, and d is the thickness of the glass sample), E_{opt} is the optical band gap, h is Plank constant, ω is the incident light angular frequency, and m is a parameter that can determine the transition type of absorption edge. For amorphous glass materials, the direct and indirect allowed transitions correspond to $m = 1/2$ and $m = 2$, respectively. B is a constant about local state in the band gap. It can be calculated with the following equation:

$$B = \frac{(4\pi c)\sigma_0}{n_0\Delta E} \quad (3)$$

In Eq. (3), c is the light speed in vacuum, σ_0 is the electrical conductivity of absolute zero, n_0 is the static refractive index, and ΔE is the local state tail width. Figs. 5 and 6 correspond to the direct and indirect band gaps, respectively. The largest value of direct band gap is 0.702 eV, whereas the smallest value of indirect band gap is 0.617 eV. The band gap values of all the glass samples are listed in Table 1.

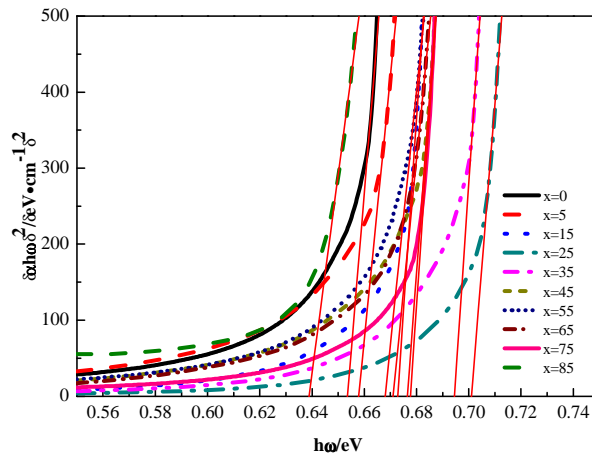


Fig. 5 Relationship between $(\alpha \cdot \hbar\omega)^2$ and $\hbar\omega$ for glass samples

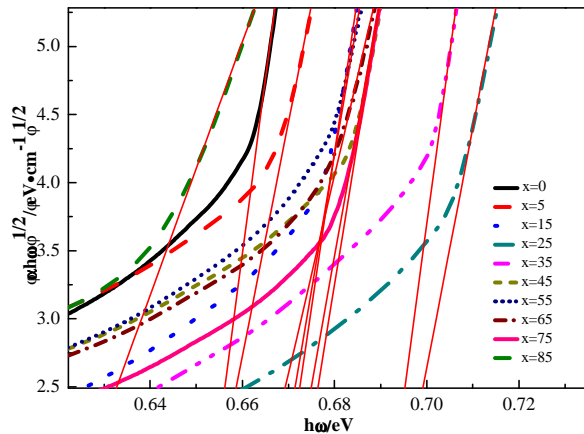


Fig. 6 Relationship between $(\alpha \cdot \hbar\omega)^{1/2}$ and $\hbar\omega$ for glass samples

3.6 Infrared transmission spectra analysis

The infrared transmitting spectra of the glass samples are shown in Fig. 7. The highest transmission was up to 53% (close to the theoretical limitation), without purification for the $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{80}(\text{CsBr})_{20}$ glass sample. All glass samples had a wide infrared transmitting window,

and the infrared cut-off wavelength could reach up to 25 μm . However, when the wavelength was more than 20 μm , the transmittance of the glass samples sharply reduced. This result can be ascribed to the multiphonon absorption produced by the Ge–Te bond vibration [5]. Some other absorption peaks can also be observed in Fig. 7. The peak at 9.9 μm resulted from Si–O covalent bond vibration. The peak at 15–20 μm can be ascribed to the presence of some oxygen contamination, such as Ge–O or Ga–O bond vibrations [18,19]. The intensity of this peak depends on the purity of raw materials. The $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{80}(\text{CsBr})_{20}$ glass sample that possessed highest transmittance was chosen for the next purification experiment to eliminate glass absorption peaks. The first method was distilling raw Te, and the purified Te was then transferred to another quartz ampoule that contained Ge, Ga, and CsBr raw materials. Another method is codoping 500 ppm Mg into a raw Te-containing ampoule. Te was then distilled with the same method mentioned previously. The succeeding steps were the same as those in the traditional melt-quenching methods. The transmission spectra of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{80}(\text{CsBr})_{20}$ glasses, which were obtained by the preceding two methods, are shown in Fig. 8. The intensity of absorption peak from 15 μm to 20 μm decreased when the first method was adopted to purify the glass. However, this method cannot thoroughly remove the absorption peaks. For the second purification method, the absorption peaks could be completely eliminated, and the transmittance of this glass can reach up to 54%.

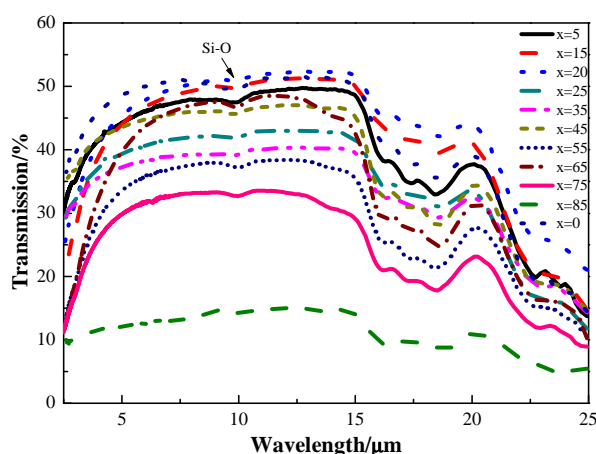


Fig. 7 Infrared transmission spectra of glass samples $[(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{100-x}(\text{CsBr})_x]$

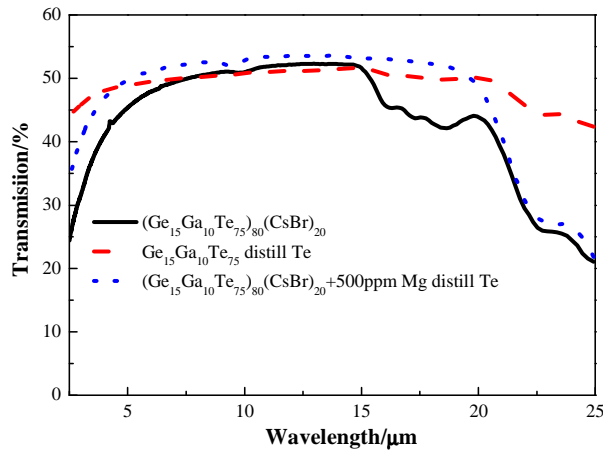


Fig. 8 Infrared spectra of $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{80}(\text{CsBr})_{20}$ glasses

4. Conclusion

A novel series of Ge–Ga–Te–CsBr glasses were prepared via a traditional vacuum melting and quenching method. The thermal, physical, and optical properties of these glasses were investigated in detail. All glass samples exhibited good thermal and physical properties. The $(\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75})_{65}(\text{CsBr})_{35}$ glass was superior to other glass samples. The ΔT of this glass was 119 °C and was higher than that of $\text{Ge}_{15}\text{Ga}_{10}\text{Te}_{75}$ glass, which is only 112 °C. The solubility of alkali-halide CsBr in Ge–Ga–Te glass can reach as much as 85 at.%, which can be distinguished as a new kind of halide-based glass. The infrared transmittance of all the glass samples showed wide optical windows, which ranged from 1.7 μm to 25 μm . These excellent properties make Ge–Ga–Te–CsBr glasses potential materials for ultrawide spectrum infrared optical applications.

Acknowledgments

This work was financially supported by the Natural Science Foundation of China (Grant Nos. 61435009, 61177087, and 61377099), National Program on Key Basic Research Project (973 Program) (Grant No. 2012CB722703), International Science & Technology Cooperation Program of China (Grant No. 2011DFA12040), Scientific Research Fund of Zhejiang Provincial Education Department (R1101263), Natural Science Foundation of Ningbo (Grant No. 2013A610118), Teaching and Research Award Program for Outstanding Young Teachers in Higher Education Institutions of MOE, P.R.C. Ningbo Optoelectronic Materials and Devices Creative Team (2009B21007), and Scientific Research Foundation of Graduate School of Ningbo University. This work was also sponsored by K. C. Wong Magna Fund of Ningbo University and the Outstanding (Postgraduate) Dissertation Growth Foundation of Ningbo University (Grant No. PY2014014).

References

- [1] C. Quémard, F. Smektala, V. Couderc, J. Phys. Chem. Solids 62 (2001) 1435-1440.
- [2] B. Bureau, X.H. Zhang, F. Smektala, J. Non-Cryst. Solids 345 (2004) 276-283.
- [3] A.A. Wilhelm, C. Boussard - Pledel, Q. Coulombier, Adv. Mater. 19 (2007) 3796-3800.
- [4] X.H. Zhang, G. Fonteneau, J. Lucas J. Non-Cryst. Solids 104 (1988) 38-44.
- [5] S. Danto, P. Houizot, C. Boussard - Pledel, Adv. Funct. Mater. 16 (2006) 1847-1852.
- [6] X.B. Yu, X. Zhao, C.H. Li, Journal of Dalian Polytechnic University 27 (2008) 155-157.
- [7] G.X. Wang, Q.H. Nie, X.S. Wang, J. Appl. Phys. 110 (2011) 043536.
- [8] S. Maurugeon, B. Bureau, C. Boussard-Plédel, J. Non-Cryst. Solids 355 (2009) 2074-2078.
- [9] W.D. Chen, F. Cazier, F. Tittel, Appl. Opt. 39 (2000) 6238-6242.
- [10] L. Aldon, P. Lippens, J. Olivier-Fourcade, Chalcogenide Letters 7 (2010) 187-196.
- [11] X.S. Wang, Q.H. Nie, G.X. Wang, Spectrochimica Acta Part A 86 (2012) 586-589.
- [12] H.J. Xu, X.S. Wang, Q.H. Nie, J. Non-Cryst. Solids 383 (2014) 212-215.
- [13] A.V. Kolobov, P. Fons, J. Tominaga, J. Phys.: Condens. Matter 16 (2004) S5103.
- [14] S. Sen, E.L. Gjersing, B.G. Aitken J. Non-Cryst. Solids 356 (2010) 2083-2088.
- [15] K.S. Andrikopoulos, S.N. Yannopoulos, A.V. Kolobov, J. Phys. Chem. Solids 68 (2007) 1074-1078.
- [16] T. Fukunaga, Y. Tanaka, K. Murase Solid State Commun. 42 (1982) 513-516.
- [17] J. Tauc, *Amorphous and liquid semiconductors*. (Plenum Publishing Corporation, 1974).
- [18] Y.J. He, X.S. Wang, Q.H. Nie, Infrared Physics & Technology 60 (2013) 129-133.
- [19] S.Q. Zhang, X.H. Zhang, M. Barillot, Opt. Mater. 32 (2010) 1055-1059.