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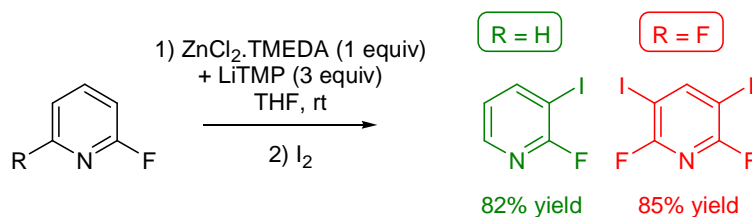
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Graphical Abstract



Deprotometalation of Substituted Pyridines and Regioselectivity-Computed CH Acidity Relationships

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Keywords: CH acidity; pyridine; mixed lithium-zinc base; substituent effect

Abstract:

A series of methoxy- and fluoro-pyridines have been deprotometalated in tetrahydrofuran at room temperature by using a mixed lithium-zinc combination obtained from ZnCl₂.TMEDA (TMEDA = *N,N,N',N'*-tetramethylethylenediamine) and LiTMP (TMP = 2,2,6,6-tetramethylpiperidino) in a 1:3

ratio, and the metalated species intercepted by iodine. Efficient functionalization at the 3 position was observed from 4-methoxy, 2-methoxy, 2,6-dimethoxy, 2-fluoro and 2,6-difluoropyridine, and at the 4 position from 3-methoxy and 2,3-dimethoxypyridine. Interestingly, clean dideprotonation was noted from 3-fluoropyridine (at C2 and C4) and 2,6-difluoropyridine (at C3 and C5).

The obtained regioselectivities have been discussed in light of the CH acidities of the substrates, determined both in the gas phase (DFT B3LYP and G3MP2B3 levels) and in THF solution. In the case of methoxypyridines, the pK_a values have also been calculated for complexes with LiCl and LiTMP.

1. Introduction

Pyridines are present in numerous biological key positions, with derivatives such as nicotine, nicotinamide (niacin), nicotinamide adenine dinucleotide phosphate (NADP) and pyridoxine (vitamin B₆). In addition to occurring in pharmaceuticals and agrochemicals,¹ pyridines are important part of organic materials.²

To functionalize regioselectively these heterocycles, deprotonative lithiation³ appears as a valuable tool; alkyllithiums and hindered lithium dialkylamides have been largely used for this purpose.⁴ Nevertheless, the low compatibility between π -deficient heteroaromatics on the one hand, and either the alkyllithiums used as base or the heteroaryllithiums generated by deprotonation on the other hand, implies very low reaction temperatures and limits the scope of the method.

Alternative deprotometalation methods have since emerged with the use of metal additives allowing chemoselective reactions to be performed.⁵ In the pyridine series, pioneering reagents combining alkyllithiums with LiDMAE (DMAE = 2-dimethylaminoethoxide) proved to direct the deprotonation to the 2 position.⁶ In addition, (R)_n(R')_nMLi-type reagents, in which M is a nonalkali metal (e.g. Cu,⁷ Zn,⁸ Cd⁹), have been reported for their ability to deprotometalate sensitive aromatics including pyridines.

In this context, we have developed a lithium-zinc combination, prepared from ZnCl₂·TMEDA (TMEDA = *N,N,N',N'*-tetramethylethylenediamine) and LiTMP (TMP = 2,2,6,6-tetramethylpiperidino) in a 1:3 ratio, capable of accomplishing room-temperature deprotometalation of a large range of

substrates.¹⁰ Probably in relation with high steric hindrance, this base proved by NMR and DFT studies to be a 1:1 mixture of the homometallic amides rather than a lithium zincate,^{10a} a result confirmed by DOSY NMR spectroscopy.¹¹ The idea of a deprotonation through LiTMP occurring first, followed by Zn(TMP)₂-mediated transmetalation, was proposed in 2008 to rationalize its synergic behavior,^{10a} and supported by others who pinpointed LiTMP·2LiCl±TMEDA as the possible active lithiating base.¹¹ This “trans-metal trapping”¹² has since been extended to other pairs behaving synergically such as mixtures of LiTMP on the one hand, and ZnCl₂·2LiCl, MgCl₂ or CuCN·2LiCl on the other hand.¹³

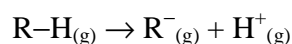
We here describe the use of this lithium-zinc combination for the deprotonative metalation of pyridines substrates and comment on the results, and notably on the regioselectivities observed, on the basis of the CH acidities in THF (THF = tetrahydrofuran).

2. Results and Discussion

Computational aspects

The literature data on CH acidity of pyridines are scanty.¹⁴ The main reasons are the necessity of using very strong bases at low temperatures and the side reactions of the azine carbanions generated.⁴ A brief review of the papers devoted to experimental and theoretical investigation of azine CH acidities is presented in our previous publications.^{10e,10h} In the present work we report the values of CH acidity of pyridines both in gas-phase ($\Delta_{\text{acid}}G$) and in THF solution ($\text{p}K_{\text{a}}$), which were calculated by means of quantum chemistry.

Data on gas-phase acidity are important for the development of an acidity scale not depending on the basicity of the solvent in which the ionization takes place. Its recommended measure is the Gibbs energy ($\Delta_{\text{acid}}G$) of deprotonation of the substance:



All the calculations were performed by using the Gaussian 03 software. Two different approaches, namely (i) the DFT B3LYP level of theory and (ii) the hybrid G3MP2B3 method,¹⁵ were used as described below.

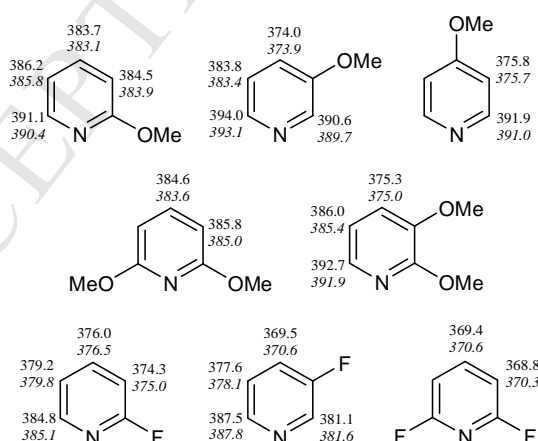
DFT B3LYP level of theory. The geometries were fully optimized by using the 6-31G(d) basis set. No symmetry constraints were implied and minima were found by the potential energy surface scans. In order to perform stationary point characterization and to calculate zero-point vibrational energies and thermal corrections, vibrational frequencies were calculated at the same level of theory. The single point energies were obtained using the 6-311+G(d,p) basis set and tight convergence criteria.

Hybrid G3MP2B3 method. G3MP2B3 is a modified version of Gaussian-3 (G3) theory for calculating energies of molecules with high accuracy. G3MP2B3 uses MP2 instead of MP4 for the basis set extension corrections and geometries and zero point vibrational energies calculated at the B3LYP/6-31G(d) level of theory.

The $\Delta_{\text{acid}}G$ values were calculated from the calculated values of Gibbs energy of species in gas phase by the following equation:

$$\Delta_{\text{acid}}G = G^0_{298}(\text{R}^-) + G^0_{298}(\text{H}^+) - G^0_{298}(\text{R-H})$$

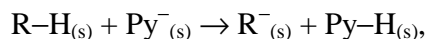
As shown in Scheme 1, the values of the gas-phase acidity calculated at DFT B3LYP and G3MP2B3 levels are in excellent agreement (maximum deviation: 1.5 kcal mol⁻¹), and are typical of those for very weak acids.



Scheme 1. Calculated $\Delta_{\text{acid}}G$ values for the studied methoxypyridines and fluoropyridines (upper straight values obtained at DFT B3LYP level, and lower italic values at G3MP2B3 level).

The solvent effects were treated by using the IEF formalism of polarized continuum model (PCM) with the default parameters for THF. The PCM energies were calculated at the B3LYP/6-311+G(d,p) level using geometries optimized for isolated structures.

The pK_a values were further calculated by means of the following homodesmotic reaction:



where Py-H is bare pyridine, whose pK_a value in THF (40.2, for position 4 of the ring) is known experimentally.

The homodesmotic reaction Gibbs energy was calculated using the following equation:

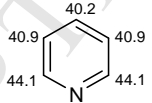
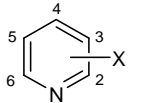
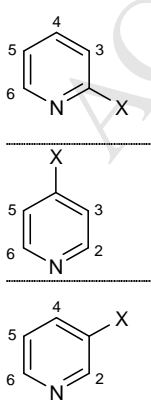
$$\Delta_r G_s = \sum_{\text{products}} G_s - \sum_{\text{reactants}} G_s.$$

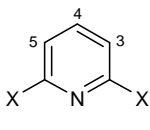
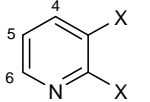
Further, it can be proved that:

$$pK_a(R-H) = 40.2 + \frac{\Delta_r G_s}{RT} \cdot \frac{1}{\ln 10}.$$

The CH acidity values for fluoro- and methoxy-pyridines in THF solution can be compared with those previously determined for chloro- and bromo-pyridines (Table 1).^{10e}

Table 1. Pyridine pK_a (THF) changes caused by substituent effects.

		ΔpK_a (THF)				
		C2	C3	C4	C5	C6
	X					
	Cl	-	-6.9	-4.5	-3.1	-3.9
	Br	-	-7.4	-5.2	-3.5	-4.4
	F	-	-6.9	-3.3	-2.2	-3.4
	X					
	Cl	-3.7	-7.9	-	-7.9	-3.7
	Br	-4.3	-8.8	-	-8.8	-4.3
	OMe	+0.4	-3.8	-	-3.8	+0.4
	X					
	Cl	-6.2	-	-8.1	-4.1	-2.5
	Br	-6.3	-	-8.8	-4.5	-3.0
	F	-7.1	-	-8.7	-3.4	-1.7
OMe	-1.3	-	-3.8	-0.1	+1.3	

	Cl	-	-9.3	-8.4	-9.3	-
	Br	-	-9.9	-9.7	-9.9	-
	F	-	-9.0	-6.4	-9.0	-
	OMe	-	+0.7	+1.2	+0.7	-
	Cl	-	-	-11.3	-6.4	-6.2
	Br	-	-	-12.2	-7.0	-6.8
	OMe	-	-	-3.0	+1.4	+1.6

While the values of $pK_a(\text{THF})$ of methoxypyridines are rather high and are of magnitude of monosubstituted benzenes,¹⁶ the insertion of halogen makes the compounds much more acidic. When present at the 2 position, fluorine exerts a short range effect similar to that of chlorine; in contrast, positions remote from fluorine are less acidified (than with chlorine and, above all, bromine), in agreement with the electronic effects of the substituents.¹⁷ The influence of the methoxy group is not quite that prominent. For molecules containing a methoxy group at the 2-, 4- and even 3-position, only adjacent sites are acidified. That the strongest acidifying effect is observed at both 2- and 4-sites, neighboring to the substituent, is a general trend for 3-substituted pyridines. In addition, this effect is stronger at the 4-position, which is far from the nitrogen lone pair, and also more affected because of concordant halogen/methoxy and nitrogen effects. In the case of 2,6-dihalogenated pyridines, the strongest acidifying effect is observed at the 3- and 5-positions of the ring. In addition, the 4-position also becomes much more acidic, notably compared to 2-monohalogenated substances, because of the cooperative long range electron-withdrawing effects of both halogens. With 2,6-dimethoxypyridine, the short range $-I$ effects exhibited at the 3- and 5-position of the cycle by the methoxy groups are concealed by their stronger long range “para” electron-donating ($+M$) effects. From 2,3-dimethoxypyridine, the “meta” electron-donating effect exhibited by the methoxy group at C2 does not offset the short range inductive acidifying effect exerted at C4 by the methoxy group at C3; as a result, the 4 position is more acidified for 2,3-dimethoxypyridine than for 2,6-dimethoxypyridine.

By using the computational approach described above, we also investigated the influence of substrate coordination to lithium species on the $pK_a(\text{THF})$ values (Figure 1). It is worth noting that, besides impacting the $pK_a(\text{THF})$ values, complexation to lithium can favor the deprotonation at a neighboring

site through complex-induced proximity effect.¹⁸ To this purpose, we considered 1:1 complexes of 3-, 4- and 2-methoxypyridine with LiCl and LiTMP. With LiCl, two types of coordination - namely through nitrogen and oxygen - were investigated, and several trends can be noted here. As expected, coordination of methoxypyridine to LiCl by nitrogen is more effective than by oxygen (with the corresponding isomeric complexes up to 8 kcal mol⁻¹ more stable). In the case of O-Li complexation, the anions formed under deprotonation at positions adjacent to the methoxy group benefit from additional cyclic stabilization (values with asterisk on Figure 1). Finally, coordination of the metal by the ring nitrogen lone pair drastically increases CH acidity of the adjacent positions, making them competitive in deprotometalation processes.

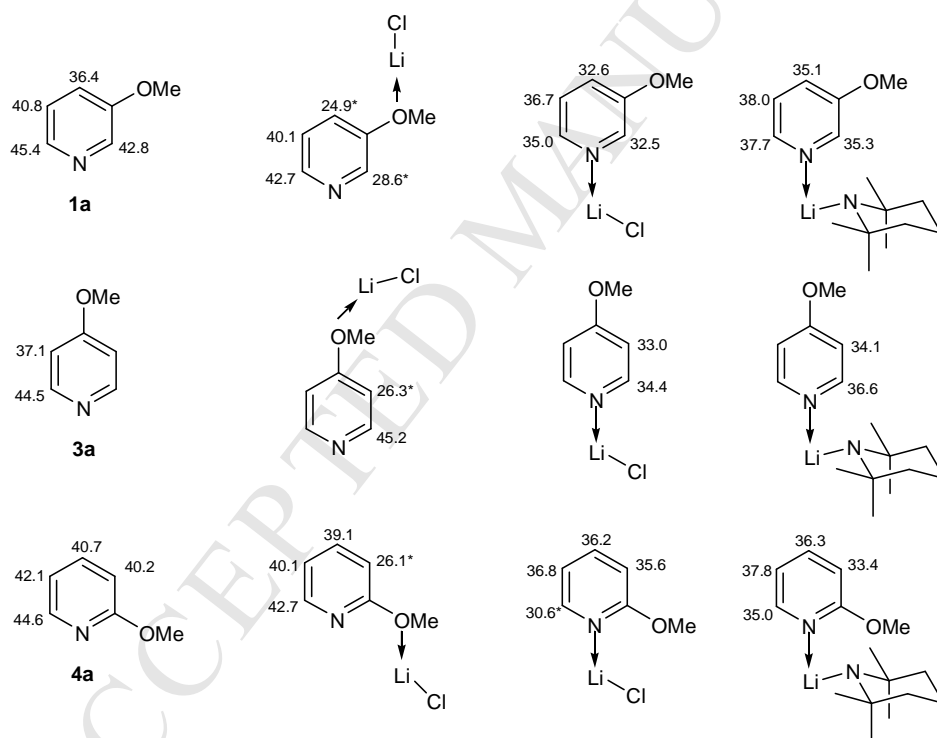


Figure 1. Calculated $pK_a(\text{THF})$ values for 3-, 4- and 2-methoxypyridine, as well as their complexes with LiCl or/and LiTMP.

Synthetic aspects

We first focused on the deprotometalation of 3-methoxypyridine (**1a**). **1a** can be deprotolithiated regioselectively at its 2 position by using BuLi·TMEDA in THF at -40 °C, as evidenced by trapping

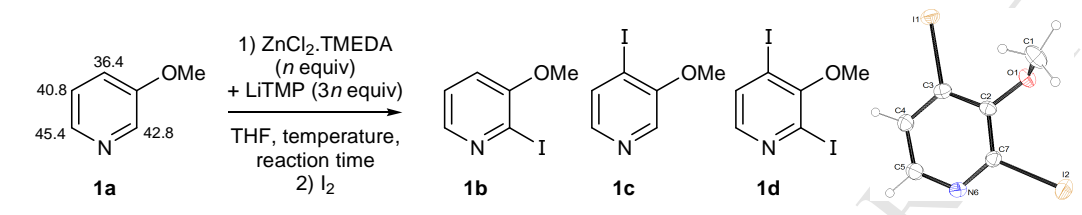
with acetaldehyde to afford the expected alcohol in 49% yield.¹⁹ A favored coordination of the metal of the base by the ring nitrogen (when compared to the methoxy substituent) is advanced to rationalize this result. By increasing the deprotometalation temperature to -23 °C, which proves possible by using less nucleophilic mesityllithium, 2-substituted 3-methoxypyridines are formed in higher yields.²⁰ Using LiDA (DA = diisopropylamino) in THF at -42 °C only proves efficient in the presence of chlorotrimethylsilane as an in situ trap, but mainly gives rise to a mixture of the 2- and 4-substituted derivatives (assumed to be the kinetic and thermodynamic products, respectively).²⁰

Magnesiumation of 3-methoxypyridine (**1a**) can be achieved selectively at its 4 position by using a dipotassium tetra(alkyl)magnesate including a polydentate N-donor, (PMDETA)₂K₂Mg(CH₂SiMe₃)₄ (PMDETA = *N,N,N',N'',N'''*-pentamethyldiethylenetriamine), in hexane at 0 °C.²¹ Similarly, **1a** can be 4-zincated by using Kondo and Uchiyama's TMP-zincate, Li(TMP)Zn(*t*Bu)₂,²² in THF at 25 °C.²³

The 1:1 mixture of LiTMP and Zn(TMP)₂, obtained from ZnCl₂·TMEDA and LiTMP,^{10a,11} was involved in the reaction with **1a** in order to better understand the behavior of this combination (Table 2). Upon treatment by the base (0.5 equiv of each metal amide) in THF for 2 h at room temperature (or 0 °C) and subsequent trapping with iodine, **1a** was converted to a mixture of the 2- and 4-iodo derivatives (**1b** and **1c**), both purified, isolated in 10 and 85% yield, respectively (entry 1). In order to see if this ratio could be modified, reactions were performed at a lower temperature or using a shorter reaction time. Whereas no reaction was noted at -40 °C (entry 2), shortening the reaction time to 10 min led to a mixture of **1b** (50% yield) and **1c** (25% yield) (entry 3). By keeping a 2 h contact at room temperature with the base, extending the base amount to 1 equiv of each metal amide led to the competitive formation of the 2,4-diiodo derivative **1d** (30% yield, entry 4). With 2 equiv of each metal amide, **1c** (58% yield) and **1d** (35% yield) were the only products isolated (entry 5); in spite of caking of the reaction mixture, the formation of **1d** proved favored by extending the contact time between base and substrate to 20 h (entry 6).^{10h} By using LiTMP (2 equiv) under similar reaction conditions, the 4-iodo derivative **1c** was obtained in 6% yield together with traces of **1b** and the 4,4'-dimer²⁴ (25%

yield). Upon a similar treatment but in the presence of TMEDA (2 equiv), **1c** and the 4,4'-dimer formed in 5 and 11% yield, respectively, together with traces of **1b**.

Table 2. Calculated $pK_a(\text{THF})$ values for 3-methoxypyridine (**1a**), deprotonation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis and ORTEP diagram (30% probability) of **1d**.

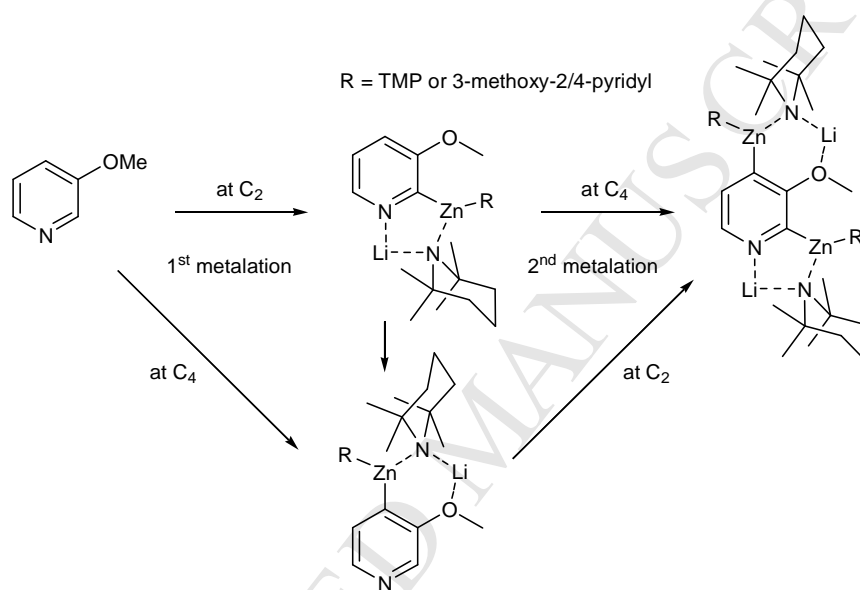


Entry	<i>n</i> (equiv)	Temperature (°C)	Reaction time	Products, yields (%)
1	0.5	20 ^a	2 h	1b , 10 ^b 1c , 85 ^b -
2	0.5	-40	2 h	- - -
3	0.5	20	10 min	1b , 50 ^b 1c , 25 ^b -
4	1	20	2 h	1b , 11 ^b 1c , 50 ^b 1d , 30 ^b
5	2	20	2 h	- 1c , 58 ^b 1d , 35 ^b
6	2	20	20 h	- 1c , 45 ^c 1d , 55 ^c

^a A similar result was obtained at 0 °C and using Et₂O as solvent at room temperature. ^b Yield after purification by column chromatography. ^c Ratio determined from the ¹H NMR spectrum of the crude.

Some conclusions can be drawn from these data. The different results above confirm that the 2- and 4-metalated 3-methoxypyridines correspond to kinetic and thermodynamic products, respectively. The kinetic products are observed by using organolithiated (alkyl/aryl) bases, and coordination of the ring nitrogen to lithium favors reaction at C2 by both proximity effect^{18a} and strongly decreasing of the corresponding pK_a values (Figure 1). Results also evidence formation of transient 2-metalated 3-methoxypyridines by using amido-containing bases such as LiDA, LiTMP and 1:1 LiTMP-Zn(TMP)₂. If chlorotrimethylsilane is not present to intercept the 2-metalated species, they change to presumably more stable 4-metalated derivatives (reversible reactions). Such an isomerization could also take place by using TMP-zincate, with lithium coordination by the pyridine nitrogen first giving rise to a 2-metalated species,²⁵ provided that this conversion is faster than the consumption of generated H-TMP by the *tert*-butyl ligands.²⁶

The formation of 2,4-dimetalated species is only observed by using 1:1 LiTMP-Zn(TMP)₂ (1 or 2 equiv) for which “trans-metal trapping”¹² can take place. If lithium is coordinated by the pyridine nitrogen in the 2-metalated 3-methoxypyridine, methoxy could be still available to stabilize a 4-metalated species. Similarly, if lithium is coordinated by the methoxy group in the 4-metalated 3-methoxypyridine, pyridine nitrogen could be free to play a role in the formation of a 2-metalated derivative (Scheme 2).



Scheme 2. Proposed pathways to the 2,4-dimetalated 3-methoxypyridine.

By contrast with **1a**, for which deprotometalation is observed at C4 by using 1:1 LiTMP-Zn(TMP)₂ in THF at room temperature for 2 h, 3-chloro and 3-bromopyridine are attacked at C2 under similar conditions.^{10e} Such a difference could be in relation with a higher ability to coordinate lithium in the case of the methoxy group, contributing to a greater stabilization of the corresponding 4-metalated derivative, as shown in Figure 2.

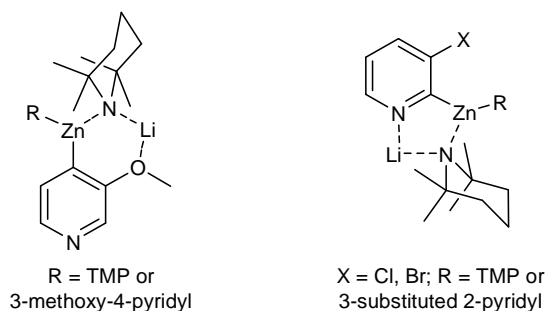


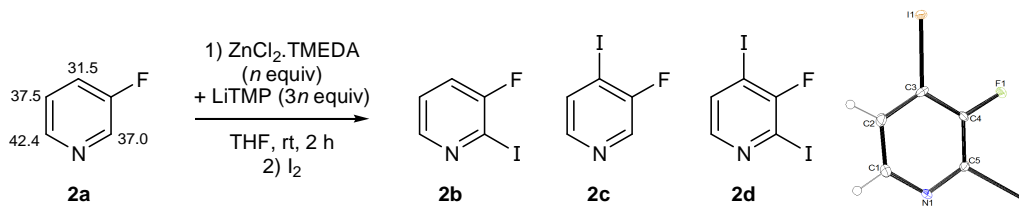
Figure 2. Proposed metalated species for 3-methoxy, 3-chloro and 3-bromopyridine.

This difference led us to study the behavior of 3-fluoropyridine (**2a**). **2a** reacts with BuLi·TMEDA at $-40\text{ }^{\circ}\text{C}$ to afford either the 2- (kinetic) or the 4- (thermodynamic) lithio species, depending on if Et₂O or more polar THF is respectively employed as solvent; the isomerization was supposed to occur through the formation of the 2,4-dilithio species.²⁷ The use of BuLi·DABCO (DABCO = 1,4-diazabicyclo[2.2.2]octane) in Et₂O at $-75\text{ }^{\circ}\text{C}$ allows **2a** to be selectively attacked next to nitrogen.^{27b} In contrast, **2a** gives 4-substituted products by using LiDA in THF at $-75\text{ }^{\circ}\text{C}$.^{28,27b} A similar regioselectivity can be reached by using BuLi·*t*BuOK in THF at $-75\text{ }^{\circ}\text{C}$ ²⁹ and LiMgBu₃ in THF at $-10\text{ }^{\circ}\text{C}$.³⁰

Thus, 3-fluoropyridine (**2a**) was involved in the reaction with the 1:1 mixture of LiTMP and Zn(TMP)₂ (Table 3). After 2 h contact with the base (0.5 equiv of each metal amide) in THF at room temperature, iodolysis furnished a mixture of the 2- and 4-iodo derivatives (**2b** and **2c**) in 57 and 37% yield, respectively (entry 1). Compared with methoxy (see above) and the other halogens (chlorine, bromine),^{10e} fluorine at C3 is less capable than the former and more than the latter of contributing to the stabilization of a 4-metalated species; thus, its ability to coordinate and stabilize a 4-metalated derivative seems to be intermediate between methoxy and the other halogens.

Increasing the base amount (to 1 equiv of each metal amide) allowed the diiodide **2d** to be obtained in 46% yield together with the 4-iodo **2c** (45% yield) and the 2-iodo **2b** (9% yield). The experiment carried out by using 2 equiv of each metal amide quantitatively provided the diiodo **2d** (entry 3); this result is different to what is observed from **1a**, and can be in relation with lower pK_a values.

Table 3. Calculated $pK_a(\text{THF})$ values for 3-fluoropyridine (**2a**), and deprotometalation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis and ORTEP diagram (30% probability) of **2d**.



Entry	<i>n</i> (equiv)	Products, yields ^a (%)		
1	0.5	2b , 57	2c , 37	2d , 1
2	1	2b , 9	2c , 45	2d , 46
3	2	2b , -	2c , 6	2d , 94

^a Yield after purification by column chromatography.

We next focused on 4- and 2-methoxypyridines (**3a** and **4a**). Concerning 4-methoxypyridine (**3a**), using LiDA in THF in the presence of chlorotrimethylsilane leads to the 3-silylated derivative (61% yield) together with the 3,5-disilylated (16%).²⁰ Deprotolithiation can be carried out at the 3 position more efficiently by using either mesityllithium in THF at -23 °C²⁰ or phenyllithium in THF at 0 °C.³¹

Due to lower LUMO levels, 2-methoxypyridine (**4a**) is more prone to nucleophilic attacks than 3- and 4-methoxypyridines (**1a** and **3a**).³² Using BuLi in THF with **4a** at temperatures between 0 and 20 °C leads to both deprotolithiation and nucleophilic addition.³³ Conducting the reaction by using LiDA and chlorotrimethylsilane as an in situ trap quantitatively furnishes the 3-trimethylsilyl derivative.³² To make efficient this LiDA-promoted deprotometalation, and extend it to other kinds of electrophilic trapping, it is possible to consume the diisopropylamine formed by reaction either with MeLi³² or with PhLi.³⁴ As for 3-methoxypyridine (**1a**), mesityllithium can be used in THF at room temperature.²⁰ Bimetallic combinations such as 1:1 LiTMP-Al(*i*Bu)₃ (“trans-metal trapping”¹² in THF at -78 °C),³⁵ putative LiCo(TMP)₃ (in THF at room temperature),³⁶ LiCu(TMP)₂ (in THF at room temperature)³⁷ and putative LiFe(TMP)₃ (in THF at room temperature),³⁸ can be employed for the same purpose. The regioselectivity of the reaction can be switched from the more acidic 3 to the less acidic 6 position by recourse to BuLi-LiDMAE in hexane at 0 °C, conditions that favor coordination- over acidity-driven reaction.³⁹

As previously for 3-methoxypyridine (**1a**), 4- and 2-methoxypyridines (**3a** and **4a**) can be converted to the 3-iodo derivatives either in 58 and 70% yield, respectively, by using (PMDETA)₂K₂Mg(CH₂SiMe₃)₄ in hexane at 0 °C,²¹ or in 92 and 70% yield by using TMP-zincate in THF at room temperature.²³

Upon treatment by the 1:1 mixture of LiTMP and Zn(TMP)₂ (0.5 equiv of each metal amide) in THF for 2 h at room temperature followed by quenching with iodine, 4-methoxypyridine (**3a**) cleanly provided the 3-iodo derivative **3b**, which was isolated in 89% yield (Table 4, entry 1). When the amount of base was increased to 1 equiv of each metal amide, no more starting material was detected but 3,5-diiodo-4-methoxypyridine **3c** (7% yield), 2,3-diiodo-4-methoxypyridine **3d** (2% yield) and 2,5-diiodo-4-methoxypyridine **3e** (traces) concomitantly formed with the monoiodide **3b** (88% yield) (Table 4, entry 2). It is worth mentioning that, by using LiTMP (1.5 equiv) without Zn(TMP)₂ in THF at 0 °C for 2 h, the yield of **3b** dropped to 38% again demonstrating the value of the “trans-metal trapping”¹² approach.

Table 4. Calculated pK_a(THF) values for 4-methoxypyridine (**3a**), and deprotometalation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis.

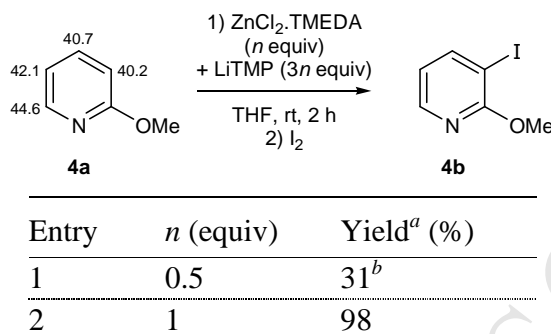
Entry	<i>n</i> (equiv)	Products, yields ^a (%)
1	0.5	3b , 89 3c , - 3d , - 3e , -
2	1	3b , 88 3c , 7 3d , 2 3e , traces

^a Yield after purification by column chromatography.

Under the same reaction conditions, 1:1 LiTMP-Zn(TMP)₂ (0.5 equiv of each metal amide) converted 2-methoxypyridine (**4a**) into the 3-iodo derivative **4b** but in a moderate 31% yield (Table 5, entry 1), a result that could be due to higher pK_a values at the 3 position. A quantitative yield was reached by

doubling the amount of base (Table 5, entry 2). Such a result could not be reproduced by using LiTMP, even in the presence of LiCl.

Table 5. Calculated $pK_a(\text{THF})$ values for 2-methoxypyridine (**4a**), and deprotometalation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis.

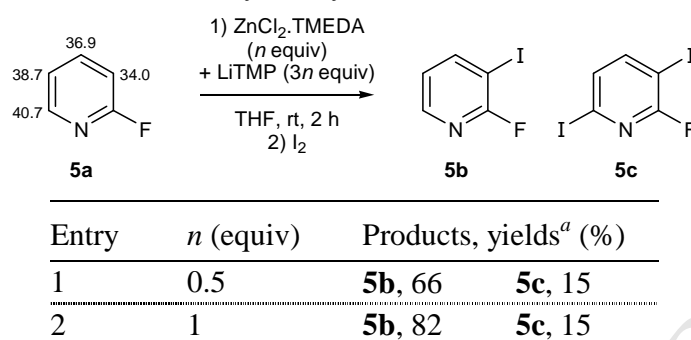


^a Yield after purification by column chromatography. ^b The rest is starting material.

Compared with 2-methoxypyridine (**4a**), 2-fluoropyridine (**5a**) is more reactive. Alkylolithiums already add nucleophilically to the latter at low temperatures, e.g. BuLi·TMEDA in Et₂O at -40 °C and, to a lesser extent, BuLi in Et₂O at -60 °C.⁴⁰ In contrast, using more protophilic LiDA⁴⁰⁻⁴¹ or BuLi·*t*BuOK,²⁹ both in THF at -75 °C, favors proton abstraction (which occurs next to fluorine). Turning to Li₂(TMP)MgBu₃³⁰ and LiCu(TMP)₂³⁷ in THF allows **5a** to be deprotometalated at -10 °C and rt, respectively.

This higher reactivity of 2-fluoropyridine (**5a**) over **4a**, probably in relation with lower pK_a values, was here evidenced by reaction with 1:1 LiTMP-Zn(TMP)₂ to give under the same reaction conditions 2-fluoro-3-iodopyridine (**5b**) as the main product together with 15% of 2-fluoro-3,6-diiodopyridine (**5c**) (Table 6). Once again, the behavior of fluorine is intermediate between those of the methoxy group and the other halogens (chlorine, bromine). Indeed, for the latter, not only 2-halo-3-iodo and 2-halo-3,6-diiodo derivatives are formed, but also 2-halo-6-iodo and 2-halo-4-iodo derivatives.^{10e}

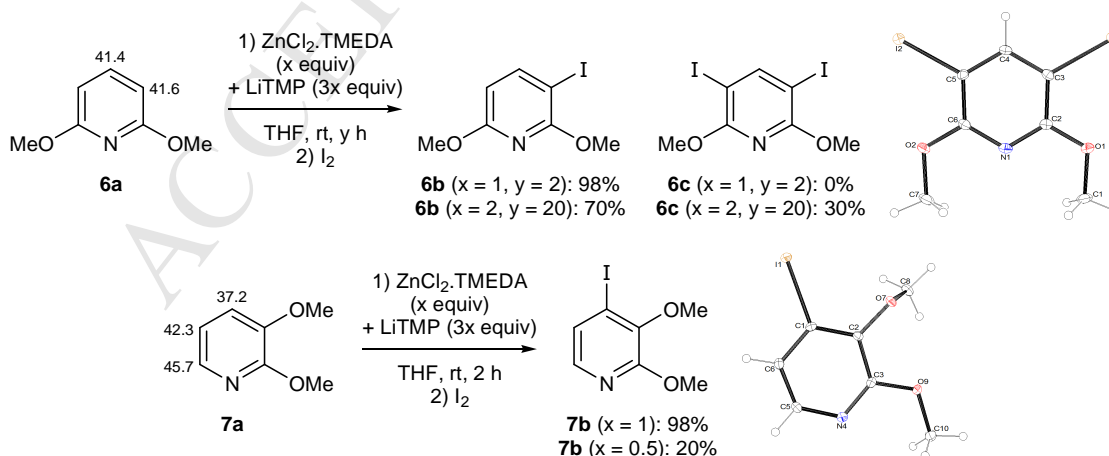
Table 6. Calculated $pK_a(\text{THF})$ values for 2-fluoropyridine (**5a**), and deprotonation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis.



^a Yield after purification by column chromatography.

Compared with 2-methoxypyridine (**4a**), 2,6-dimethoxypyridine (**6a**) is less prone to nucleophilic attacks. BuLi can thus be employed to functionalize efficiently the 3 position.⁴² The bimetallic base LiCo(TMP)₃ can also be used in THF at room temperature for the same purpose.³⁶ In the case of 2,3-dimethoxypyridine (**7a**), 2 equiv of BuLi are required to perform an efficient functionalization at the 4 position.⁴³

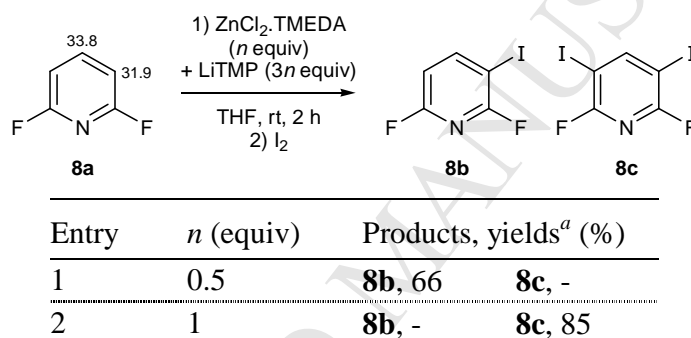
Both compounds **6a** and **7a** could be quantitatively deprotonated next to the methoxy group by using 1:1 LiTMP-Zn(TMP)₂ (1 equiv of each metal amide), a result evidenced by subsequent iodolysis. From **6a**, dimetalation was only observed when the amount of base was doubled and the reaction time extended to 20 h. Reducing the amount of base in the case of **7a** led to a lower 20% yield (Scheme 3).



Scheme 3. Calculated $pK_a(\text{THF})$ values for 2,6-dimethoxypyridine (**6a**) and 2,3-dimethoxypyridine (**7a**), deprotonation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis and ORTEP diagram (30% probability) of **6c** and **7b**.

We finally studied the behavior of 2,6-difluoropyridine (**8a**). **8a** can be cleanly deprotonated at its 3 position upon treatment with LiDA in THF at $-75\text{ }^{\circ}\text{C}$.^{44,41b,c} A similar reaction is also possible by using $\text{Li}_2(\text{TMP})\text{MgBu}_3$ in THF at $-10\text{ }^{\circ}\text{C}$.³⁰ In contrast to what can be observed by employing LiDA, using 1:1 LiTMP-Zn(TMP)₂ in THF at room temperature led either to mono or to dideprotonation (0.5 or 1 equiv of each metal amide, respectively) of 2,6-difluoropyridine (**8a**), as demonstrated by subsequent interception with iodine to furnish either the monoiodide **8b** (66% yield) or the diiodide **8c** (85% yield) (Table 7).

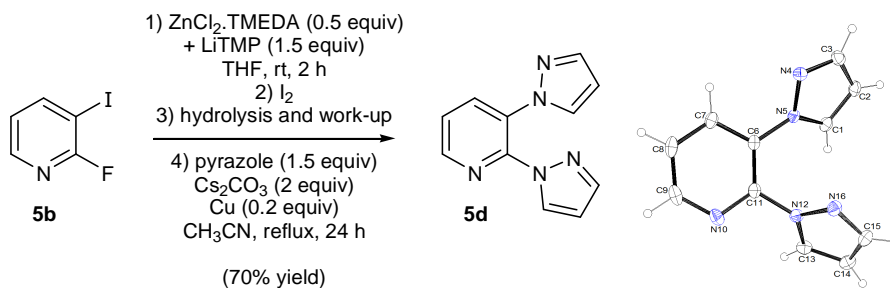
Table 7. Calculated $\text{p}K_{\text{a}}(\text{THF})$ values for 2,6-difluoropyridine (**8a**), and deprotonation using in situ prepared 1:1 LiTMP-Zn(TMP)₂ followed by iodolysis.



^a Yield after purification by column chromatography.

Compared with 2,6-dimethoxypyridine (**6a**), **8a** is more prone to dimetalation, probably in relation with lower $\text{p}K_{\text{a}}$ values. Compared with **8a**, pyridines bearing heavier halogens (chlorine, bromine) give under similar reaction conditions more than one product,^{10e} a result probably resulting from both steric and longer range acidifying effects.⁴⁵

Finally, we have shown that it is possible to associate deprotonation-iodination of pyridines with *N*-arylation of azoles for the generation of C,N'-linked bis-heterocycles. To this purpose, after hydrolysis and work-up, we involved the crude containing the iodide **5b** in the reaction with pyrazole (2 equiv) using metal copper (0.2 equiv) as transition metal, cesium carbonate (2 equiv) as base, and acetonitrile as solvent at its reflux temperature for 24 h.⁴⁶ Under these conditions, both substitution of fluorine and *N*-arylation were noted, as evidenced with the formation of **5d** in 70% yield (Scheme 4).



Scheme 4. Deprotometalation-iodination of **5a** followed by *N*-arylation of pyrazole with the crude iodide **5b** and ORTEP diagram (30% probability) of **5d**.

3. Conclusion

In summary, the basic mixture prepared from 1:3 $\text{ZnCl}_2 \cdot \text{TMEDA}$ -LiTMP, and supposed to be a 1:1 mixture of LiTMP and $\text{Zn}(\text{TMP})_2$, allows reactions that are not reached by its lithium precursor. Besides more efficient monodeprotonations, for example from 2-methoxypyridine, dideprotonations can be efficiently achieved for substrates benefiting from lower $\text{p}K_a$ values such as 3-fluoropyridine and 2,6-difluoropyridine. Such results are in accordance with a “trans-metal trapping”¹² allowing the pyridyllithiums to be trapped by zinc species as they are formed.

4. Experimental

4.1. General

All the reactions were performed in Schlenk tubes under an argon atmosphere. THF was distilled over sodium/benzophenone. Liquid chromatography separations were achieved on silica gel Merck-Geduran Si 60 (63–200 μm). Melting points were measured on a Kofler apparatus. IR spectra were taken on a Perkin–Elmer Spectrum 100 spectrometer. ^1H and ^{13}C Nuclear Magnetic Resonance (NMR) spectra were recorded on a Bruker Avance III spectrometer at 300 and 75 MHz, respectively. ^1H chemical shifts (δ) are given in ppm relative to the solvent residual peak, and ^{13}C chemical shifts are relative to the central peak of the solvent signal.⁴⁷

4.1.1. Crystallography

The sample was studied with graphite monochromatized Mo-K α radiation ($\lambda = 0.71073 \text{ \AA}$). The X-ray diffraction data were collected by using either APEXII, Bruker-AXS diffractometer (compounds **1d**, **2d** and **5d**) or D8 VENTURE Bruker AXS diffractometer (compounds **6c** and **7b**). The structure was solved by direct methods using the SIR97 program,⁴⁸ and then refined with full-matrix least-square methods based on F^2 (SHELX-97)⁴⁹ with the aid of the WINGX program.⁵⁰ All non-hydrogen atoms were refined with anisotropic atomic displacement parameters. H atoms were finally included in their calculated positions. Molecular diagrams were generated by ORTEP-3 (version 2.02).⁵⁰

4.2. General procedure 1: To a stirred, cooled (0 °C) solution of 2,2,6,6-tetramethylpiperidine (0.25 mL, 1.5 mmol) in THF (2-3 mL) were successively added BuLi (about 1.6 M hexanes solution, 1.5 mmol) and, 5 min later, ZnCl₂·TMEDA⁵¹ (0.13 g, 0.50 mmol). The mixture was stirred for 15 min at 0 °C before introduction of the substrate (1.0 mmol) at 0-10 °C. After 2 h at room temperature, a solution of I₂ (0.38 g, 1.5 mmol) in THF (4 mL) was added. The mixture was stirred overnight before addition of an aqueous saturated solution of Na₂S₂O₃ (4 mL) and extraction with AcOEt (3 x 20 mL). The combined organic layers were dried over MgSO₄, filtered and concentrated under reduced pressure. Purification by chromatography on silica gel (the eluent is given in the product description) led to the compounds described below.

4.2.1. 2-Iodo-3-methoxypyridine (1b). The general procedure 1, but with a 10 min contact time between the base and the substrate instead of 2 h, using 3-methoxypyridine (0.10 mL) gave **1b** (eluent: heptane-AcOEt 20:80) in 50% yield as a yellow oil: ¹H NMR (CDCl₃) δ 3.86 (s, 3H), 6.97 (dd, 1H, $J = 8.1$ and 1.5 Hz), 7.17 (dd, 1H, $J = 8.1$ and 4.5 Hz), 7.95 (dd, 1H, $J = 4.5$ and 1.5 Hz); ¹³C NMR (CDCl₃) δ 56.4 (CH₃), 111.6 (C), 116.9 (CH), 123.6 (CH), 142.5 (CH), 155.3 (C). These data are in accordance with those previously described.⁵²

4.2.2. 4-Iodo-3-methoxypyridine (1c). The general procedure 1 using 3-methoxypyridine (0.10 mL) gave **1c** (eluent: heptane-AcOEt 20:80) in 85% yield as a pale yellow powder: mp 88-90 °C; IR (ATR): 720, 815, 1015, 1065, 1196, 1250, 1281, 1403, 1473, 1558 cm⁻¹; ¹H NMR (CDCl₃) δ 3.98 (s, 3H), 7.72

(d, 1H, $J = 5.1$ Hz), 7.87 (d, 1H, $J = 5.1$ Hz), 8.10 (s, 1H). The ^1H NMR data are in accordance with those previously described.²³ ^{13}C NMR (CDCl_3) δ 57.1 (CH_3), 97.4 (C), 132.9 (CH), 134.4 (CH), 143.0 (CH), 155.2 (C).

4.2.3. 3-Fluoro-2-iodopyridine (2b). The general procedure 1 using 3-fluoropyridine (86 μL) gave **2b** (eluent: CH_2Cl_2) in 57% yield as a yellow oil: ^1H NMR (CDCl_3) δ 7.29-7.42 (m, 2H), 8.25 (dt, 1H, $J = 4.5$ and 1.5 Hz). The ^1H NMR data are in accordance with those previously described.⁵³

4.2.4. 3-Iodo-4-methoxypyridine (3b). The general procedure 1 using 4-methoxypyridine (0.10 mL) gave **3b** (eluent: heptane-AcOEt 20:80) in 89% yield as a pale yellow powder: mp 62-64 $^\circ\text{C}$ (lit.^{10h} 64 $^\circ\text{C}$); ^1H NMR (CDCl_3) δ 3.95 (s, 3H), 6.78 (d, 1H, $J = 5.7$ Hz), 8.40 (d, 1H, $J = 5.4$ Hz), 8.75 (s, 1H). The ^1H NMR data are in accordance with those previously reported.^{10h} ^{13}C NMR (CDCl_3) δ 55.7 (CH_3), 84.6 (C), 106.5 (CH), 150.1 (CH), 156.8 (CH), 163.2 (C).

4.2.5. 2-Fluoro-3,6-diiodopyridine (5c). The general procedure 1 using 2-fluoropyridine (86 μL) gave **5c** (eluent: heptane-AcOEt 10:90) in 15% yield as a pale yellow powder: mp 88 $^\circ\text{C}$ (lit.^{10d} 89 $^\circ\text{C}$); IR (ATR): 671, 727, 823, 871, 1010, 1114, 1135, 1226, 1257, 1368, 1416, 1531, 1548, 3508 cm^{-1} ; ^1H NMR (CDCl_3) δ 7.38 (dd, 1H, $J = 7.8$ Hz), 7.75 (t, 1H, $J = 8.0$ Hz). These ^1H NMR data are in accordance with those previously described.^{10d}

4.2.6. 2,6-Difluoro-3-iodopyridine (8b). The general procedure 1 using 2,6-difluoropyridine (91 μL) gave **8b** (eluent: heptane-AcOEt 10:90) in 66% yield as a white powder: mp < 50 $^\circ\text{C}$; ^1H NMR (CDCl_3) δ 6.69 (ddd, 1H, $J = 8.4$, 3.0 and 0.9 Hz), 8.19 (td, 1H, $J = 8.1$ and 7.8 Hz); ^{13}C NMR (CDCl_3) δ 69.3 (dd, C, $J = 40$ and 5.9 Hz), 108.4 (dd, CH, $J = 35$ and 5.9 Hz), 153.6 (dd, CH, $J = 7.4$ Hz), 160.3 (dd, C, $J = 241$ and 14 Hz), 162.0 (dd, C, $J = 247$ and 13 Hz). These data are in accordance with those previously reported.³⁰

4.3. General procedure 2: To a stirred, cooled (0 $^\circ\text{C}$) solution of 2,2,6,6-tetramethylpiperidine (0.50 mL, 3.0 mmol) in THF (5 mL) were successively added BuLi (about 1.6 M hexanes solution, 3.0 mmol) and, 5 min later, $\text{ZnCl}_2 \cdot \text{TMEDA}$ ⁵¹ (0.26 g, 1.0 mmol). The mixture was stirred for 15 min at 0

°C before introduction of the substrate (0.5 mmol) at 0-10 °C. After 2 h at room temperature, a solution of I₂ (0.76 g, 3.0 mmol) in THF (8 mL) was added. The mixture was stirred overnight before addition of an aqueous saturated solution of Na₂S₂O₃ (8 mL) and extraction with AcOEt (3 x 30 mL). The combined organic layers were dried over MgSO₄, filtered and concentrated under reduced pressure. Purification by chromatography on silica gel (the eluent is given in the product description) led to the compounds described below.

4.3.1. 2,4-Diiodo-3-methoxypyridine (1d).^{10h} The general procedure 2 using 3-methoxypyridine (50 μL) gave **1d** (eluent: heptane-AcOEt 20:80) in 35% yield as a pale yellow powder: mp 146-148 °C; IR (ATR): 698, 810, 840, 919, 978, 1012, 1185, 1247, 1358, 1452, 1523, 1540, 1717, 2931, 3345 cm⁻¹; ¹H NMR (CDCl₃) δ 3.92 (s, 3H), 7.66 (d, 1H, *J* = 5.1 Hz), 7.74 (d, 1H, *J* = 5.1 Hz); ¹³C NMR (CDCl₃) δ 61.2 (CH₃), 100.9 (C), 115.2 (C), 134.3 (CH), 146.8 (CH), 157.1 (C). **Crystal data for 1d** (CCDC 1450720, room temperature): C₆H₅I₂NO, *M* = 360.91, orthorhombic, *P b c a*, *a* = 13.2964(7), *b* = 8.6329(3), *c* = 15.5558(7) Å, *V* = 1785.60(14) Å³, *Z* = 8, *d* = 2.685 g cm⁻³, *μ* = 6.982 mm⁻¹. A final refinement on *F*² with 2035 unique intensities and 93 parameters converged at $\omega R(F^2) = 0.0730$ (*R*(*F*) = 0.0311) for 1679 observed reflections with *I* > 2σ(*I*).

4.3.2. 3-Fluoro-2,4-diiodopyridine (2d). The general procedure 2 using 3-fluoropyridine (43 μL) gave **2d** (eluent: heptane-AcOEt 90:10) in 94% yield as a yellow powder: mp 102 °C (lit.⁵⁴ 102 °C); ¹H NMR (CDCl₃) δ 7.63 (dd, 1H, *J* = 4.8 and 4.2 Hz), 7.82 (dd, 1H, *J* = 5.1 and 0.9 Hz); ¹³C NMR (CDCl₃) δ 91.1 (d, C, *J* = 17 Hz), 105.4 (d, C, *J* = 32 Hz), 134.0 (d, CH, *J* = 1.1 Hz), 146.9 (d, CH, *J* = 6.2 Hz), 158.6 (d, C, *J* = 254 Hz). **Crystal data for 2d** (CCDC 1450721, T = 150 K): C₅H₂FI₂N, *M* = 348.88, triclinic, *P -1*, *a* = 7.1039(4), *b* = 7.8955(5), *c* = 7.9442(5) Å, *α* = 109.945(2), *β* = 105.973(2), *γ* = 102.495(2) °, *V* = 378.21(4) Å³, *Z* = 2, *d* = 3.063 g cm⁻³, *μ* = 8.244 mm⁻¹. A final refinement on *F*² with 1737 unique intensities and 83 parameters converged at $\omega R(F^2) = 0.1171$ (*R*(*F*) = 0.0389) for 1639 observed reflections with *I* > 2σ(*I*).

4.3.3. 3,5-Diiodo-2,6-dimethoxypyridine (6c). The general procedure 2, but with a contact time of 20 h with the base, using 2,6-dimethoxypyridine (66 μL) gave **6c** (eluent: heptane- CH_2Cl_2 60:40) in 30% yield as a pale yellow powder: mp 126 $^\circ\text{C}$; IR (ATR): 708, 736, 908, 999, 1038, 1235, 1247, 1288, 1315, 1385, 1361, 1459, 1552, 2948 cm^{-1} ; ^1H NMR (CDCl_3) δ 3.96 (s, 6H), 8.16 (s, 1H); ^{13}C NMR (CDCl_3) δ 54.9 (2 CH_3), 66.8 (2C), 156.8 (CH), 161.3 (2C). **Crystal data for 6c** (CCDC 1450723, T = 150 K): $\text{C}_7\text{H}_7\text{I}_2\text{NO}_2$, $M = 390.94$, monoclinic, $P 2_1/c$, $a = 8.2579(3)$, $b = 8.6432(3)$, $c = 14.7789(6)$ \AA , $\beta = 104.7150(10)^\circ$, $V = 1020.24(7)$ \AA^3 , $Z = 4$, $d = 2.545$ g cm^{-3} , $\mu = 6.128$ mm^{-1} . A final refinement on F^2 with 2337 unique intensities and 112 parameters converged at $\omega R(F^2) = 0.0433$ ($R(F) = 0.0174$) for 2190 observed reflections with $I > 2\sigma(I)$.

4.4. General procedure 3: To a stirred, cooled (0 $^\circ\text{C}$) solution of 2,2,6,6-tetramethylpiperidine (0.25 mL, 1.5 mmol) in THF (2-3 mL) were successively added BuLi (about 1.6 M hexanes solution, 1.5 mmol) and, 5 min later, $\text{ZnCl}_2 \cdot \text{TMEDA}^{51}$ (0.13 g, 0.50 mmol). The mixture was stirred for 15 min at 0 $^\circ\text{C}$ before introduction of the substrate (0.5 mmol) at 0-10 $^\circ\text{C}$. After 2 h at room temperature, a solution of I_2 (0.38 g, 1.5 mmol) in THF (4 mL) was added. The mixture was stirred overnight before addition of an aqueous saturated solution of $\text{Na}_2\text{S}_2\text{O}_3$ (4 mL) and extraction with AcOEt (3 x 20 mL). The combined organic layers were dried over MgSO_4 , filtered and concentrated under reduced pressure. Purification by chromatography on silica gel (the eluent is given in the product description) led to the compounds described below.

4.4.1. 3-Fluoro-4-iodopyridine (2c). The general procedure 3 using 3-fluoropyridine (43 μL) gave **2c** (eluent: hexane- CH_2Cl_2 80:20) in 45% yield as a yellow powder: mp 94 $^\circ\text{C}$ (lit.⁵³ 96 $^\circ\text{C}$); IR (ATR): 1415, 1475, 1550, 1570, 3060 cm^{-1} ; ^1H NMR (CDCl_3) δ 7.75 (t, 1H, $J = 5.1$ Hz), 8.10 (d, 1H, $J = 5.1$ Hz), 8.35 (s, 1H); ^{13}C NMR (CDCl_3) δ 92.4 (d, C, $J = 23$ Hz), 133.7 (s, CH), 137.1 (d, CH, $J = 26$ Hz), 145.6 (d, CH, $J = 5.1$ Hz), 159.1 (d, C, $J = 256$ Hz).

4.4.2. 3,5-Diiodo-4-methoxypyridine (3c). The general procedure 3 using 4-methoxypyridine (51 μL) gave **3c** (eluent: heptane-AcOEt 20:80) in an estimated 7% yield. It was identified by NMR: ^1H NMR (CDCl_3) δ 3.91 (s, 3H), 8.72 (s, 2H).

4.4.3. 2,3-Diiodo-4-methoxypyridine (3d). The general procedure 3 using 4-methoxypyridine (51 μL) gave **3d** (eluent: heptane-AcOEt 20:80) in an estimated 2% yield. It was identified by NMR: ^1H NMR (CDCl_3) δ 3.93 (s, 3H), 6.68 (d, 1H, $J = 5.4$ Hz), 8.20 (d, 1H, $J = 5.4$ Hz).

4.4.4. 2,5-Diiodo-4-methoxypyridine (3e). The general procedure 3 using 4-methoxypyridine (51 μL) gave **3e** (eluent: heptane-AcOEt 20:80) as traces. It was identified by NMR: ^1H NMR (CDCl_3) δ 3.91 (s, 3H), 7.11 (s, 1H), 8.44 (s, 1H).

4.4.5. 3-Iodo-2-methoxypyridine (4b). The general procedure 3 using 2-methoxypyridine (53 μL) gave **4b** (eluent: heptane-AcOEt 80:20) in 98% yield as a yellow oil: ^1H NMR (CDCl_3) δ 3.95 (s, 3H), 6.60 (dd, 1H, $J = 7.5$ and 4.8 Hz), 7.98 (dd, 1H, $J = 7.5$ and 1.8 Hz), 8.08 (dd, 1H, $J = 5.1$ and 1.8 Hz); ^{13}C NMR (CDCl_3) δ 54.7 (CH_3), 79.9 (C), 118.3 (CH), 146.5 (CH), 148.0 (CH), 161.9 (C). These NMR data are in accordance with those previously described.³⁶

4.4.6. 2-Fluoro-3-iodopyridine (5b). The general procedure 3 using 2-fluoropyridine (43 μL) gave **5b** (eluent: heptane-AcOEt 10:90) in 82% yield as a white powder: mp < 50 $^\circ\text{C}$; IR (ATR): 737, 794, 843, 1018, 1065, 1137, 1253, 1370, 1407, 1424, 1557, 1576, 3383 cm^{-1} ; ^1H NMR (CDCl_3) δ 6.93 (ddd, 1H, $J = 7.5$, 5.1 and 2.1 Hz), 8.09-8.16 (m, 2H). These data are in accordance with those described previously.^{10d}

4.4.7. 3-Iodo-2,6-dimethoxypyridine (6b). The general procedure 3 using 3-methoxypyridine (50 μL) gave **6b** (eluent: heptane-AcOEt 10:90) in 98% yield as a yellow oil: IR (ATR): 670, 807, 950, 1002, 1020, 1050, 1112, 1191, 1234, 1265, 1309, 1372, 1411, 1462, 1567, 2948, 2985 cm^{-1} ; ^1H NMR (CDCl_3) δ 3.89 (s, 3H), 3.96 (s, 3H), 6.15 (d, 1H, $J = 8.1$ Hz), 7.80 (d, 1H, $J = 8.1$ Hz); ^{13}C NMR (CDCl_3) δ 53.5 (CH_3), 54.3 (CH_3), 65.5 (C), 103.4 (CH), 149.3 (CH), 160.5 (C), 163.1 (C). These data are in accordance with those previously described.³⁶

4.4.8. 4-Iodo-2,3-dimethoxypyridine (7b). The general procedure 3 using 3-methoxypyridine (50 μL) gave **7b** (eluent: heptane-AcOEt 10:90) in 98% yield as a pale yellow powder: mp < 50 $^{\circ}\text{C}$; IR (ATR): 659, 769, 819, 855, 986, 1020, 1157, 1215, 1383, 1460, 1562, 2938 cm^{-1} ; ^1H NMR (CDCl_3) δ 3.83 (s, 3H), 3.96 (s, 3H), 7.22 (d, 1H, $J = 5.4$ Hz), 7.51 (d, 1H, $J = 5.4$ Hz); ^{13}C NMR (CDCl_3) δ 53.7 (CH_3), 60.0 (CH_3), 101.7 (C), 126.6 (CH), 141.4 (CH), 144.3 (C), 156.8 (C). **Crystal data for 7b** (CCDC 1450724, T = 150 K): $\text{C}_7\text{H}_8\text{INO}_2$, $M = 265.04$, orthorhombic, Pbc_a , $a = 8.0585(4)$, $b = 13.4414(6)$, $c = 15.7058(7)$ \AA , $V = 1701.21(14)$ \AA^3 , $Z = 8$, $d = 2.070$ g cm^{-3} , $\mu = 3.715$ mm^{-1} . A final refinement on F^2 with 1957 unique intensities and 102 parameters converged at $\omega R(F^2) = 0.0408$ ($R(F) = 0.0177$) for 1730 observed reflections with $I > 2\sigma(I)$.

4.4.9. 2,6-Difluoro-3,5-diiodopyridine (8c). The general procedure 3 using 2,6-difluoropyridine (45 μL) gave **8c** (eluent: heptane-AcOEt 10:90) in 85% yield as a pale yellow powder: mp 102 $^{\circ}\text{C}$; IR (ATR): 666, 728, 811, 1047, 1268, 1363, 1423, 1567, 2922 cm^{-1} ; ^1H NMR (CDCl_3) δ 8.50 (t, 1H, $J = 7.7$ Hz); ^{13}C NMR (CDCl_3) δ 70.7-71.3 (m, 2C), 160.6 (t, CH, $J = 2.8$ Hz), 160.7 (dd, 2C, $J = 243$ and 13 Hz).

4.5. General procedure 4: To a stirred, cooled (0 $^{\circ}\text{C}$) solution of 2,2,6,6-tetramethylpiperidine (0.25 mL, 1.5 mmol) in THF (2-3 mL) were successively added BuLi (about 1.6 M hexanes solution, 1.5 mmol) and, 5 min later, $\text{ZnCl}_2 \cdot \text{TMEDA}^{51}$ (0.13 g, 0.50 mmol). The mixture was stirred for 15 min at 0 $^{\circ}\text{C}$ before introduction of the substrate (1.0 mmol) at 0 - 10 $^{\circ}\text{C}$. After 2 h at room temperature, a solution of I_2 (0.38 g, 1.5 mmol) in THF (4 mL) was added. The mixture was stirred overnight before addition of an aqueous saturated solution of $\text{Na}_2\text{S}_2\text{O}_3$ (4 mL) and extraction with AcOEt (3 x 20 mL). The combined organic layers were dried over MgSO_4 , filtered and concentrated under reduced pressure. To the crude iodide were added Cs_2CO_3 (0.65 g, 2.0 mmol), Cu powder (13 mg, 0.20 mmol), the azole (1.5 mmol) and MeCN (5 mL) and the resulting mixture was heated under reflux for 24 h. Filtration over celite[®], washing with AcOEt, removal of the solvent and purification by chromatography on silica gel (the eluent is given in the product description) led to the compound described below.

4.5.1. 2,3-Bis(1H-1-pyrazolyl)pyridine (5d). The general procedure 4 using 2-fluoropyridine (86 μL) and pyrazole (0.10 mL) gave **5d** (eluent: heptane-AcOEt 90:10) in 70% yield as a yellow powder: mp 98 $^{\circ}\text{C}$; IR (ATR): 757, 808, 937, 1018, 1036, 1050, 1262, 1297, 1393, 1458, 1479, 1521, 1584, 3090, 3122 cm^{-1} ; ^1H NMR (CDCl_3) δ 6.29 (t, 1H, $J = 2.3$ Hz), 6.34 (t, 1H, $J = 2.3$ Hz), 7.06 (d, 1H, $J = 2.4$ Hz), 7.40 (dd, 1H, $J = 8.1$ and 4.8 Hz), 7.61-7.67 (m, 3H), 8.05 (dd, 1H, $J = 7.8$ and 1.5 Hz), 8.50 (dd, 1H, $J = 4.8$ and 1.8 Hz); ^{13}C NMR (CDCl_3) δ 107.6 (CH), 107.7 (CH), 123.6 (CH), 129.8 (CH), 130.2 (C), 130.6 (CH), 136.2 (CH), 141.5 (CH), 142.0 (CH), 144.9 (C), 147.8 (CH). **Crystal data for 5d** (CCDC 1450722, T = 150 K): $\text{C}_{11}\text{H}_9\text{N}_5$, $M = 211.23$, orthorhombic, $P 2_1 2_1 2_1$, $a = 8.7934(9)$, $b = 8.9384(7)$, $c = 13.0819(14)$ \AA , $V = 1028.22(17)$ \AA^3 , $Z = 4$, $d = 1.365$ g cm^{-3} , $\mu = 0.090$ mm^{-1} . A final refinement on F^2 with 1369 unique intensities and 145 parameters converged at $\omega R(F^2) = 0.0937$ ($R(F) = 0.0392$) for 1211 observed reflections with $I > 2\sigma(I)$.

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References and Notes

1. Eicher, T.; Hauptmann, S.; Speicher, A., *The Chemistry of Heterocycles*, 2nd ed., Wiley-VCH, New York, 2003.
2. See for example: (a) Kim, M.; Lee, J. Y. *Chem. Asian J.* **2012**, *7*, 899-902; (b) Lin, H.-W.; Lu, C.-W.; Lin, L.-Y.; Chen, Y.-H.; Lin, W.-C.; Wong, K.-T.; Lin, F. *J. Mater. Chem. A* **2013**, *1*, 1770-1777.
3. (a) Gschwend, H. W.; Rodriguez, H. R. *Org. React.* **1979**, *26*, 1-360; (b) Beak, P.; Snieckus, V. *Acc. Chem. Res.* **1982**, *15*, 306-312; (c) Snieckus, V. *Chem. Rev.* **1990**, *90*, 879-933; (d) Gant, T. G.; Meyers, A. I. *Tetrahedron* **1994**, *50*, 2297-2360; (e) Schlosser, M. *Organometallics in Synthesis* **2002**, 2nd ed. (Ed.: M. Schlosser), Wiley: New York, Chapter I.

4. (a) Queguiner, G.; Marsais, F.; Snieckus, V.; Epszajn, J. *Adv. Heterocycl. Chem.* **1991**, *52*, 187-304; (b) Mongin, F.; Queguiner, G. *Tetrahedron* **2001**, *57*, 4059-4090; (c) Schlosser, M.; Mongin, F. *Chem. Soc. Rev.* **2007**, *36*, 1161-1172.
5. (a) Mulvey, R. E. *Acc. Chem. Res.* **2009**, *42*, 743-755; (b) Haag, B.; Mosrin, M.; Ila, H.; Malakhov, V.; Knochel, P. *Angew. Chem. Int. Ed.* **2011**, *50*, 9794-9824; (c) Mongin, F.; Harrison-Marchand, A. *Chem. Rev.* **2013**, *113*, 7563-7727.
6. (a) Gros, P.; Fort, Y. *Eur. J. Org. Chem.* **2002**, 3375-3383; (b) Gros, P. C.; Fort, Y. *Eur. J. Org. Chem.* **2009**, 4199-4209.
7. Harford, P. J.; Peel, A. J.; Chevallier, F.; Takita, R.; Mongin, F.; Uchiyama, M.; Wheatley, A. E. H. *Dalton Trans.* **2014**, *43*, 14181-14203.
8. (a) Mulvey, R. E.; Mongin, F.; Uchiyama, M.; Kondo, Y. *Angew. Chem. Int. Ed.* **2007**, *46*, 3802-3824; (b) Wunderlich, S. H.; Knochel, P. *Angew. Chem. Int. Ed.* **2007**, *46*, 7685-7688; (c) Jaric, M.; Haag, B. A.; Unsinn, A.; Karaghiosoff, K.; Knochel, P. *Angew. Chem. Int. Ed.* **2010**, *49*, 5451-5455.
9. Mongin, F.; Uchiyama, M. *Curr. Org. Chem.* **2011**, *15*, 2340-2361.
10. (a) L'Helgoual'ch, J. M.; Seggio, A.; Chevallier, F.; Yonehara, M.; Jeanneau, E.; Uchiyama, M.; Mongin, F. *J. Org. Chem.* **2008**, *73*, 177-183; (b) Seggio, A.; Lannou, M.-I.; Chevallier, F.; Nobuto, D.; Uchiyama, M.; Golhen, S.; Roisnel, T.; Mongin, F. *Chem. Eur. J.* **2007**, *13*, 9982-9989; (c) Seggio, A.; Chevallier, F.; Vaultier, M.; Mongin, F. *J. Org. Chem.* **2007**, *72*, 6602-6605; (d) Snégaroff, K.; Komagawa, S.; Chevallier, F.; Gros, P. C.; Golhen, S.; Roisnel, T.; Uchiyama, M.; Mongin, F. *Chem. Eur. J.* **2010**, *16*, 8191-8201; (e) Snégaroff, K.; Nguyen, T. T.; Marquise, N.; Halauko, Y. S.; Harford, P. J.; Roisnel, T.; Matulis, V. E.; Ivashkevich, O. A.; Chevallier, F.; Wheatley, A. E. H.; Gros, P. C.; Mongin, F. *Chem. Eur. J.* **2011**, *17*, 13284-13297; (f) Chevallier, F.; Halauko, Y. S.; Pecceu, C.; Nassar, I. F.; Dam, T. U.; Roisnel, T.; Matulis, V. E.; Ivashkevich, O. A.; Mongin, F. *Org. Biomol. Chem.* **2011**, *9*, 4671-4684; (g) Chevallier, F.; Blin, T.; Nagaradja, E.; Lassagne, F.; Roisnel, T.; Halauko, Y. S.; Matulis, V. E.; Ivashkevich, O. A.; Mongin, F. *Org. Biomol. Chem.* **2012**, *10*, 4878-4885; (h) Kadiyala, R. R.; Tilly, D.; Nagaradja, E.; Roisnel, T.; Matulis, V. E.; Ivashkevich, O. A.; Halauko, Y. S.; Chevallier, F.; Gros, P. C.; Mongin, F. *Chem. Eur. J.* **2013**, *19*, 7944-7960; (i) Nagaradja, E.; Chevallier, F.; Roisnel, T.; Dorcet, V.; Halauko, Y. S.; Ivashkevich, O. A.; Matulis, V. E.; Mongin, F. *Org. Biomol. Chem.* **2014**, 1475-1487; (j) Marquise, N.; Bretel, G.; Lassagne, F.; Chevallier, F.; Roisnel, T.; Dorcet, V.; Halauko, Y. S.; Ivashkevich, O. A.; Matulis, V. E.; Gros, P. C.; Mongin, F. *RSC Adv.* **2014**, *4*, 19602-19612; (k) Ameer Messaoud, M. Y.; Bentabed-Ababsa, G.; Hedidi, M.; Derdour, A.; Chevallier, F.; Halauko, Y. S.; Ivashkevich, O. A.; Matulis, V. E.; Picot, L.; Thiéry, V.; Roisnel, T.; Dorcet, V.; Mongin, F. *Beilstein J. Org. Chem.* **2015**, *11*, 1475-1485; (l) Nagaradja, E.; Bentabed-Ababsa, G.; Scalabrini, M.; Chevallier, F.; Philippot, S.; Fontanay, S.; Duval, R. E.; Halauko, Y. S.; Ivashkevich, O. A.; Matulis, V. E.; Roisnel, T.; Mongin, F. *Bioorg. Med. Chem.* **2015**, *23*, 6355-6363.
11. García-Álvarez, P.; Mulvey, R. E.; Parkinson, J. A. *Angew. Chem. Int. Ed.* **2011**, *50*, 9668-9671.
12. Fuentes, M. A.; Kennedy, A. R.; Mulvey, R. E.; Parkinson, J. A.; Rantanen, T.; Robertson, S. D.; Snieckus, V. *Chem. Eur. J.* **2015**, *21*, 14812-14822.
13. Becker, M. R.; Knochel, P. *Angew. Chem. Int. Ed.* **2015**, *54*, 12501-12505.
14. Concerning simulation of heteroaromatic pK_a values in DMSO, see: Shen, K.; Fu, Y.; Li, J.-N.; Liu, L.; Guo, Q.-X. *Tetrahedron* **2007**, *63*, 1568-1576.
15. Baboul, A. G.; Curtiss, L. A.; Redfern, P. C.; Raghavachari, K. *J. Chem. Phys.* **1999**, *110*, 7650-7657.
16. Fraser, R. R.; Bresse, M.; Mansour, T. S. *J. Am. Chem. Soc.* **1983**, *105*, 7790-7791.
17. Hansch, C.; Leo, A.; Taft, R. W. *Chem. Rev.* **1991**, *91*, 165-195.

18. (a) Whisler, M. C.; MacNeil, S.; Snieckus, V.; Beak, P. *Angew. Chem. Int. Ed.* **2004**, *43*, 2206-2225. For examples in the pyridine series, see: (b) Dhau, J. S.; Singh, A.; Kasetti, Y.; Bhatia, S.; Bharatam, P. V.; Brandão, P.; Félix, V.; Singh, K. N. *Tetrahedron* **2013**, *69*, 10284-10291.
19. Marsais, F.; Le Nard, G.; Queguiner, G. *Synthesis* **1982**, 235-237.
20. Comins, D. L.; LaMunyon, D. H. *Tetrahedron Lett.* **1988**, *29*, 773-776.
21. Baillie, S. E.; Bluemke, T. D.; Clegg, W.; Kennedy, A. R.; Klett, J.; Russo, L.; de Tullio, M.; Hevia, E. *Chem. Commun.* **2014**, *50*, 12859-12862.
22. Kondo, Y.; Shilai, M.; Uchiyama, M.; Sakamoto, T. *J. Am. Chem. Soc.* **1999**, *121*, 3539-3540.
23. Blair, V. L.; Blakemore, D. C.; Hay, D.; Hevia, E.; Pryde, D. C. *Tetrahedron Lett.* **2011**, *52*, 4590-4594.
24. Its ¹H NMR spectrum proved identical to that previously described: Dehmlow, E. V.; Slegers, A. *Liebigs Ann. Chem.* **1992**, 953-959.
25. (a) Uchiyama, M.; Matsumoto, Y.; Nobuto, D.; Furuyama, T.; Yamaguchi, K.; Morokuma, K. *J. Am. Chem. Soc.* **2006**, *128*, 8748-8750; (b) Kondo, Y.; Morey, J. V.; Morgan, J. C.; Naka, H.; Nobuto, D.; Raithby, P. R.; Uchiyama, M.; Wheatley, A. E. H. *J. Am. Chem. Soc.* **2007**, *129*, 12734-12738.
26. Clegg, W.; Conway, B.; Hevia, E.; McCall, M. D.; Russo, L.; Mulvey, R. E. *J. Am. Chem. Soc.* **2009**, *131*, 2375-2384.
27. (a) Marsais, F.; Mallet, M.; Queguiner, G.; Pastour, P. C. *R. Hebd. Séances Acad. Sci., Ser. C* **1972**, *275*, 1535-1538; (b) Marsais, F.; Queguiner, G. *Tetrahedron* **1983**, *39*, 2009-2021.
28. Gribble, G. W.; Saulnier, M. G. *Tetrahedron Lett.* **1980**, *21*, 4137-4140.
29. Shi, G.-q.; Takagishi, S.; Schlosser, M. *Tetrahedron* **1994**, *50*, 1129-1134.
30. Awad, H.; Mongin, F.; Trécourt, F.; Queguiner, G.; Marsais, F.; Blanco, F.; Abarca, B.; Ballesteros, R. *Tetrahedron Lett.* **2004**, *45*, 6697-6701.
31. Trécourt, F.; Mallet, M.; Mongin, O.; Gervais, B.; Queguiner, G. *Tetrahedron* **1993**, *49*, 8373-8380.
32. Trécourt, F.; Mallet, M.; Marsais, F.; Queguiner, G. *J. Org. Chem.* **1988**, *53*, 1367-1371.
33. Thomas, E. W. *J. Org. Chem.* **1986**, *51*, 2184-2191.
34. Mallet, M. *J. Organomet. Chem.* **1991**, *406*, 49-56.
35. (a) Naka, H.; Uchiyama, M.; Matsumoto, Y.; Wheatley, A. E. H.; McPartlin, M.; Morey, J. V.; Kondo, Y. *J. Am. Chem. Soc.* **2007**, *129*, 1921-1930. Concerning the behavior of this lithium-aluminum base, see: (b) Armstrong, D. R.; Crosbie, E.; Hevia, E.; Mulvey, R. E.; Ramsay, D. L.; Robertson, S. D. *Chem. Sci.* **2014**, *5*, 3031-3045.
36. Dayaker, G.; Chevallier, F.; Gros, P. C.; Mongin, F. *Tetrahedron* **2010**, *66*, 8904-8910.
37. Nguyen, T. T.; Marquise, N.; Chevallier, F.; Mongin, F. *Chem. Eur. J.* **2011**, *17*, 10405-10416.
38. Nagaradja, E.; Chevallier, F.; Roisnel, T.; Jouikov, V.; Mongin, F. *Tetrahedron* **2012**, *68*, 3063-3073.
39. Gros, P.; Fort, Y.; Caubère, P. *J. Chem. Soc., Perkin Trans. 1* **1997**, 3071-3080.
40. Marsais, F.; Granger, P.; Queguiner, G. *J. Org. Chem.* **1981**, *46*, 4494-4497.
41. (a) Güngör, T.; Marsais, F.; Queguiner, G. *J. Organomet. Chem.* **1981**, *215*, 139-150; (b) Gupta, L.; Hoepker, A. C.; Singh, K. J.; Collum, D. B. *J. Org. Chem.* **2009**, *74*, 2231-2233; (c) Gupta, L.; Hoepker, A. C.; Ma, Y.; Viciu, M. S.; Faggini, M. F.; Collum, D. B. *J. Org. Chem.* **2013**, *78*, 4214-4230.
42. Khanapure, S. P.; Biehl, E. R. *Heterocycles* **1990**, *31*, 505-516.
43. Trécourt, F.; Mallet, M.; Mongin, O.; Queguiner, G. *J. Org. Chem.* **1994**, *59*, 6173-6178.
44. (a) Schlosser, M.; Rausis, T. *Eur. J. Org. Chem.* **2004**, 1018-1024; (b) Shutske, G. M.; Roehr, J. E. *J. Heterocycl. Chem.* **1997**, *34*, 789-795.
45. (a) Mongin, F.; Curty, C.; Marzi, E.; Leroux, F. R.; Schlosser, M. *ARKIVOC* **2015**, 48-65; (b) Mongin, F. *Chimia* **2016**, *70*, 48-52.

46. These conditions were used with success in: (a) Zhu, R.; Xing, L.; Wang, X.; Cheng, C.; Su, D.; Hu, Y. *Adv. Synth. Catal.* **2008**, *350*, 1253-1257. See also: (b) Hedidi, M.; Bentabed-Ababsa, G.; Derdour, A.; Roisnel, T.; Dorcet, V.; Chevallier, F.; Picot, L.; Thiéry, V.; Mongin, F. *Bioorg. Med. Chem.* **2014**, *22*, 3498-3507.
47. Gottlieb, H. E.; Kotlyar, V.; Nudelman, A. *J. Org. Chem.* **1997**, *62*, 7512-7515.
48. Altomare, A.; Burla, M. C.; Camalli, M.; Cascarano, G. L.; Giacovazzo, C.; Guagliardi, A.; Moliterni, A. G. G.; Polidori, G.; Spagna, R. *J. Appl. Crystallogr.* **1999**, *32*, 115-119.
49. Sheldrick, G. M. *Acta Crystallogr., Sect. A* **2008**, *A64*, 112-122.
50. Farrugia, L. J. *J. Appl. Crystallogr.* **2012**, *45*, 849-854.
51. Isobe, M.; Kondo, S.; Nagasawa, N.; Goto, T. *Chem. Lett.* **1977**, 679-682.
52. Mongin, F.; Trécourt, F.; Mongin, O.; Queguiner, G. *Tetrahedron* **2002**, *58*, 309-314.
53. Rocca, P.; Marsais, F.; Godard, A.; Queguiner, G. *Tetrahedron* **1993**, *49*, 49-64.
54. Rocca, P.; Cochenec, C.; Marsais, F.; Thomas-dit-Dumont, L.; Mallet, M.; Godard, A.; Queguiner, G. *J. Org. Chem.* **1993**, *58*, 7832-7838.