Synthesis and structural properties of Aza[n]helicene platinum complexes: control of cis and trans stereochemistry

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Abstract

The synthesis and the structural characterization of azahelicene platinum complexes obtained from \textit{cis}-PtCl\textsubscript{2}(NCEt)(PPh\textsubscript{3}) and from ligands that differ both in terms of the position of the N atom and the number of fused rings, is reported. These square planar complexes of general formula PtCl\textsubscript{2}(nHm)(PPh\textsubscript{3}) (n = 4,5, m = 5,6) display mainly a \textit{cis} configuration. However, by X-ray crystallographic analysis we show that for both PtCl\textsubscript{2}(4H6)(PPh\textsubscript{3}) and PtCl\textsubscript{2}(5H6)(PPh\textsubscript{3}) there is a chirality control of the \textit{cis}-trans stereochemistry. Indeed, while starting from a racemic mixture of aza[6]helicene, Pt complexes with a \textit{cis} configuration are invariably obtained, the more thermodynamically stable \textit{trans} isomers are formed when using enantiopure ligands. We further corroborated these results by NMR analysis in solution.

Introduction

Helicenes are a class of non-planar, polycyclic aromatic compounds whose helical backbone, formed by a variable number of ortho-fused benzene or other aromatic rings, renders them intrinsically chiral despite the absence of any chiral center.\textsuperscript{1} The extended $\pi$-conjugation of these systems explains some of their most important physical-chemical properties, including their high chiroptical properties and redox activity.\textsuperscript{1a,b,h,i} These systems generally exhibit larger intersystem crossing rates and larger magnetic dipole moments than their planar analogs. Owing to the long lifetime of their triplet state and to their tendency to form $\pi$-$\pi$ stacking, these compounds may be particularly useful as building blocks in materials for optoelectronic applications.\textsuperscript{1b,h,i}
particular, helicene/transition metal derivatives have recently been described in the literature as promising candidates in a variety of light emitting and sensing devices. The photoluminescence and chiroptical properties of helicenes combined with the visible light emission of transition metals are in fact very attractive features of their complexes due to the strong UV absorbance of the ligand and the possibility of energy transfer to the metal ion. Moreover, the transmission of chirality to transition metal complexes may allow the design of chiral luminescent materials, enantioselective sensors, chiroptical switches, and magnetochiral compounds. Helicene derivatives have also been extensively tested for their chiral selectivity in DNA binding. Moreover, possible roles of these compounds in the context of different biomedical and biotechnological applications such as drug development and biosensing have been identified. Another important aspect is the new reactivity features that may originate from the different solubility properties of racemates and pure enantiomers, as recently demonstrated by some of us. Therefore, there is considerable interest in the structural diversity of helicene-based coordination complexes and in the control of their supramolecular self-assembly properties via crystal engineering approaches.

Herein, we report the synthesis of aza[n]helicene/platinum complexes of general formula $\text{Cl}_2\text{Pt}(\text{nHm})(\text{PPh}_3)$ ($\text{nHm}: n$-aza[m]helicene, $n = 4, 5$, $m = 5, 6$), displaying either a cis or a trans configuration (Scheme 1). Five of these complexes were characterized by X-ray crystallography. Interestingly, they all displayed different packing arrangements. The tendency of the helicene ligand to form $\pi-\pi$ stacking and the ability of the platinum complexes to form hydrogen bonds involving the chlorine atoms allow for the formation of intramolecular and intermolecular stabilizing interactions. Finally, after our first serendipitous discovery that starting from either racemic mixtures or enantiopure forms of the 4-aza[6]helicene ligand ($\text{4H6}$) can affect the
cis/trans stereochemistry of the resulting Pt complex, we set out to investigate whether other chiral ligands such as 5H6 would display the same feature.

Scheme 1. Molecular structures and atom numbering for ligands 4H5, 5H5, 4H6 and 5H6 and for their corresponding Pt complexes 1 – 4 (a and b stands for cis and trans stereochemistry).

cis-PtCl₂(NCEt)(PPh₃) 5a, i) Propionitrile, 1-4 days, RT. ii) Toluene, 24 hrs, reflux.
Results and Discussion

_Cis_ Pt-complexes of aza[5]helicenes

Inspired by the interesting properties revealed by the complex between Pt and 4-aza[6]helicene,4 we carried out the synthesis and the structural characterization of a number of Pt complexes of aza[5]helicene and aza[6]helicene ligands with N atom in different positions within the fused ring system (positions 4 and 5). At first, the relatively small 4- and 5-aza[5]helicenes were tested as ligands.6 Our first attempt to synthesize these aza-helicene/platinum complexes starting from PtCl2(NCR)2, which is generally recognized as a good starting material,7 did not result in the desired products but in a complex mixture. However, an efficient strategy to synthesize aza-helicene platinum complexes 1 and 2 was later developed, consisting in the substitution of the nitrile group in the _cis_-PtCl2(NCEt)(PPh3) complex (5a) by the 4- or 5-aza[5]helicene ligands in propionitrile at room temperature (RT) and under continuous stirring (Scheme 2).5 Complexes 1 and 2 were obtained from 4H5 and 5H5 respectively in moderate to good yields (Scheme 1). The 31P NMR of 1 and 2 showed one signal at 3.1 ppm, with 195Pt-31P coupling constants of 3660-3670 Hz (see Supporting Information, SI). These complexes were also characterized by 195Pt NMR which showed doublets at -3625 and -3522 ppm for 1 and 2, respectively, with the same J(195Pt-31P) coupling constants. Moreover, the aza[5]helicenes 1H NMR chemical shifts undergo several modifications upon coordination to the PtCl2(PPh3) moiety. For example, in complex 1, the H5 proton of the 4H5 moiety undergoes a 1.4 ppm down-field shift upon coordination. Similarly in complex 2, the H4 proton of the 5H5 moiety undergoes a 1.4 ppm down-field shift. In addition, the hydrogen H6 proton in the complex 2 shows a coupling constant of 4 Hz that is absent in the free ligand, and which is assigned as a 4JH-P.
coupling. Finally, the ESI mass spectrometry of both complexes 1 and 2 afforded a peak at m/z 813.2 corresponding to the cationized [PtCl(nH5)(PPh3)CH3CN]⁺ (see SI).

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\text{cis-PtCl}_2(PPh_3)(\text{EtCN}) + n\text{H}_5 \xrightarrow{\text{r.t., 1-4 days}} \text{cis-PtCl}_2(PPh_3)(nH5) + \text{EtCN}
\]

\[\text{cis-5a} \quad \text{n-aza[5]helicene} \quad 1,2 \quad \text{EtCN} \]

**Scheme 2.** Synthesis of cis platinum complexes 1 and 2 from 4H5 and 5H5 ligands.

Medium size single crystals of complexes 1 and 2 suitable for X-ray structural studies were obtained by slow evaporation at RT. Complex 1 crystallizes in the triclinic space group P-1, with one cis-PtCl2(4H5)(PPh3) molecule and one solvent molecule (dichloromethane) in the asymmetric unit. Owing to the presence of an inversion center, two molecules with opposite M and P handedness for the 4-aza[5]helicene ligand are present in the unit cell. In each complex 1, the platinum center is coordinated by two chlorine ligands, one 5-aza[6]helicene ligand and one PPh3 (Figures 1 and 2). The molecule shows the typical square planar geometry of platinum(II) complexes with a slight distortion from the ideal 90° arrangement (the N4-Pt-P and N4-Pt-Cl1 angles are 94.2(3)° and 87.1(3)°, respectively) presumably due to steric hindrance of the ligand. However, the Pt atom and its four coordinated atoms are quite coplanar (out of plane distortion ca. 2°). Furthermore, the distance between the platinum and the chlorine atom trans to the phosphine (Pt-Cl1: 2.343(3) Å) is longer than the distance between the platinum and the chlorine atom trans to the helicene (Pt-Cl2: 2.291(3) Å). These values are in full agreement with literature data from similar complexes, and with the stronger effect of phosphine compared to pyridine-type ligands. In the complex, the helicity of the helicene ligand is 46.46° and is therefore comparable to the helicity of the free 4-aza[5]helicene ligand (51.13°, see ref. 6), suggesting that no ligand distortion occurs upon metal coordination.
Interestingly, the trapped dichloromethane molecules, sitting in the inversion center, form weak C-H···Cl interactions with the first ring of the helicene and with one phenyl ring of the PPh₃ group. Such interactions extend along the x crystallographic axis forming a 1D chain (Figure 2a). In addition, intermolecular C-H···π (Figure 2b) and inter/intramolecular π−π interactions (Figure 2c) can be observed between the phenyl rings of the PPh₃ ligand and either the central or the terminal aromatic rings of the poly-fused helicene backbone, within two heterochiral complexes formed by M- and P-4-aza[5]helicene ligands. Note however that in solution the aza[5]helicenes are known to be configurationally unstable.⁶

![Figure 1. X-ray crystallographic structures of cis Pt-complexes 1, 2, 3a, 4a and trans Pt-complex M-4b.](image)

Figure 1. X-ray crystallographic structures of cis Pt-complexes 1, 2, 3a, 4a and trans Pt-complex M-4b.
Figure 2. Crystal structure of 1 showing a) the arrangement along the x axis, with H-bonding interactions involving the CH₂Cl₂ solvent molecules (disordered, two positions found), b) the intermolecular CH-π interactions and c) the inter/intramolecular π-π interactions between PPh₃ and the pyridyl cycle, and between the two central rings of two 4H5 in the heterochiral assembly of cis-PtCl₂(4H5)(PPh₃) (1).

Complex 2 crystallizes in the monoclinic space group C2/c. Owing to the centrosymmetric nature of the space group, molecules with opposite M and P handedness are present in the unit cell (Figure 3). This complex is characterized by the presence of a network of interactions whereby each molecule (A) is linked to three others (B, C, D) by two types of hydrogen bonds. These can be classified as horizontal links (see, for instance, Cl1(A)...H8(B) and Cl3(C)...H8(A) of Figure 3a) and vertical links (see Cl1(A)...H6(D) and Cl2(D)...H6(A) in
Figure 3a). The latter seem to suggest a cooperative effect between molecules A and D. Moreover, π-π interactions are sketched in Figure 3b. Intramolecular π-π stacking can be observed between the aromatic ring of the phosphine and the azahelicene ligands on the same metal complex, while intermolecular π-π interactions occur between helicene ligands from different metal complexes, with the interacting molecules showing the same chirality (either M or P).

Overall, the cis arrangement of the chlorine ligands in Pt complexes 1 and 2 demonstrates that the stereochemistry is maintained during the replacement of propionitrile by the 4- or 5-aza[5]helicene ligands 4H5 and 5H5. Indeed, Belli Dell’Amico et al. observed that the cis isomer 5a is more stable than the trans isomer (trans-5b, vide infra) in a propionitrile solution at room temperature.9

**Cis and trans Pt-complexes of aza[6]helicenes**

The synthesis of 4- and 5-aza-[6]helicene Pt complexes 3a,b and 4a,b from 4H610 and 5H611 was accomplished by a different strategy. It was indeed observed that the nitrile substitution reaction at RT was very slow, whereas the reaction would occur overnight upon heating. Indeed, the use of non-nitrile based solvents with a sufficiently high boiling point, e.g., toluene, promoted complex formation. The reaction was monitored by 31P-NMR, which showed the complete disappearance of the starting reactants and the onset of new signals at different chemical shifts and with different J couplings. The detailed characterization of the reaction product was obtained by 1H-NMR and ESI mass spectrometry (see ref. 4 and SI).
Figure 3. Intermolecular interactions in cis-PtCl$_2$(5H5)(PPh$_3$)$_2$: hydrogen bonds (a), π-π interactions between aromatic part of the complexes (b).

Previous studies$^9$ demonstrated that in the cis-5a complex the nitrile group can easily be substituted by another ligand in refluxing toluene. The substitution is achieved via isomerization to the trans complex 5c– which is more reactive than the cis one likely because of the trans-
effect of the phosphine ligand – and the possible formation of the chlorido-bridged trans-[PtCl(μ-Cl)(PPh3)]2 dimer trans-5c (Scheme 3). The latter can then be opened by many ligands in a relatively straightforward way. Moreover, the presence of the phosphorus atom in the triphenylphosphine ligand allows the use of 31P-NMR as a fast and unambiguous structural characterization method. Indeed, complex formation could be followed over time by monitoring the variation of the 31P chemical shifts and heteronuclear J couplings.

Scheme 3. Generic reaction pathway for 4- and 5-aza[6]helicene-platinum complexes with either cis (3a, 4a) or trans (3b, 4b) stereochemistry. Equilibrium process of platinum precursor as a mixture of cis-5a, trans-5b and μ-chlorido bridged trans-5c.
We have previously reported that the reaction of racemic 4-aza[6]helicene 4H6 with cis-PtCl₂(NCEt)PPh₃ 5 in refluxing toluene for one night resulted in the precipitation the cis-isomeric complex 3a while the reaction of enantiopure P- and M-4H6 under the same conditions yielded the enantiopure trans-isomeric complex P- and M-3b. This surprising result was identified as a dynamic process, namely a crystallization-induced diastereoselective transformation, due to the different solubilities of racemic and enantiopure 4-aza[6]helicene Pt complexes that displace the cis-trans equilibrium of Pt-precursors 5a-c in refluxing toluene (see Scheme 3). As already described in ref. 4, complex 3a crystallizes in the triclinic P-1 space group with the presence of M and P 4-aza[6]helicenes. The geometry around the platinum atom is square planar (Figure 1). The central metal is coordinated by two chlorine ligands in a cis mutual position, one 4-aza[6]helicene ligand and one PPh₃ group. Weak intramolecular π-π interactions were found to occur between one phenyl ring of PPh₃ and the pyridyl ring (centroid-centroid distance 3.852 Å) and a set of hydrogen bonds involving the Cl ligands through the whole crystalline network.

Following the uncommon behavior of 4H6 towards Pt complexation, we set out to investigate whether the same results could be observed with similar ligands such as 5H6. Indeed, using the same procedure with racemic 5H6, i.e., reacting it with cis-5a in refluxing toluene overnight, resulted in the precipitation of a yellow solid that was identified as the cis-Pt complex 4a (vide infra). In this new complex, differences in the 1H chemical shifts between the non coordinated and the coordinated 5H6 ligand could be observed. For example, the H2 and H15 protons appeared as ddd signals at 6.62 and 6.73 ppm in complex 4a while they resonate at the same chemical shift (6.8 ppm) in the free 5H6 ligand. The H6 proton is greatly influenced by the metal coordination and appears as an up-field shifted singlet at 9.16 ppm with a ¹⁹⁵Pt-H coupling
constant of 45 Hz. The $^{31}$P NMR spectrum of 4a shows a single signal at 6.9 ppm, with a $^{195}$Pt-P coupling constant of 3850 Hz, very similar to the one observed for complex 3a (see SI).

On the contrary, when using enantiopure P- or M-5H6 and under the same reaction conditions, the enantiopure P- or M-trans-Pt complexes 4b were obtained in reasonably good yields. These new trans complexes were first characterized by $^{31}$P NMR which displayed signals that were different from complex 4a (2.8 ppm, $^1J_{P,Pt} = 3645$ Hz, see SI). Similarly, the $^1$H-NMR spectrum displayed signals at different chemical shifts than 4a and was finally identified as the trans-isomeric complex by X-ray crystallography (vide infra). Clearly, these results show that 5H6 displays the same behavior as 4H6, the reactivity being triggered by the enantiopurity of the starting helicene ligand (racemic vs. enantiopure) and corresponding to a crystallization-induced diastereoselective transformation. To our knowledge, this is a very uncommon behavior in helicenes coordination chemistry.

Note however that the cis Pt complexes 3a and 4a display different behaviors in CH$_2$Cl$_2$ solutions. It was in fact observed by $^1$H NMR spectroscopy and electronic circular dichroism (vide infra) that rac-4a slowly isomerizes to the more thermodynamically stable complex rac-4b after several hours, while rac-3a remains unchanged. For this reason, the isolation of enantiopure samples of the cis complex 4a by HPLC separation over a chiral stationary phase is unattainable, as well as $^{195}$Pt and $^{13}$C NMR spectra of reasonably good quality. At this stage it is interesting to note that 3a appears more stable than 4a since the latter isomerizes to 4b while complex 3a remains unchanged in solution. Theoretical calculations have recently emphasized the role of $\pi-\pi$ interactions between the PPh$_3$ ligand and the 4aza[6]helicene in stabilizing some conformations of 3a. It is therefore most probable that the $\pi-\pi$ interactions taking place in 4a are not sufficient for stabilizing the cis configuration.
Despite its low solution stability and easy transformation to trans complex 4b, single crystals of rac-4a were obtained and characterized by X-ray crystallography. Complex 4a crystallizes in the triclinic space group P-1. Its X-ray crystallographic structure, depicted in Figures 1 and 4, reveals a square planar geometry around the platinum atom which is coordinated by two chlorine ligands in a cis mutual position, one 5-aza[6]helicene ligand and one PPh3. Owing to steric hindrance, a small distortion from the ideal 90° geometry is observed (the N5-Pt-P and N5-Pt-Cl1 angles are 94.27(17)° and 86.46(7)°, respectively). As for complex 3a, the distance between the platinum atom and the chlorine atom trans to the phosphine ligand (Pt-Cl1: 2.349(2) Å) is longer than the distance between the platinum atom and the chlorine atom trans to the helicene ligand (Pt-Cl2: 2.2933(16) Å). Interestingly, weak intramolecular π−π interactions take place between one phenyl of the PPh3 ligand and the pyridyl ring (centroid-centroid distance 3.774 Å). As discussed for 3a, owing to the steric hindrance of the helicene moiety, the PPh3 is stacked on one side of the pyridyl ring, thus generating planar chirality, with the pyPtCl2 defining the chiral plane.4,14 Indeed, C6-N-Pt-P and C4a-N-Pt-Cl1 torsion angles of -79.63 and -87.38° (pM-chirality) are measured in the cis-4a molecule having the M-4-aza[6]helicene ligand, suggesting that the M-helicity induces a fixed pM-chiral planar sense.4 The crystal packing is arranged to form parallel layers on the yz plane, as shown in Figure 4. The individual layers are connected by hydrogen bonds (distances 2.8-2.9 Å) of the type Cl2···H11 (helicene) and Cl1···H (PPh3) shown in Figure 4 in horizontal and vertical projection, respectively. Heterochiral assemblies are also found thanks to π−π interactions between azahelicenes of opposite handedness (Figure 5a). Finally, the crystal packing of 4a reveals a set of several different intermolecular CH···Cl hydrogen bonds that contribute to lattice stabilization (Figure 5b).
**Figure 4.** Arrangement of *cis* complex 4a in the *yz* plane. Hydrogen bonds between Cl and H atoms are emphasized in red.

**Figure 5.** a) π-π interaction between helicenic rings of two adjacent heterochiral *P*- and *M*-4a. b) Selected Cl···H interactions between two heterochiral *P*- and *M*-4a.
Single crystals of enantiopure $M$-$4b$ could be grown and characterized by X-ray crystallography. Complex $M$-$4b$ crystallizes in the orthorhombic chiral space group $P2_12_12_1$, with one molecule in the asymmetric unit and four molecules in the unit cell. While the Pt center shows an almost undistorted square planar geometry (Figures 1 and 6) (angles N-Pt-P: 170.70(11)$^\circ$, Cl1-Pt-Cl2: 174.67(5)$^\circ$, N-Pt-Cl1: 86.67(11)$^\circ$, N-Pt-Cl2: 90.48(11)$^\circ$), the two chlorine atoms are now in a mutual trans configuration. As a result, the two Pt-Cl bond distances are similar (Pt-Cl1: 2.3121(11) Å, Pt-Cl2: 2.2823(11) Å). A helicity angle$^\theta$ of 53.66$^\circ$ was measured in the solid state, showing that no ligand distortion occurs upon complexation. Several H…Cl hydrogen bonds are found between neighboring molecules (Cl2…H4: 2.707 Å, Cl1…H10: 2.932 Å, Cl2…Hm(PPh3): 2.968 Å, Cl2…Hp(PPh3): 3.163 Å) (see Figures 6a,b).

The electronic circular dichroism (ECD) spectra of the enantiopure samples of $M$- and $P$-$4b$ were also measured (Figure 7). The mirror-image ECD spectra display the characteristic bands of helicene derivatives, and are very similar to the ECD spectra of the $5H6$ ligand,$^{10}$ showing that in this case the coordination to PtCl$_2$(PPh$_3$) moiety has no influence on the chiroptical properties of the complexes.
**Figure 6.** Arrangement of *trans* complex 4b a) in the xz plane and b) in the yz plane. Hydrogen bonds between Cl and H atoms are emphasized in red.
Figure 7. Mirror-image ECD spectra of \( P \) (red) and \( M \) (blue) \textit{trans} Pt complexes 4b (CH\(_2\)Cl\(_2\), 5 \( 10^{-5} \) M). Insert: Magnified ECD bands in the 400-460 nm region.

Conclusions

We have reported the synthesis and the structural characterization of a number of novel \textit{cis} and \textit{trans} azahelicene/metal complexes using 5- and 6-membered helicenes with the N atom in position 4 and 5. The very different packing arrangements shown by these complexes is modulated by the different position of the nitrogen atom in the helicene moiety. While square planar (SP-4) Pt(II) complexes of general formula LL'PtX\(_2\) displaying the usual \textit{cis}/\textit{trans} isomerism is well-known in coordination chemistry\textsuperscript{15}, the effect of either the racemic or the enantiopure forms of aza[6]helicenes on such stereochemistry is an uncommon phenomenon\textsuperscript{4}. Such isomerism can have important practical implications in biology as with Pt(NH\(_3\))\(_2\)Cl\(_2\) (\textit{cis}-platin) which is an efficient antitumor drug, while its \textit{trans} isomer is ineffective\textsuperscript{16a,b}. Therefore, the control of the stereochemistry of SP-4 platinum complexes may provide an efficient tool for the development of anticancer drugs\textsuperscript{16} as well as the design of innovative molecular materials for optical applications\textsuperscript{1h}. 
General

Toluene and diethylether were freshly distilled under argon from appropriate drying agent. All other reagents and solvents were bought from Sigma-Aldrich and used either as such or after drying where required. All complexation reactions were done under argon using standard Schlenk techniques. NMR spectra $^1$H and $^{31}$P were recorded at room temperature on a Bruker Avance 500 spectrometer equipped with a QNP switchable probe $^{13}$C-$^{31}$P-$^{19}$F-$^1$H and operating at proton resonance frequency of 500 MHz. Chemical shifts are reported in parts per million (ppm) relative to Me$_4$Si as external standard; coupling constants are expressed in Hz. $^{31}$P-NMR downfield chemical shifts are expressed with a positive sign, in ppm, relative to external 85% H$_3$PO$_4$. $^{195}$Pt-NMR spectra were performed on a Bruker DRX 500 spectrometer, equipped with a BBI broadband probe; spectra were referenced to a solution of H$_2$PtCl$_6$ in D$_2$O. Some of the samples were measured by the CRMPO - Rennes.

All spectra were obtained using deuterated dichloromethane or chloroform as solvents. The terms s, d, t, q, m indicate respectively singlet, doublet, triplet, quartet, multiplet; b is for broad, dd is doublet of doublets, ddd is doublet of doublets of doublets, AB is the AB spin system. Positive-ion ESI-MS was performed on Esquire 3000 plus ion-trap mass spectrometer (Bruker Daltonik, Bremen, Germany) equipped with an ESI source. Sample solutions were introduced into the ion source at a flow-rate 4 μL m$^{-1}$, capillary voltage 3.8 kV, drying gas temperature 250 °C, drying gas flow rate 5 L m$^{-1}$, nebulizer pressure 14 psi. Nitrogen was used as both nebulizing gas and drying gas. Specific rotations (in deg cm$^2$g$^{-1}$) were measured in a 1 dm thermostated quartz cell on a Perkin Elmer-341 polarimeter. Circular dichroism (in M$^{-1}$cm$^{-1}$) was measured on a Jasco J-815 Circular Dichroism Spectrometer (Biosit platform - Rennes). cis-PtCl$_2$(NCEt)(PPh$_3$) complex 5a, 9 4-aza[6]helicene 4H$_6$$^{10}$ and 5-aza[6]helicene 5H$_6$$^{11}$ were prepared according to literature procedures. The synthesis of complexes 3a,b was described elsewhere.$^4$

CCDC numbers: 1421169 (1), 1421176 (2), 940838 (rac-3a), 947611 (rac-4a), 1062437 (M-4b).
Syntheses

* CIS-PtCl2(PPh3)(4H5): complex 1*

A 50 mL Schlenk tube was loaded with *cis*-PtCl$_2$(NCEt)(PPh$_3$) (50 mg, 0.086 mmol), 4-aza[5]helicene 4 (24 mg, 0.086 mmol) and propionitrile (5 mL). The mixture was stirred at room temperature for 4 days, then dried under vacuum and washed several times with Et$_2$O. 59 mg of yellow solid were recovered (85 % yield) and identified as the product. Single crystals were obtained by slow evaporation of a dichloromethane solution. $^1$H NMR (CDCl$_3$, 500 MHz): $\delta$ 9.65 (d, $^3$J$_{H-H}$ = 8.7, 1H, H5); 9.21 (dd, $^3$J$_{H-H}$ = 8.5,1H, H3); 9.01 (dd, $^2$J$_{H-H}$ = 8.2 Hz, 1H, H1); 8.39 (dd, $^2$J$_{H-H}$ = 8.7 Hz, 1H, H14); 8.28 (d, $^2$J$_{H-H}$ = 9.2 Hz, 1H, H6); 7.97–7.85 (m, 11H, H7/H8/H9/H10/H11, PPh$_3$); 7.55–7.35 (m, 11H, H12/H13, PPh$_3$); 7.28 (ddd, 1H, H2). $^{31}$P NMR (CDCl$_3$, 202 MHz): 3.1, $^1$J$_{P-Pt}$ = 3670 Hz. $^{195}$Pt NMR (CDCl$_3$, 107 MHz): $-$3625 (d), $^1$J$_{Pt-P}$ = 3670 Hz.

* CIS-PtCl2(PPh3)(5H5): complex 2*

A 50 mL Schlenk tube was loaded with *cis*-PtCl$_2$(NCEt)(PPh$_3$) (50 mg, 0.086 mmol), 5-aza[5]helicene 5 (29 mg, 0.104 mmol) and propionitrile (8 mL). The mixture was stirred at room temperature for 1 night. The product precipitates as a yellow powder, to afford 35 mg of product (43 % yield). The solid was crystallized from CH$_2$Cl$_2$/heptane. $^1$H NMR (CDCl$_3$, 500 MHz): 9.77 (d, $^3$J$_{H-H}$ = 4.2, 1H, H6); 9.66 (dd, $^3$J$_{H-H}$ = 8.3, $^4$J$_{H-H}$ = 1.0, 1H, H4); 8.52 (dd, $^3$J$_{H-H}$ = 8.7, $^4$J$_{H-H}$ = 0.8, 1H, H14); 8.50 (dd, $^3$J$_{H-H}$ = 9.8, $^4$J$_{H-H}$ = 0.7, 1H, H1); 8.04 (d, $^3$J$_{H-H}$ = 8.2, H8); 8.01 (d, $^3$J$_{H-H}$ = 9.1, 1H, H9/H10); 7.99 (d, $^3$J$_{H-H}$ = 8.2, 1H, H7); 7.95 (dd, $^3$J$_{H-H}$ = 8.1, $^4$J$_{H-H}$ = 1.2, 1H, H11); 7.91–7.86 (m, 6H, PPh$_3$); 7.80 (ddd, $^3$J$_{H-H}$ = 8.4, 7.0, $^4$J$_{H-H}$ = 1.4 H, H3); 7.56 (ddd, $^3$J$_{H-H}$ = 8.0, 6.7, $^4$J$_{H-H}$ = 1.1 1H, H12) 7.48–7.44 (m, 9H, PPh$_3$); 7.35 (ddd, $^3$J$_{H-H}$ = 8.1, 6.7, $^4$J$_{H-H}$ = 1.1 1H, H13); 7.32 (ddd, $^3$J$_{H-H}$ = 8.4, 7.0, $^4$J$_{H-H}$ = 1.4 1H, H2). $^{31}$P NMR (CDCl$_3$, 202 MHz): 3.1, $^1$J$_{P-Pt}$ = 3660 Hz. $^{195}$Pt NMR (CDCl$_3$, 107 MHz): $-$3522 (d), $^1$J$_{Pt-P}$ = 3660 Hz.

* CIS-PtCl2(PPh3)(5H6): complex 4a*

The synthetic procedure for this complex is the same reported above, using 5-aza[6]helicene 15 as ligand. We obtained a yellow solid as product, with a yield of 69%. The product was
crystallized from a CH$_2$Cl$_2$/iPr$_2$O mixture. $^1$H NMR (CD$_2$Cl$_2$, 500 MHz): 9.44 (dd, $^3$J$_{H-H} = 8.5$, $^4$J$_{H-H} = 1.1$, 1H, $H4$); 9.16 (s, $^3$J$_{H-Pt} = 45$, 1H, $H6$); 8.08 (AB, $^3$J$_{H-H} = 8.4$, 1H, $H8$); 7.98-7.86 (m, 4H, H9/H10/H11/H12); 7.79 (dd, $^3$J$_{H-H} = 8.1$ Hz, $^4$J$_{H-H} = 1.4$, 1H, $H13$); 7.56 (dd, $^3$J$_{H-H} = 8.9$, $^4$J$_{H-H} = 0.8$ Hz, 1H, $H16$); 7.52-7.47 (m, 6H, H PPh$_3$); 7.32 (AB, $^3$J$_{H-H} = 8.4$, 1H, $H7$); 7.30 (ddd, $^3$J$_{H-H} = 8.5$, 6H, $H PPh3$); 7.25 (dd, $^3$J$_{H-H} = 8.1$, 1H, $H1$); 7.21 (ddd, $^3$J$_{H-H} = 8.1$, 1H, $H14$); 7.14-7.10 (m, 3H, H PPh$_3$); 7.06–7.02 (m, 6H, H PPh$_3$); 6.73 (ddd, $^3$J$_{H-H} = 8.5$, 6H, $^4$J$_{H-H} = 1.4$, 1H, $H15$); 6.62 (ddd, $^3$J$_{H-H} = 8.5$, 6H, $^4$J$_{H-H} = 1.4$, 1H, $H2$). $^{31}$P NMR (CD$_2$Cl$_2$, 202 MHz): 6.9, $^1$J$_{P-Pt} = 3853$ Hz.

$P$-(+) and $M$-(-)-Trans-PtCl$_2$(PPh$_3$)(5H$_6$): complex 4b

A 10 mL Schlenk tube was loaded with cis-PtCl$_2$(NCEt)(PPh$_3$) (17.4 mg, 0.0332 mmol), a little excess of enantiopure 5-aza[6]helicene $P$-(+)5H$_6$ (15 mg, 0.0334 mmol) and toluene as solvent (3.0 mL). The mixture was heated under stirring in a sealed tube at 130°C for one night; no precipitation was observed. The yellow solution thus obtained was then completely dried under vacuum, allowing 14.6 mg of compound (59% yield) to be recovered. The same procedure was carried out in parallel using $M$-(-)4H$_6$ as the ligand, yielding the $M$-(-)-4b complex, as confirmed by NMR experiments. $^1$H NMR (CD$_2$Cl$_2$, 400 MHz): $\delta$ 9.88 (d, $^3$J$_{H-H} = 4.2$, $^4$J$_{H-Pt} = 24$ Hz, 1H, $H6$), 9.52 (d, $^3$J$_{H-H} = 8.4$, 1H, $H4$), 8.23 (m, 2H), 8.19 (d, $^3$J$_{H-H} = 8.3$, 1H), 8.19 (d, $^3$J$_{H-H} = 8.3$, 1H), 8.09 (d, $J = 8.7$, 1H), 8.03 (d, $^3$J$_{H-H} = 8.7$, 1H), 8.00 (d, $^3$J$_{H-H} = 8.7$, 1H), 7.96 - 7.87 (m, 7H), 7.58 (m, 1H, $H3$), 7.62 - 7.52 (m, 11H), 7.29 (unres. t, $^3$J$_{H-H} = 7.4$, 1H, $H14$), 6.82 (unres. t, $^3$J$_{H-H} = 7.7$, 1H, $H2$), 6.77 (unres. t, $^3$J$_{H-H} = 7.7$, 1H, $H15$). $^{13}$C NMR (126 MHz, CD$_2$Cl$_2$) $\delta$ 155.89 (CH), 142.63 (C), 136.89 (C), 135.57 (CHx3), 135.48 (CHx3), 132.99 (C), 132.71 (C), 132.62 (C), 131.53(CH), 131.51 (CHx2), 131.16 (CH), 129.72 (C), 129.66 (CH), 129.61 (C), 129.58 (CH), 129.10 (C), 128.99 (CH), 128.63 (CHx3), 128.59 (CH), 128.54 (CHx3), 128.48 (C), 128.43 (CH), 128.37 (CH), 127.41 (CH), 127.38 (CH), 126.93 (CH), 126.57 (CH), 126.55 (CH), 126.35 (CH), 126.19 (CH), 126.16 (C), 126.03 (C), 126.00 (C), 122.74 (C). $^{31}$P NMR (CD$_2$Cl$_2$, 162 MHz): 2.8, $^1$J$_{P-Pt} = 3645$ Hz. $^{195}$Pt NMR (CDCl$_3$, 107 MHz): $-3525$, $^1$J$_{P-Pt} = 3653$ Hz.
X-ray data

The diffraction data for 1 and 2 were recorded at room temperature with a Bruker X8 Prospector APEX-II/CCD diffractometer equipped with a focusing mirror (Cu-Kα radiation, λ = 1.54056 Å). The diffraction data for 3 and 4 were recorded at 150 K with an APEX II Bruker-AXS with Mo-Kα radiation (λ = 0.71073 Å). The structures were determined using direct methods and refined (based on F2 using all independent data) by full-matrix least-square methods (SHELXTL 97). 17 All non-hydrogen atoms were located from different Fourier maps and refined with anisotropic displacement parameters. Hydrogen atoms were added in riding positions.

ASSOCIATED CONTENT

Supporting Information. Further details on NMR, ESI MS, crystallographic information, HPLC separations can be found in the Supporting Information. This material is available free of charge via the Internet at http://pubs.acs.org.

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**Notes**

$^5$ Increasing the temperature resulted in the formation of side products among which the \textit{trans} complex as observed by $^{31}$P NMR and $^{195}$Pt NMR. Moreover, the target complexes showed degradation after few days in solution of chlorinated solvents like chloroform and dichloromethane. For this reason, satisfactory $^{13}$C NMR spectra could not be obtained.

$^#$ The helicity is defined as the dihedral angle between the two terminal rings of the helicene derivative.

$^\$ The $p$ label in $pM$ and $pP$ is used to differentiate planar from helical chirality.
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