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Phosphorescent columnar hybrid materials containing polyionic inorganic nanoclusters

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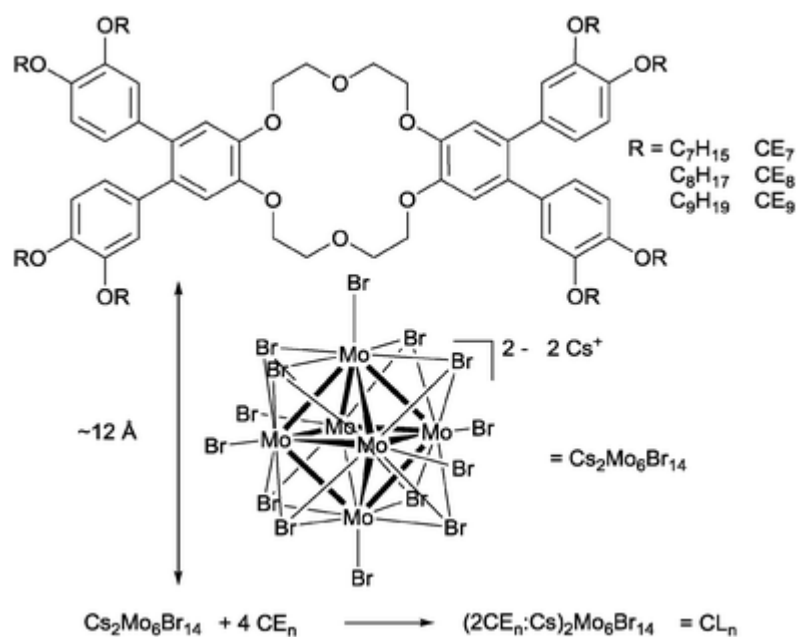
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Abstract

The ternary polyionic inorganic compound Cs₂Mo₆Br₁₄ and 18-crown-6 ethers bearing two o-terphenyl units have been combined to design phosphorescent columnar liquid crystalline hybrid materials. The obtained host–guest complexes are very stable even at high temperatures. Depending on their surrounding atmosphere, these hybrids switch reversibly from a high-to-low luminescence state and show a very stable emission intensity up to 140 °C.

Columnar liquid crystals (CLCs) have attracted much interest in materials chemistry since 1977 when Chandrasekhar discovered them during his studies of discotic benzene hexaalkanoates.¹ These materials self-organise spontaneously usually into columns, show self-healing properties and have been shown to be useful in the design of many devices including photovoltaic solar cells, field effect transistors, one dimensional conductors and light emitting diodes.² It was also found that non-disk-shaped molecules can form columnar mesophases. Recently, they were also used to organise small amounts of bulky inorganic components such as gold nanoparticles,³ gold nanorods,⁴ or quantum dots over a wide range.⁵ Addition of new functionalities to CLCs, e.g., phosphorescence, appears then, to be particularly appealing in terms of application prospects but still remains very challenging in terms of LC phase or emission stability. We present herein the first hybrid CLC containing up to 34.4 wt% of polyionic, bulky, transition metal cluster (TMC) salts. TMCs are aggregates of metal atoms linked by metal–metal bonds.⁶ In particular, octahedral TMC based compounds of the general formula A_nM₆Xⁱ₈X^a₆ where A = alkali cation, M = Mo or Re, Xⁱ = halogen or chalcogen, X^a = halogen or organic ligands are known for their phosphorescence with evidenced dual emission for Cs₂Mo₆Br₁₄.⁷ We recently developed three strategies to integrate these isotropic emitters in smectic or nematic and cholesteric LCs.^{8,9} However, up to now, a TMC containing CLC remains unknown. To design such self-assembling materials, we present herein a supramolecular approach in which the nanocluster alkali counter cations are complexed within the cavity of 18-crown-6 ether derivatives CE_n that were previously used by some of us to design CLC (Scheme 1).¹⁰ Indeed, although the LC properties of CE_n ligands and their respective complexes with monocharged species have been studied,¹¹ we report for the first time the use and the influence of bulky polyionic inorganic species on the mesomorphic properties of CE_n as well as the surprising influence of the supramolecular organization on the luminescence properties of hybrids.



Scheme 1 Representation of CE_n ligands, Cs₂Mo₆Br₁₄ and reaction leading to CL_n complexes.

The dibenzo[18]crown-6 derivatives CE_n with two lateral tetraalkoxy o-terphenyl units, as well as the Cs₂Mo₆Br₁₄ cluster ternary salt, were prepared according to the reported procedure with conforming analytical data.^{10,12}

This TMC salt was chosen for its high thermal stability, a prerequisite for developing new classes of clusters containing thermotropic hybrid materials. Complexes were obtained by mixing 1 equivalent of Cs₂Mo₆Br₁₄ dissolved in acetone with 4 equivalents of CE_n dissolved in dichloromethane. The reaction mixture was then stirred at 50 °C for 14 hours. After evaporation of solvents, the residue was taken up by dichloromethane to remove, by filtration through a microfilter, any traces of unreacted TMC, and heated under a high vacuum to yield the respective dry complexes (2CE_n:Cs)₂Mo₆Br₁₄ = CL_n. Their purity was verified by elemental analysis (ESI). ¹H NMR experiments performed on free CE_n and their respective complexes CL_n did not reveal many changes due to the Cs⁺ complexation within the crown ether cavity (ESI,† Fig. S1 and S2). Only a slight upfield shift of signals corresponding to the methylene groups of the crown ring from 4.26 ppm to 3.96 ppm was observed upon complexation. This shift could be due to (1) a small decrease in the dihedral angle of the OCH₂CH₂O fragment upon ion complexation to accommodate the bulky Cs⁺ cation and (2) an electric field caused by the presence of the salt. It is noteworthy that the complexation of smaller ions such as K⁺ within the macrocyclic cavity induced the opposite effect.¹¹ As previously observed,⁹ it is mandatory to maintain a Cs : CE_n ratio of 1 : 2 in order to maintain a full homogeneity of the cluster salt/CE_n complex.

The LC properties of hybrids were studied by polarized optical microscopy (POM), differential scanning calorimetry (DSC, ESI,† Fig. S4–S6) and, small and wide angle X-ray scattering (SAXS, WAXS, ESI,† Fig. S7–S14), the results of which are shown in Table 1. All complexes show enantiotropic LC phases extending over a wide temperature range. This behaviour largely contrasts with the native CE_n precursor, for which LC phases were observed at lower temperatures and over narrower ranges. However, the increase in the mesophase stability as compared to the parent CE_n precursor is similar to that encountered when soft anions such as iodide or bridging anions such as hexafluorophosphate or

thiocyanate are used to counterbalance the positive charge of a potassium cation embedded within the crown cavity.¹¹ This unexpected result is of particular interest as host-guest complexes between the small 18-crown-6 cavity and the sterically demanding Cs⁺ cation are usually supposed to be weak. It could eventually be attributed to the polyionic character of the TMC salt and the strong interactions existing between the complexed Cs⁺ and [Mo₆Br₁₄]²⁻ dianion which insure a better stability of the supramolecular assemblies upon heating. Surprisingly, the DSC thermogram of CL₈ reveals, upon heating, the formation of two LC phases, whereas upon cooling, only one first order transition corresponding to the isotropic-to-LC phase transition is observed (ESI,† Fig. S5). Fig. 1 shows the optical fan-shaped focal conic textures observed upon cooling for all samples, which are typical indicators of columnar mesophases. Heating CL₈ from 100 °C up to 160 °C did not lead to any noticeable change in the observed textures. Irradiating the samples with UV light illustrates their excellent homogeneity thanks to the well-known red NIR emission of TMC (vide infra). The same experiments performed at 200 °C also reveal a very good homogeneity showing that the interactions between the crown centre and Cs⁺ cations are strong enough to maintain the integrity of the complex even at very high temperatures (ESI,† Fig. S3).

Table 1 Thermodynamic and X-ray diffraction data for CL_n compounds

	Transition	T/°C	ΔC_p/ΔH	Mesophase	Lattice spacing/Å
	a In kJ mol⁻¹ K⁻¹. b In kJ mol⁻¹. c Observed on cooling from the isotropic state; Cr: crystalline state; Col_r: columnar rectangular; Col_h: columnar hexagonal; I: isotropic state.				
CL₉	Cr → Col _r	129.8	7.28 ^b	Col _r	a = 66.3
	Col _r → I	150.8 (148.7) ^c	2.36 ^b	131 °C	b = 44.4
				p2mm	h = 4.0
					Z = 1
CL₈	g → Col _r	91.6	0.43 ^a	Col _r	a = 62.5
	Col _r → Col _h	146.6	2.54 ^b	119 °C	b = 43.8
	Col _h → I	164.4 (162.2) ^c	2.50 ^b	p2mm	h = 4.5
				Col _h	Z = 1
				150 °C	a = 35.5
			p6mm	h = 4.5	
					Z = 1
CL₇	g → Col _r	96.8 (96.9)	0.141 ^a	Col _r	a = 63.1
	Col _r → I	159.2 (147.9) ^c	2.19 ^b		

Transition	T/°C	$\Delta C_p/\Delta H$	Mesophase	Lattice spacing/Å
			127 °C	b = 43.0
			p2mm	h = 4.4
				Z = 1

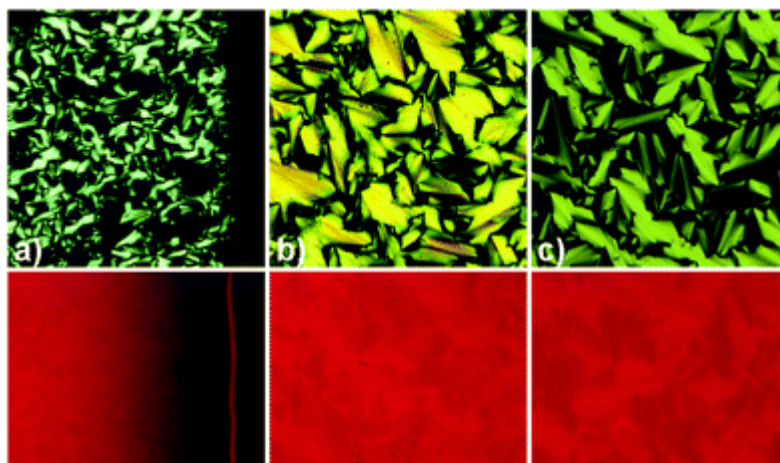


Fig. 1 Polarized Optical Micrographs of (a) CL_9 ($T = 150\text{ °C}$), (b) CL_8 ($T = 163\text{ °C}$) and (c) CL_7 ($T = 150\text{ °C}$) under white light (top) and UV irradiation (bottom).

SAXS and WAXS patterns recorded at 127 °C, 119 °C and 131 °C for CL_7 , CL_8 and CL_9 , respectively, contained sharp and intense reflections in the small angle region. We first analysed the XRD data, to investigate possible lamello-columnar packing, examples of which were independently reported by Donnio, Ziessel and Marcelis for metallomesogens,¹³ donor–acceptor-triads,¹⁴ side-chain LC polymers¹⁵ or triazines with pendant triphenylenes.¹⁶ We first checked whether sharp Bragg reflections indicating a smectic order were present for complexes CL_n . However, neither characteristic layer reflections in the SAXS diffractogram nor the second halo in the wide angle region were visible. Thus, a lamello-columnar phase or a coexisting lamellar and columnar packing could be ruled out.

Hence, the indexation of observed patterns allowed the identification of a rectangular columnar phase with p2mm symmetry for the three complexes (see Tables S1–S4, ESI,† for a full indexation). All WAXS patterns contained a broad and large signal around 4.4 Å arising from the fluid-like behaviour of the molten alkyl chains and from the stacking of the aromatic units within the columns confirming the liquid crystalline nature of all complexes. For CL_8 , additional measurements were performed at 150 °C to characterize the unexpected second mesophase detected via DSC upon heating. Three signals with reciprocal d-spacings in the ratio 1 : $1/\sqrt{3}$: 1/2 were observed in the SAXS pattern and are in good accordance with the formation of a 2D hexagonal lattice with a p6mm symmetry (Col_h). In this case, the unit cell parameter ($a = 35.5\text{ Å}$) was calculated by looking at the position of the most intense scattering signal assigned to the (10) reflection. A comparison of previously published XRD data for metal-free crown ether CE_{11} ¹⁷ with those for CL_n (Table 1) revealed that the symmetry of the Col_h lattice changed upon complexation. While the lattice parameter a remains relatively unaffected, i.e. 66.1 Å (CE_{11}) vs. 66.3 Å (CL_9) despite the difference of 2 CH_2 units in the

side chain, the b values almost doubled from 25.1 Å (CE_{11}) to 44.4 Å (CL_9). Halos were in a similar range (4.6 Å vs. 4.0 Å). To rationalize these results, we propose that the nanocluster, because of its bulkiness, cannot be accommodated within the $c2mm$ lattice of CE_n , resulting in a $p2mm$ lattice with an elongated b axis, so that the cluster fits in the centre of the unit cell and is surrounded by 4 stacks of crown ethers CE_9 , with an intracolumnar distance of 4 Å (Fig. 2). This results in a height of 12 Å for the unit cell and the (1 : 4) cluster/ CE_9 stoichiometry is maintained. In that case the inorganic moieties volumic fraction is around 9 vol%. For the corresponding Col_h phase of CL_8 , the observed a value should allow a hexagonal packing of the discotic crowns CE_8 , while each $\text{Mo}_6\text{Br}_{14}^{2-}$ unit is encapsulated by three crowns and leads to a volumic fraction of 22 vol% for $\text{Cs}_2\text{Mo}_6\text{Br}_{14}$.

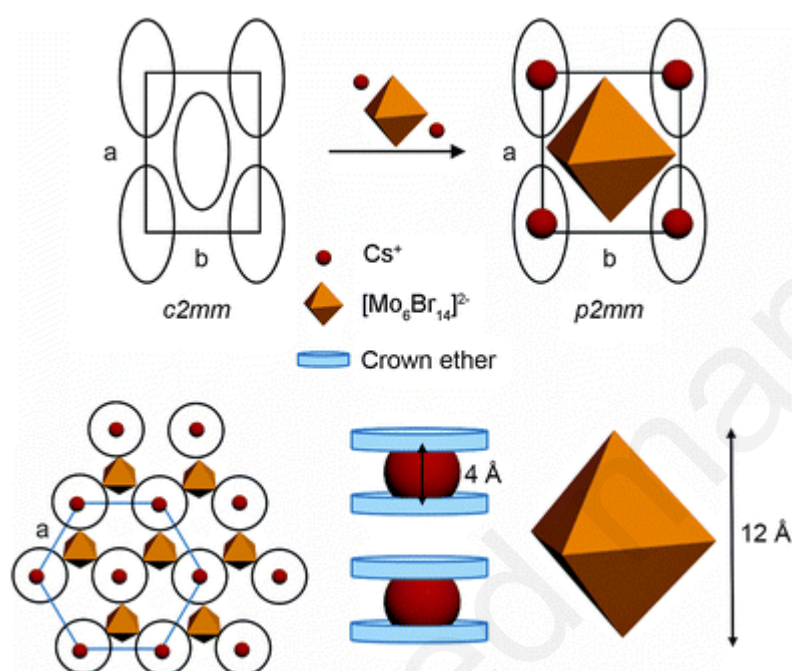


Fig. 2 Proposed packing models of CL_n compounds in the Col_r (top, $n = 7-9$) and Col_h (bottom, $n = 8$) mesophase.

As expected, the emission properties of $[\text{Mo}_6\text{Br}_{14}]^{2-}$ are retained within the hybrid materials (see ESI,† Fig. S15).¹⁸ The photoluminescence spectra ($\lambda_{\text{exc}} = 380$ nm) show a broad and intense band extending from 550 nm up to 950 nm. The influence of the hybrid nanostructures on the inorganic emitter emission was investigated by temperature-dependent luminescence studies. Samples were heated up to the isotropic state and slowly cooled to reach a nanostructured glassy or crystalline state at 25 °C prior to measurements. The phosphorescence signals were collected at every degree upon heating (see the ESI,† Fig. S16 and S17). As previously reported,¹⁹ if the emission intensity variation upon heating is only due to an increase of non-radiative deactivation processes, it should follow an exponential law of the type: $I(T) = I_0 \exp(-A/T)$ where A is constant and I_0 the emission intensity at a given temperature. Thus, reporting $\ln(I/I_{300}) = f(1/T)$ should give a straight line and this is what is observed for $\text{Cs}_2\text{Mo}_6\text{Br}_{14}$ (Fig. 3). Surprisingly, increasing the temperature from 25 °C up to 140 °C does not have much effect on the ability of the hybrid materials to emit light as an emission intensity loss of only less than 10% was recorded (the intensity loss at 140 °C is about 77% for $\text{Cs}_2\text{Mo}_6\text{Br}_{14}$ in its powder form). At higher temperatures, the obtained curves significantly deviate from the straight line with a similar behaviour for CL_7 and CL_9 . The separation of the clusters from each other by isolating triangular arrays of crown ethers might

be the reason for the increased luminescence intensity of CL_8 in the high temperature Col_h phase as compared to the low temperature Col_l phase. Presumably, the crown ethers act as an isolating matrix for luminescent clusters, making these hybrids attractive alternatives with regards to polymers for applications requiring temperatures around 100 °C.²⁰ Thus, the columnar nanostructures have a strong influence on the emissive ability of hybrids despite the isotropic nature of the inorganic emitter. Let us emphasize that such unprecedented high stability of the emission intensity vs. temperature is particularly rare and has never been reported so far for TMC or TMC-containing hybrid materials.

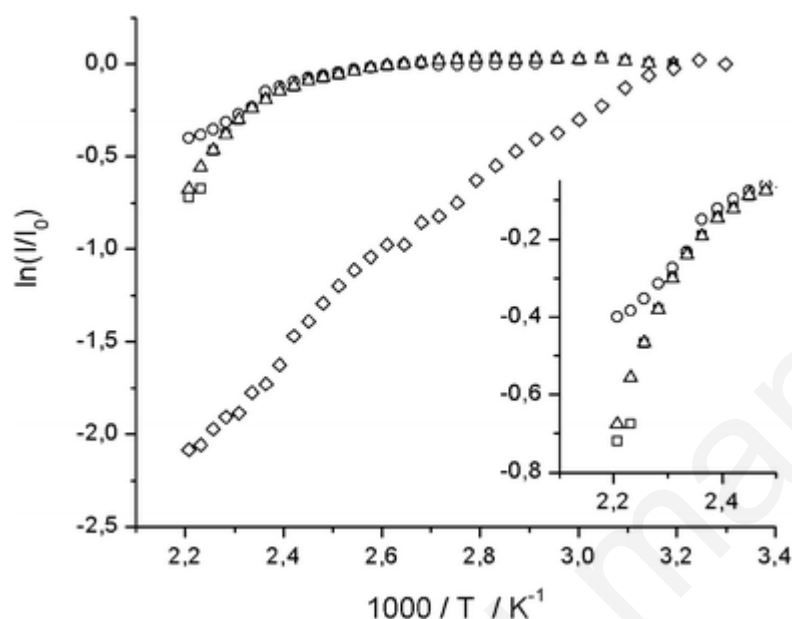


Fig. 3 Evolution of luminescence intensity at 700 nm with temperature for $\text{Cs}_2\text{Mo}_6\text{Br}_{14}$ (diamond), CL_9 (triangle), CL_8 (circle) and CL_7 (square). Inset: Zoom of the high temperature area for hybrids.

As illustrated by absolute quantum yield measurements performed in an air or N_2 atmosphere, the emission intensity can also be modulated by modifying the local atmosphere of the material. In all cases, a reversible and instantaneous switch from 0.05 or 0.07 in air up to 0.20, 0.26 or 0.30 in an N_2 atmosphere was measured for CL_7 , CL_8 and CL_9 , respectively (Fig. S18 and Video S1, ESI†). This phenomenon, which shows that these materials might be used as efficient oxygen sensors, is due to the O_2 permeability of CL_n and is based on the ability of the excited states of the phosphorescent metal clusters to be quenched by molecular triplet oxygen, thus creating singlet oxygen.²¹ Although we demonstrated recently the same effect when nanoclusters are introduced into gas-permeable polymer matrices,²² this is the first time that such behaviour is reported in a liquid crystalline material.

In conclusion, we show herein, that nanometric luminescent transition metal cluster compounds are easily integrated in a CLC hybrid material by using a supramolecular approach. In these materials, Cs^+ cluster unit counter cations are complexed within the cavity of crown ether derivatives functionalized by o-terphenyl units that are well known to generate columnar liquid crystalline phases. We show that the introduction of cluster salts results in columnar rectangular phases of $p2mm$ symmetry and that their thermal stability is largely enhanced as compared to the parent crown ether CE_n . Surprisingly, a second columnar hexagonal phase was observed on heating for CL_8 . Luminescence studies reveal that the red

NIR clusters emission stability upon heating is unexpectedly largely enhanced as compared to its precursor namely Cs₂Mo₆Br₁₄. Moreover, we demonstrate for the first time that modifying the atmosphere surrounding the LC hybrids with an inert gas induces a large, instantaneous and reversible increase of its emission efficiency thanks to the good O₂ permeability of the LC matrix. The columnar mesophases observed for these novel clustomesogens in combination with their luminescence properties make them promising candidates for oxygen sensors, optoelectronic or photovoltaic devices. As a further step towards such applications, the mesophase temperature ranges need to be lowered by molecular design. Work towards this goal is currently in progress.

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References

1. S. Chandrasekhar, B. K. Sadashiva and K. A. Suresh, *Pramana*, 1977, 9, 471–480 CrossRef CAS .
2. J. W. Goodby, P. J. Collings, T. Kato, C. Tschierske, H. F. Gleeson and P. Raynes, *Handbook of Liquid Crystals: Applications of Liquid Crystals*, Wiley-VCH Verlag GmbH & Co. KGaA, 2nd edn, 2014, vol. 8 Search PubMed .
3. S. Kumar and V. Lakshminarayanan, *Chem. Commun.*, 2004, 1600–1601 RSC .
4. X. Feng, L. Sosa-Vargas, S. Umadevi, T. Mori, Y. Shimizu and T. Hegmann, *Adv. Funct. Mater.*, 2015, 25, 1180–1192 CrossRef CAS .
5. S. Kumar and L. K. Sagar, *Chem. Commun.*, 2011, 47, 12182–12184 RSC ; M. Kumar and S. Kumar, *RSC Adv.*, 2015, 5, 1262–1267 RSC .
6. F. A. Cotton, *Inorg. Chem.*, 1964, 3, 1217–1220 CrossRef CAS .
7. T. G. Gray, C. M. Rudzinski, D. G. Nocera and R. H. Holm, *Inorg. Chem.*, 1999, 38, 5932–5933 CrossRef CAS PubMed ; A. W. Maverick and H. B. Gray, *J. Am. Chem. Soc.*, 1981, 103, 1298–1300 CrossRef ; K. Costuas, A. Garreau, A. Bulou, B. Fontaine, J. Cuny, R. Gautier, M. Mortier, Y. Molard, J.-L. Duvail, E. Faulques and S. Cordier, *Phys. Chem. Chem. Phys.*, 2015, 17, 28574–28585 RSC .
8. M. Prevot, M. Amela-Cortes, S. k. Manna, S. Cordier, T. Roisnel, H. Folliot, L. Dupont and Y. Molard, *J. Mater. Chem. C*, 2015, 3, 5152–5161 RSC ; S. M. Wood, M. Prévôt, M. Amela-Cortes, S. Cordier, S. J. Elston, Y. Molard and S. M. Morris, *Adv. Opt. Mater.*, 2015, 3, 1368–1372 CrossRef CAS ; Y. Molard, A. Ledneva, M. Amela-Cortes, V. Circu, N. G. Naumov, C. Meriadec, F. Artzner and S. Cordier, *Chem. Mater.*, 2011, 23, 5122–5130 CrossRef ; Y. Molard, F. Dorson, V. Circu, T. Roisnel, F. Artzner and S. Cordier, *Angew. Chem., Int. Ed.*, 2010, 49, 3351–3355 CrossRef PubMed .

9. S. K. Nayak, M. Amela-Cortes, C. Roiland, S. Cordier and Y. Molard, *Chem. Commun.*, 2015, 51, 3774–3777 RSC .
10. A. Schultz, S. Laschat, A. Saipa, F. Giesselmann, M. Nimtz, J. L. Schulte, A. Baro and B. Miehlich, *Adv. Funct. Mater.*, 2004, 14, 163–168 CrossRef CAS .
11. M. Kaller, S. Tussetschlaeger, P. Fischer, C. Deck, A. Baro, F. Giesselmann and S. Laschat, *Chem. – Eur. J.*, 2009, 15, 9530–9542 CrossRef CAS PubMed .
12. K. Kirakci, S. Cordier and C. Perrin, *Z. Anorg. Allg. Chem.*, 2005, 631, 411–416 CrossRef CAS .
13. M. Stepien, B. Donnio and J. L. Sessler, *Chem. – Eur. J.*, 2007, 13, 6853–6863 CrossRef CAS PubMed ; T. Cardinaels, K. Driesen, T. N. Parac-Vogt, B. Heinrich, C. Bourgogne, D. Guillon, B. Donnio and K. Binnemans, *Chem. Mater.*, 2005, 17, 6589–6598 CrossRef ; R. Ziessel, L. Douce, A. El-Ghayoury, A. Harriman and A. Skoulios, *Angew. Chem., Int. Ed.*, 2000, 39, 1489–1493 CrossRef ; A. El-Ghayoury, L. Douce, A. Skoulios and R. Ziessel, *Angew. Chem., Int. Ed.*, 1998, 37, 1255–1258 CrossRef .
14. K.-Q. Zhao, L.-L. An, X.-B. Zhang, W.-H. Yu, P. Hu, B.-Q. Wang, J. Xu, Q.-D. Zeng, H. Monobe, Y. Shimizu, B. Heinrich and B. Donnio, *Chem. – Eur. J.*, 2015, 21, 10379–10390 CrossRef CAS PubMed .
15. I. Tahar-Djebbar, F. Nekelson, B. Heinrich, B. Donnio, D. Guillon, D. Kreher, F. Mathevet and A.-J. Attias, *Chem. Mater.*, 2011, 23, 4653–4656 CrossRef CAS .
16. C. P. Umesh, A. T. M. Marcelis and H. Zuilhof, *Liq. Cryst.*, 2015, 42, 1450–1459 CrossRef CAS .
17. M. Kaller, C. Deck, A. Meister, G. Hause, A. Baro and S. Laschat, *Chem. – Eur. J.*, 2010, 16, 6326–6337 CrossRef CAS PubMed .
18. S. Cordier, Y. Molard, K. A. Brylev, Y. V. Mironov, F. Grasset, B. Fabre and N. G. Naumov, *J. Cluster Sci.*, 2015, 26, 53–81 CrossRef CAS ; S. Cordier, F. Grasset, Y. Molard, M. Amela-Cortes, R. Boukherroub, S. Ravaine, M. Mortier, N. Ohashi, N. Saito and H. Haneda, *Inorg. Organomet. Polym. Mater.*, 2015, 25, 189–204 CrossRef ; M. Amela-Cortes, A. Garreau, S. Cordier, E. Faulques, J.-L. Duvail and Y. Molard, *J. Mater. Chem. C*, 2014, 2, 1545–1552 RSC .
19. S. Suarez, O. Mamula, D. Imbert, C. Piguet and J. C. G. Bunzli, *Chem. Commun.*, 2003, 1226–1227 RSC .
20. M. Shi, Y. Guan, J. Zhang and X. Wang, *RSC Adv.*, 2015, 5, 34900–34907 RSC .
21. J. A. Jackson, M. D. Newsham, C. Worsham and D. G. Nocera, *Chem. Mater.*, 1996, 8, 558–564 CrossRef CAS ; J. A. Jackson, C. Turro, M. D. Newsham and D. G. Nocera, *J. Phys. Chem.*, 1990, 94, 4500–4507 CrossRef .
22. M. Amela-Cortes, S. Paofai, S. Cordier, H. Folliot and Y. Molard, *Chem. Commun.*, 2015, 51, 8177–8180 RSC .

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Electronic supplementary information (ESI) available: Experimental details, DSC traces, POM micrographs, SAXS and WAXS patterns, temperature dependent emission spectra, video of CL7 under irradiation alternatively in an air or inert atmosphere. See DOI: 10.1039/c5cc09110a

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