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Use of Single Metal Fragments for Cluster Building: Synthesis, Structure and Bonding of Heterometalla-boranes

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ABSTRACT

Synergic property of the CO ligand in general can stabilize metal complexes at lower oxidation states. Utilizing this feature of CO ligand, we have recently isolated and structurally characterized a highly fluxional molybdenum complex [{Cp*Mo(CO)$_2$}$_2${μ-η^2:η^2-B_2H_4}] (Cp* = η^5-C_5Me_5) (**2**) comprising diborane(4) ligand. Compound **2** represents a rare class of bimetallic diborane(4) complex corresponding to a singly bridged *Cs* structure. In an attempt to isolate the tungsten analogue of **2**, [{Cp*W(CO)$_2$}$_2${μ-η^2:η^2-B_2H_4}], we have isolated an rare vertex-fused cluster [(Cp*W)$_3$WB_9H_{18}] (**5**). Having a structural alikeness with dimolybdenum alkyne complex [{CpMo(CO)$_2$}$_2$$C_2H_2$], we have further explored the chemistry of **2** with CO gas that yielded homoleptic tri-molybdenum complex, [(Cp*Mo)$_3$(μ-H)$_2$(μ_3-H)(μ-CO)$_2$$B_4H_4$] (**4**). In an attempt to replace the 16 electron {Cp*MoH(CO)$_2$} moiety in **4** with isolobal fragment {W(CO)$_5$}, we treated the intermediate, obtained from the reaction of Cp*$MoCl_4$ and $LiBH_4$, with mono metal carbonyl fragment {W(CO)$_5$.thf}. The reaction indeed yielded two bimetallic clusters [(Cp*Mo)$_2$$B_4H_8$W(CO)$_4$], **7** and [(Cp*Mo)$_2$$B_4H_6$W(CO)$_5$], **8** that seem to have generated by the replacement of one {BH} or {BH_3} vertex from [(Cp*Mo)$_2$$B_5H_9$] respectively. All of the compounds have been characterized by various spectroscopic analyses and single crystal X-ray diffraction studies. Electron counting rules and molecular orbital analyses provided further insight into the electronic structure of all these molecules.

INTRODUCTION

Boron hydrides i.e., boranes are an exceptionally diverse class[1], which are of vast importance to many areas of chemistry, both from fundamental and application standpoint.[2-4] Several research groups are actively involved on this, especially diboranes and its derivatives, primarily due to its mimic of alkane counterparts.[5] The results not only enhance our understanding of chemical

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bonding[6-7] of B_nH_{n+2} (n = 0, 2, 4) but also reveal their applications into homogeneous catalysis.[8-9] The diboration reaction introduce boryl groups across the unsaturated organic species which offers valued synthetic routes.[10-11] As a result, the diboration reaction turn out to be the most-used organometallic reaction in organic synthesis of bimetallic-diborane(4).[12]

Diborane(4) compounds have been utilized as relevant starting materials to generate electron precise metal–boron bonds.[13] For example, borylene species of the type [{$LMn(CO)_2$}$_2$BR] (L = η^5-C_5H_5 or η^5-C_5Me_5; R = ^{t}Bu or NMe_2) have been synthesised from diborane(4) reagents.[14] Hydrogen substituted diborane compounds, particularly, bi- or tri-metallic diborane(4) compounds, are generally very rare.[15] The derivatives of B_2H_6 as well as unsubstituted B_2H_6 are found to be stabilized by the metal fragments.[16-17] Our recent findings of bimetallic diborane(4), [{$Cp^*Mo(CO)_2$}$_2${μ-η^2:η^2-B_2H_4}] (**2**) and diborene(2) species, [{$Cp^*Mo(CO)_2$}$_2$$B_2H_2W(CO)_4$][18] led us to explore the chemistry of **2** further, not only to increase the library but also to have substituted products within the realm of transition-metal borane chemistry.

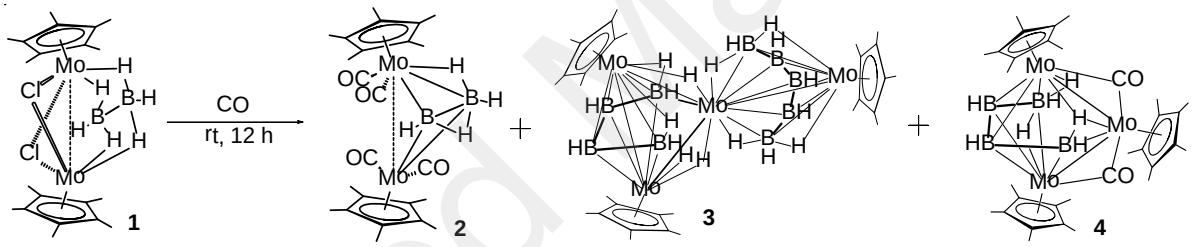
The progress of metallaborane chemistry expanded significantly with the advancement of generalized synthetic route of group 4-9 metallaboranes.[19-20] The electronic contributions of both metal and boranes in polyhedral metallaborane clusters are expressed differently than the individuals.[21] As a result, this makes the metallaborane chemistry more versatile and it continues to surprise with the various structural types and unusual reactivities.[22] Understanding the structures of such clusters required the early development of theories on chemical bonding, e.g. Wade–Mingos or Jemmis rules for electron counting.[23,24] Herein, in this article, we report the syntheses and structural characterizations of various metallaborane clusters using CO ligands or mono metal carbonyl fragments, [M'$(CO)_5$.thf], (M' = Mo and W). In addition, we have

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performed the theoretical calculations on the ground of density functional theory (DFT) to provide further insight into their bonding and electronic structures.

RESULTS AND DISCUSSION

The metallaborane clusters of early transition metals typically yielded uncommon problems that challenged the cluster electron counting rules.[25] Therefore, as a part of our ongoing research on cluster growth chemistry, we intended to explore the metallaboranes of group-6 transition metals.[18,26] As a result, we treated $[(Cp^*Mo)_2(\mu\text{-}Cl)_2B_2H_6]$[27] with CO, which resulted in the formation of compounds **2, 3**[28] and **4** (Scheme 1). A preliminary result of diborane(4) complex **2** has recently been reported.[18] The detailed spectral and structural characterization of **2** and **4** is discussed below.



Scheme 1. Synthesis of molybdaborane clusters, **2-4.**

$[\{Cp^*Mo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}B_2H_4\}]$ (2): Reaction of $[(Cp^*Mo)_2(\mu\text{-}Cl)_2B_2H_6]$[27] (**1**) at ambient temperature with CO gas produces **2** that has been characterized by ^{11}B, ^{1}H and ^{13}C NMR, and IR. The mass spectrum and IR data of **2** suggests that it is a derivative of **1**. Although, the room temperature ^{11}B NMR was uninformative, at -40 °C it showed two broad resonances appeared at δ = 27.3 and -31.1 ppm, whereas the ^{1}H NMR showed three broad signals at δ = -11.3, -0.86 and 3.28 ppm along with the signal due to Cp* ligands.

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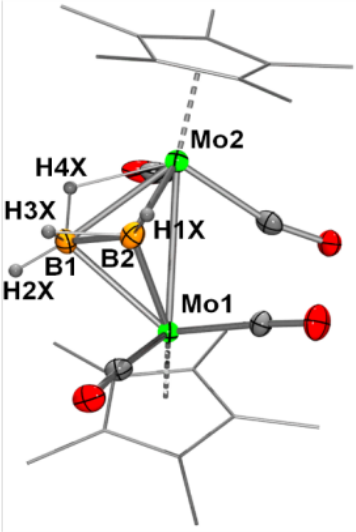


Figure 1. Molecular structure of **2** (hydrogen atoms on the Cp* ligands are omitted for clarity). Selected bond lengths (Å) and angles (deg): B1-B2 1.700(12), B1-Mo1 2.348(9), B1-Mo2 2.420(9), B2-Mo1 2.222(8), B2-Mo2 2.366(8), Mo1-Mo2 3.0617(8), B1-H2X 1.12(10), B1-H4X 1.13(7), B1-H3X 1.09(9), B2-H1X 1.25(11), B2-H3X 1.46(9), B2-B1-Mo2 67.5(4), Mo1-B1-Mo2 79.9(3), B2-B1-H2X 139(5), Mo1-B1-H4X 115(4), Mo2-B1-H4X 47(4), H2X-B1-H4X 109(6), B2-B1-H3X 58(5), Mo1-B1-H3X 114(5), Mo2-B1-H3X 104(5), B1-B2-Mo1 72.2(4), B1-B2-Mo2 70.9(4), Mo1-B2-Mo2 83.7(3), B1-B2-H1X 155(5), Mo1-B2-H1X 128(5), Mo2-B2-H1X 121(5), Mo1-B2-H3X 105(4), Mo2-B2-H3X 95(4), B2-Mo1-B1 43.6(3), B1-Mo1-Mo2 51.1(2), B2-Mo2-Mo1 46.1(2), B1-Mo2-Mo1 49.0(2).

In order to determine the solid-state structure of **2**, the X-ray structure analysis was undertaken. The solid-state X-ray structure (Figure 1) can be seen as [{Cp*Mo(CO)$_2$}$_2${μ-η^2:η^2-B$_2$H$_4$}]. Most noteworthy feature of **2** is the presence of a [B$_2$H$_4$] moiety that is linked between two [Cp*Mo(CO)$_2$] moieties such that the B-B bond positions perpendicular to the Mo-Mo bond. The B-B bond distance of 1.700(12) Å in **2** is significantly shorter as compared to the base stabilized diborane(4) complexes, for example, [(PMe$_3$)$_2$B$_2$H$_4$ZnCl$_2$] (1.814(6) Å), [(PMe$_3$)$_2$B$_2$H$_4$Cr(CO)$_4$]

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(1.748(11) Å) and $[(PMe_3)_2B_2H_4W(CO)_4]$ (1.74(41) Å).[29] The Mo–Mo bond distance of 3.0617(8) Å is consistent with single bond order which is comparable to those of $[\{CpMo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}C_2R_2\}]$ (R = H/Et/Ph) (avg. 2.971(1) Å) and $[\{CpMo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}P_2\}]$ (3.022(1) Å).[30] Among the different isomers of diborane(4), the singly bridged isomer with C_s symmetry is very rare and beside **2** the only such example hypothesised is $[Co_2(CO)_6B_2H_4]$.[31]

Although the Cp-Mo-Mo-Cp axis in $[\{CpMo(CO)_2\}_2]$ is nearly linear, the structures of both $[\{CpMo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}C_2H_2\}]$ and **2** show distinct Mo-Mo-(Cp/Cp*) angles (Chart 1). The dihedral angle of Cp/Cp*-M-M-Cp/Cp* (147.6° vs 160.3°, see Chart 1) is related to the Cp/Cp*-M-M angle. The longer distance of B-B in comparison to that of C-C would require a more diffuse orbitals on the Cp*M-M-Cp* in **2**. This can be achieved by bending the Cp*s farther away from the Mo-Mo axis, i.e. Cp*-M-M angle has to decrease. However, this decrease brings the two Cp*s together. The molecule reduces the Cp*---Cp* non-bonding interactions by increasing the Cp*-M-M-Cp* dihedral angle. The opening of this angle and the positioning of the carbonyls and Cp* rings over Mo-Mo bond results from the longer B-B distance (in relation to C-C) that requires a rehybridization at the metal centre for better Mo-B bonding.

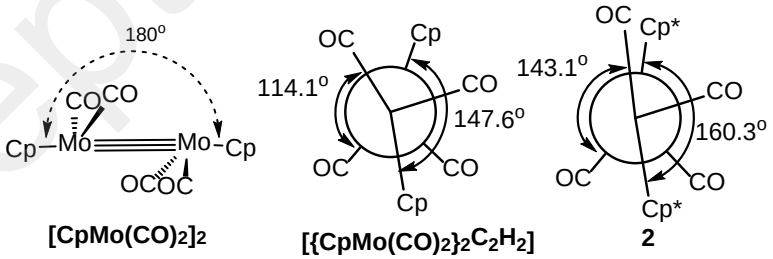


Chart 1. Schematic representation of the dihedral angles of Cp'···Mo-M···Cp' in $[CpMo(CO)_2]_2$, $[\{CpMo(CO)_2\}_2C_2H_2]$ and **2** (Cp' = Cp or Cp*).

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Electron counting rules along with isolobal analogy suggest a close resemblance of **2**[32] with [{CpMo(CO)$_2$}$_2${μ-η^2:η^2-C$_2$H$_2$}]. Thus, compound **2** is a true mimic of Cotton's dimolybdenum alkyne complexes, [{CpMo(CO)$_2$}$_2${μ-η^2:η^2-C$_2$R$_2$}] (R = H/Et/Ph).[30a] This further reveals that the [B$_2$H$_4$] fragment apparently binds with the fragment [Cp*Mo(CO)$_2$]$_2$ (16 electron at each metal center) through four-electrons so that each Mo atom attain the 18 electron count, which is vitually similar to the bonding of RC≡CR in [{CpMo(CO)$_2$}$_2${μ-η^2:η^2-C$_2$R$_2$}]. On the other hand, adding 6 skeletal electron fragment B$_2$H$_4$ (2 BH + 2 H) to two {CpMo(CO)$_2$} moieties (6 skeletal electron) makes a total of 12 skeletal electrons or 6 skeletal electron pairs (seps). [23,33] This leads to a *nido* geometry and makes compound **2** isoelectronic to [{CpMo(CO)$_2$}$_2${μ-η^2:η^2-C$_2$R$_2$}].

Homoleptic cluster, [(Cp*Mo)$_3$(μ-H)$_2$(μ_3-H)(μ-CO)$_2$B$_4$H$_4$] (4): Compound **4** was isolated as a brown solid in 10% yield. The mass spectrum of **4** showed a molecular ion peak at *m/z* 807. The IR spectra showed strong absorption bands at 1787 and 1729 cm^{-1} corresponding to bridging CO ligands. The ^{1}H NMR displayed two signals at δ = 1.95 and 1.91 ppm in 1:2 ratio that rationalizes the presence of two different types of Cp* ligands. This implies two chemically dissimilar Mo environments. The presence of Cp* and CO ligands have also been supported by ^{13}C NMR. In addition, the ^{1}H NMR shows three resonances in the upfield region at δ = -13.17, -9.35 and -8.85 in 1:1:1 ratios, which may be due to the presence of bridged Mo-H-Mo and Mo-H-B protons. The ^{11}B{^{1}H} NMR spectrum shows four resonances at δ = 88.3, 83.2, 46.3 and 23.5 ppm in 1:1:1:1 ratio, which demonstrates the presence of four dissimilar boron atoms in **4**. However, identity of the molecule was unambiguous and definitive structural characterization was determined by X-ray diffraction analysis on a suitable crystal obtained from a solution of compound **4** in hexane-CH_2Cl_2 solvent mixture at -10 °C.

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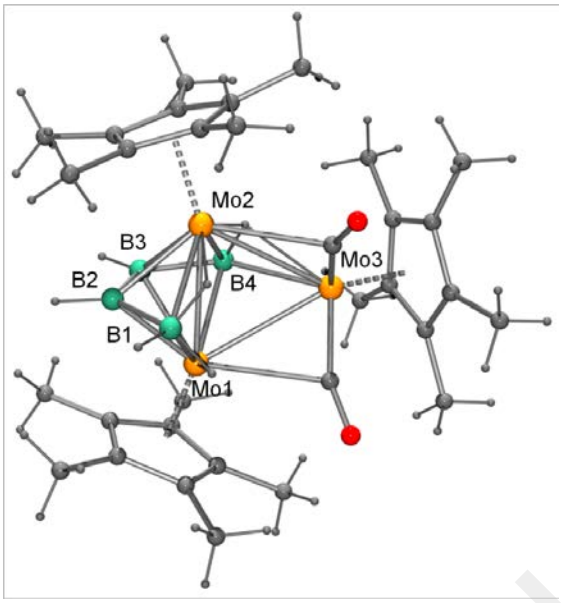


Figure 2. Molecular structure and labeling diagram for **4**. Selected bond lengths (Å) and bond angles (°): Mo1-Mo2 2.8338(5), Mo2-Mo3 2.9843(5), B1-B2 1.731(8), B2-B3 1.713(8), B3-B4 1.708(7), B1-Mo1 2.305(6), B1-Mo2 2.307(6), B2-Mo2 2.180(5), B2-Mo1 2.205(5), B4-Mo2 2.277(6), B4-Mo1 2.342(5), Mo1-Mo2-Mo3 61.479(13), Mo2-Mo1-Mo3 61.752(13), Mo1-Mo3-Mo2 56.769(12).

The molecular structure of **4**, shown in Figure 2, can be seen as $[(Cp^*Mo)_3(\mu\text{-}H)_2(\mu_3\text{-}H)(\mu\text{-}CO)_2B_4H_4]$ which is fully consistent with the solution spectroscopic data. The polyhedral skeletal electron pair theory (PSEPT)[23,33] showed that compound **4** can be viewed as 6 sep (skeletal electron pair) *oblato-nido* hexagonal bipyramidal cluster. As a result, the core geometry appears to be similar to $[(Cp^*Mo)_2B_5H_9]$ (**I**)[34] where one of the $\{BH_2\}$ vertices in the open face has been replaced by isoelectronic fragment $\{Cp^*Mo(CO)_2\}$**.** This resulted in a small stabilization of the LUMO but destabilization of the HOMO as compared to **I** (Table 1). Note that all the three Mo-Mo bond distances are dissimilar in **4**. Although, the axial Mo1-Mo2 distance is similar to that of **I** (2.833 vs 2.81Å), the peripheral Mo-Mo distances are significantly

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longer (3.004 and 2.946 Å). This has also been reflected in the WBI (Wiberg bond index) values; for example, WBI of Mo1-Mo2 (0.73) is ca. 1.5 times greater then the other Mo-Mo bonds (Mo2-Mo3: 0.39 and Mo1-Mo3: 0.51). The MO analyses of **4'** (a Cp analogue of **4**) show that both the HOMO and HOMO-1 (Figures 3b-c and S21) mostly involves Mo-Mo bonding. Further, to gain some insights into the Mo-Mo interactions in **4'** we have carried out the NBO analysis. The NLMOs clearly demonstrate the Mo-Mo interactions (Figures 3a and S22). The molecular graph, Laplacian plot of electron density and WBI values indicate that the CO ligands are closer to Mo3 that can also be seen from the X-ray structure analyses.

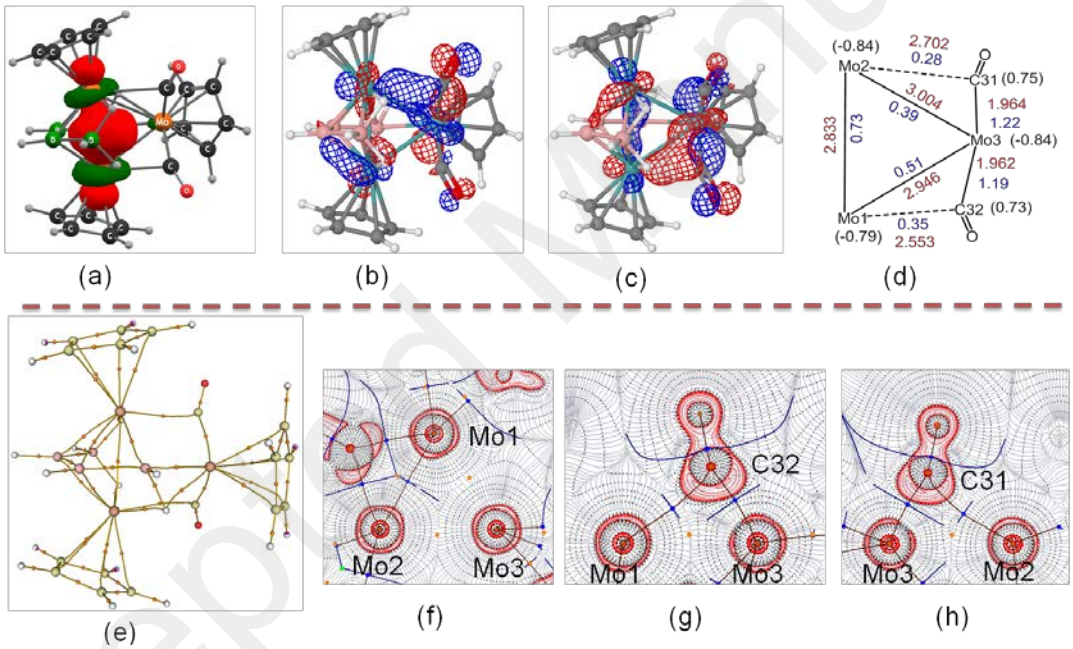
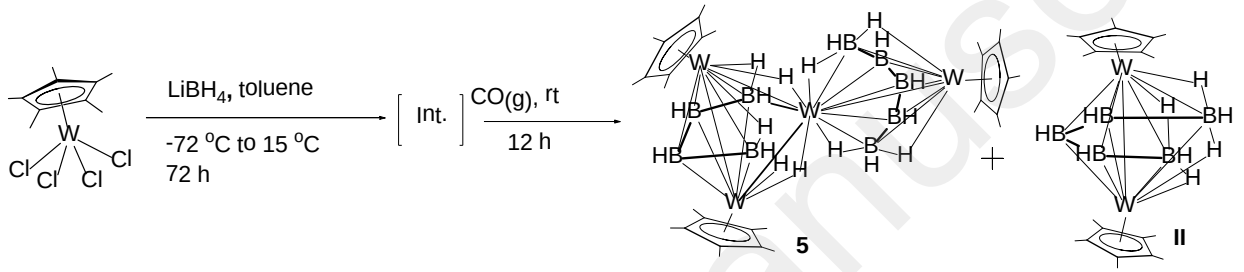


Figure 3. Investigating the Mo-Mobonding in **4'**. (a) natural localized molecular orbital (NLMO) obtained from NBO analysis, (b) HOMO, (c) HOMO-1, (d) The Mo_3-core (WBIs are shown in blue, bond distances in red and natural charges in parenthesis), (e) molecular graphs for the model **1'**, contour line diagrams of the Laplacian of electron density inMo_3 plane (f) and Mo-C-Moplanes (g-h) [Solid red lines indicate areas of charge concentration ($\nabla^2\rho(r) < 0$), while dashed black lines show areas of charge depletion ($\nabla^2\rho(r) > 0$)].

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Table 1. DFT calculated energies of the HOMO and LUMO (eV), HOMO-LUMO gaps (ΔE = E_{LUMO} - E_{HOMO}, eV) of $[(CpM)_2B_5H_9]$ (M = Mo(I') and W(II')), **3'-5', 7'** and **8'.**

	I'	**II'**	**3'**	**4'**	**5'**	**7'**	**8'**
HOMO	-4.40	-4.84	-4.99	-4.60	-4.84	-5.07	-3.52
LUMO	-1.32	-1.44	-3.25	-3.19	-2.99	-3.63	-1.10
ΔE	3.08	3.40	1.74	1.41	1.85	1.44	2.42



Scheme 2. Synthesis of tungstaborane cluster, **5.**

$[(Cp^*M)_3MB_9H_{18}]$ [M= Mo (3), W (5)]: Isolation and structural characterization of diborane(4) molybdenum complex **2** led us to explore the chemistry of tungsten system mainly with an objective to isolate the analogue of **2**, $\{Cp^*W(CO)_2\}_2B_2H_4$. Although the objective of isolating analogous **2** was not achieved, we have isolated a vertex-fused tungsten cluster $[(Cp^*W)_3WB_9H_{18}]$ (**5**). Compound **5** was isolated as a byproduct in parallel to the formation of $[(Cp^*W)_2B_5H_9]$ (**II**).[35] These compounds can be separated by preparative thin-layer chromatography (TLC), allowing the characterization of pure **5**. The mass spectrum of **5** showed a molecular ion peak [*m/z*] corresponding to $C_{30}H_{63}W_4B_9$, while the IR spectrum confirms the absence of CO ligands. The $^{11}B\{^1H\}$ NMR spectrum displayed a set of nine resonances between δ = 55.3 and 21.9 ppm (from high frequency to low frequency), distributed over a chemical shift range of *ca.* 30 ppm.[36] Besides the BH terminal protons, the ^1H NMR showed highly shielded upfield resonances which may be due to the presence of bridging W-H-B protons. In addition,

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the ^{1}H and ^{13}C NMR spectra of **5** are in consistent with the existence of two chemically different Cp* ligands.

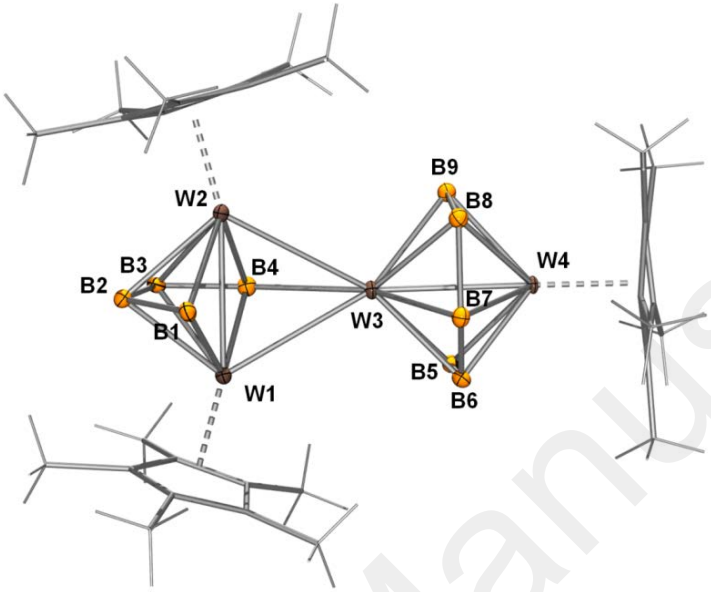


Figure 4. Molecular structure of **5**. Ellipsoids are set at 30% probability (hydrogen atoms on the Cp* ligands are omitted for clarity) Selected bond lengths (Å) and bond angles (degree).: W1-W2 2.839(2), W1-W3 3.013(2), W3-W4 2.856(2), W2-W3 3.046(2), W1-B1 2.258(4), W1-B4 2.6181(5), B1-B2 1 699(6), B2-B3 1.760(6), B8-B9 1.838(8), B7-B8 1.770(8), W1-W3-W2 55.8(5), W1-W2-W3 61.4(5), W2-W1-W3 62.6(5), W1-B1-W2 77.1(13), B1-B2-B3 120.0(3).

The single crystal X-ray structure analysis was performed on a brown crystal of **5** that was obtained from the concentrated hexane solution at -4 ^{o}C. As shown in Figure 4, the molecular structure of **5** can be viewed as a fusion of two clusters through a common single atom vertex. The X-ray structure and the solution state spectroscopic data are consistent with the molybdenum analogue, $[(Cp^{*}Mo)_3MoB_9H_{18}]$ (**3**).[28] Although, all the hydrogen atoms were not located in the X-ray diffraction study, evidence for their presence has been unequivocally supported by ^{1}H NMR spectroscopy. Core structure of **5** is composed of two independent

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clusters, *oblato-nido* $\{W_2B_5\}$ and *oblato-arachno* $\{W_2B_4\}$ hexagonal bipyramid clusters fused in perpendicular fashion through a naked tungsten atom (W3). The dihedral angle between the basal planes of the corresponding B_5- and B_4- belts is 85.53°. This brings six B atoms and four hydride ligands in the coordination sphere of W3 in expense of a Cp* ligand of compound **II**. Both the cluster fragments and their structural parameters resemble with those of **II**. However, the solid state X-ray structure reveals two different sets of W-W bond lengths [W1-W2: 2.839(2), W3-W4 (2.856(2) Å and W1-W3: 3.013(2), W2-W3: 3.046(2) Å] with the average W-B and B-B bond distances comparable with the reported tungstaborane compounds (Table 2). Longer W-W bond lengths may be attributed to the unique chemical environment of the naked W3. This observation is in line with the Laplacian plot (Figure 5b) as well as ELF plot (Figure 5c), as both the diagrams reveal high electron density concentration along W1-W2 bond than W1-W3 and W2-W3. Similar bonding situation has been observed in $[\{1\text{-}(Cp^*Ru)(\mu\text{-}H)B_4H_9\}_2Ru]$ that showed shorter metal-metal bond distance compared to $[1,2\text{-}(Cp^*Ru)_2(\mu\text{-}H)B_4H_9]$.[37] The only difference between **3** and **5** is their ^{11}B NMR chemical shifts. The ^{11}B resonances of **5** appeared at upfield region as compared to **3**. This can well be understood from the fact that more charge accumulation occurred over B centers in **5** compare to **3** (Figure 5e).

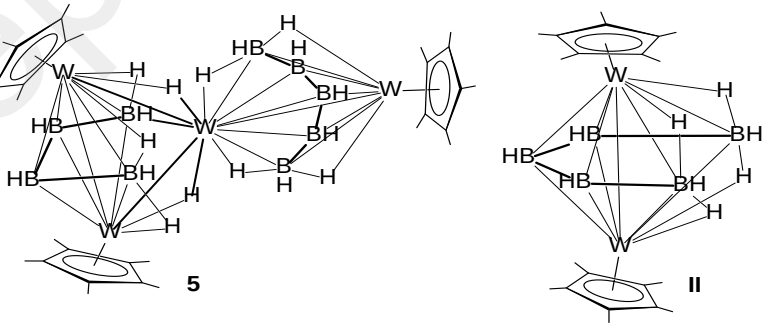


Chart 2. Structural equivalence of **5** with (**II**).

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The 18-electron count rule for **5** becomes more apparent when one considers that the W atom in $[(Cp^*W)_2B_5H_9]$ (**II,** Chart 2) receives five electrons from the Cp* ligand. This is comparable to a case in which $\{(Cp^*W)_2B_4H_9\}$ fragment in **5** contributes five electron to the bare tungsten atom (W3, Figure 4). Thus, the replacement of one of the Cp* ligands from **II** with a five electron donor fragment $\{(Cp^*W)_2B_4H_9\}$, the 18-electron count for the bare W atom satisfied and hence this may be considered as an analogue to **II**. Both the NBO and MO analysis of **5** and **II** support these arguments (Figures 5a,d and S23-S24). Further, the HOMO shows large localization of electron density on the metal centers (Figure 5a and S23).

$$\underset{\text{CVE = 44}}{[(Cp^*W)_2B_4H_{10}]} + \underset{\text{CVE = 46}}{[(Cp^*W)_2B_5H_9]} \longrightarrow \underset{\text{CVE = 84}}{[(Cp^*W)_3WB_9H_{18}]} + \underset{\text{CVE = 6}}{Cp^*H} \qquad \text{Eqn. 1}$$

Cluster condensation or fusion by vertex, edge, or face sharing is quite familiar in polyhedral borane, metallaborane and carborane chemistry.[38] Electron counting of these types of clusters can well be understood following Mingos's fusion formalism[38f] and Jemmis's mno rule.[24] Existence of cluster **II** led us to propose a plausible pathway for the formation of compound **5**. From Eqn 1, it is reasonable to assume that, during the progress of the reaction, $[(Cp^*W)_2B_4H_{10}]$ and **II** underwent a cluster condensation reaction to produce **5** with the elimination of a Cp*. As a result $\{(Cp^*W)_2B_4H_9\}$ acts as a ligand for the formation of $\{(Cp^*W)WB_5H_9]$ (Chart 1).[39]

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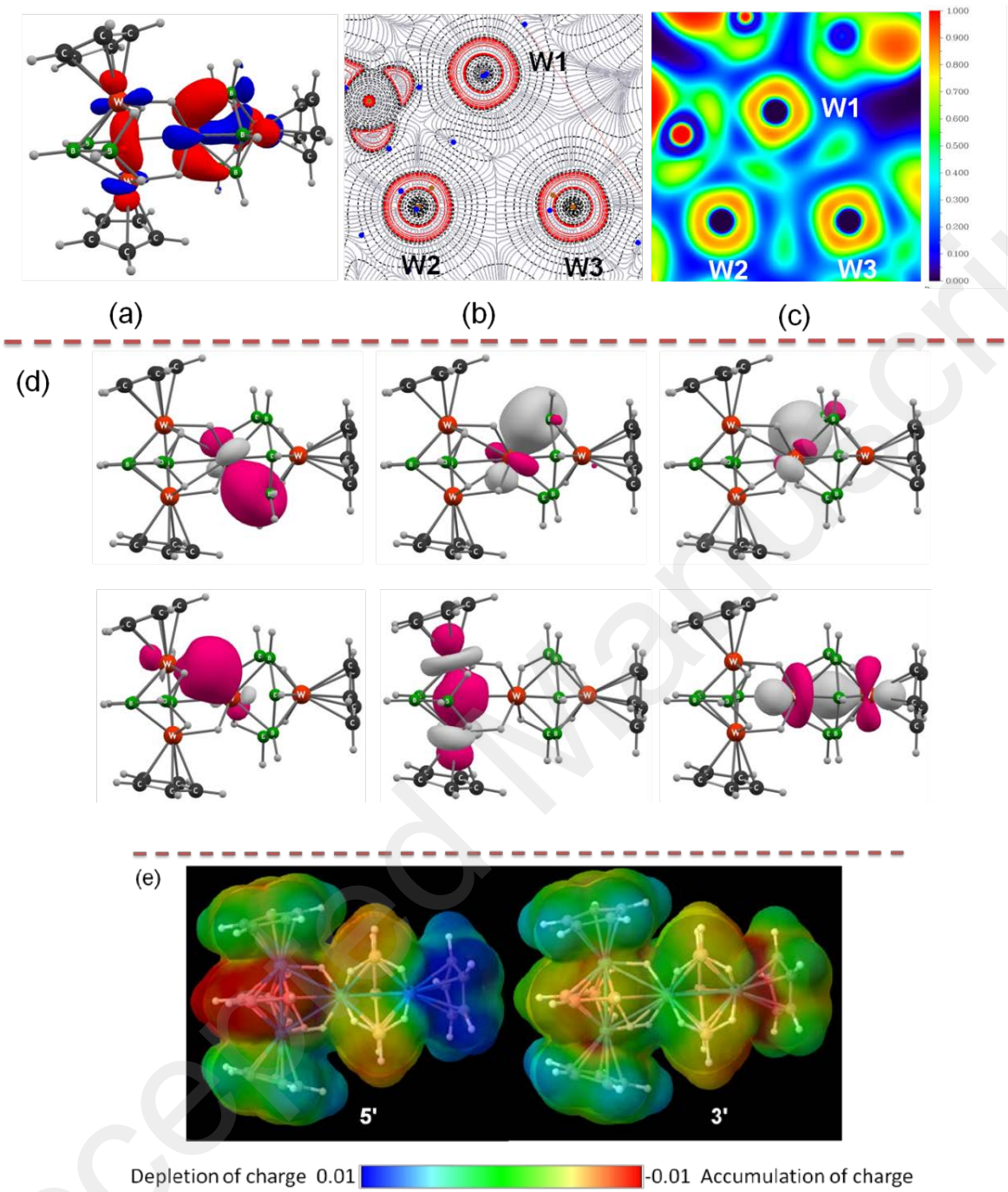
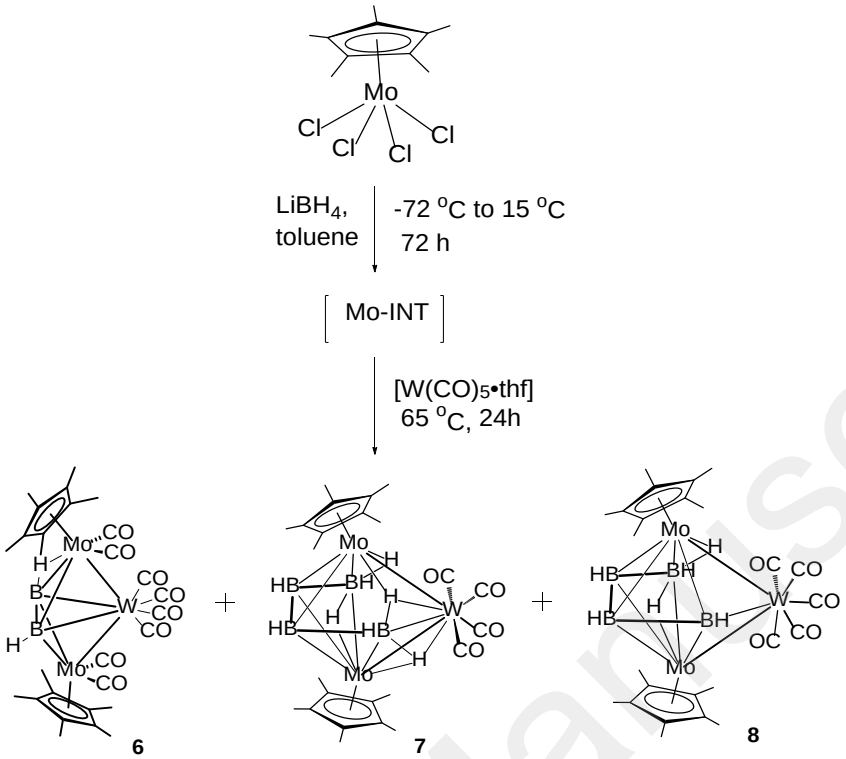


Figure 5. Bonding analyses of **5'** [(a) shows the HOMO, (b) contour line diagram of the Laplacian of electron density in the plane of W1-W2-W3, (c) showing ELF plot in the plane of W1-W2-W3, (d) selected natural localized molecular orbital (NLMO) obtained from NBO analysis, (e) Electrostatic potential analysis].

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Scheme 3. Synthesis of diborene(2), (**6)** and heterometallaboranes (**7** and **8**). [Mo-INT] is an *insitu* generated intermediate, obtained from the reaction of [Cp*$MoCl_4$] and [$LiBH_4$] at -72 °C.

[(Cp*Mo)$_2$B$_4$H$_8$W(CO)$_4$], 7 and [(Cp*Mo)$_2$B$_4$H$_6$W(CO)$_5$], 8: In order to replace {Cp*Mo $(CO)_2H$} unit in **4** with {$W(CO)_5$}, we treated the intermediate, obtained from the reaction of [Cp*$MoCl_4$] and [$LiBH_4$], with {$W(CO)_5$.thf}. The reaction yielded two trimetallic clusters [$(Cp^*Mo)_2B_4H_8$-$W(CO)_4$], **7** and [$(Cp^*Mo)_2B_4H_6W(CO)_5$], **8** along with the formation of diborene(2) complex, [{Cp*$Mo(CO)_2$}$_2B_2H_2$ $W(CO)_4$ (**6**), reported recently in a communication.[18] Since the landmark report of N-heterocyclic carbene (NHCs) stabilized diborene compounds by Robinson, the chemistry of unsaturated boron-boron species has extended significantly and in this respect, **6** is an exceptional entry to this diborene(2) species.[40,41] Compounds **7** and **8** were isolated as brown and green solids in moderate yields.

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Room temperature $^{11}B\{^{1}H\}$ NMR of compounds **7** and **8** appeared between δ = 27 and 103 ppm (from high frequency to low frequency), with signals of **8** being relatively downfield shifted. The IR spectra of both **7** and **8** showed the stretching frequencies corresponding to the CO ligands and BH_t. The ^{1}H NMR spectra of **7** and **8** reveal the presence of one equivalent of Cp* proton. Further, the mass spectrometric data suggest a close association between compounds **7** and **8**.

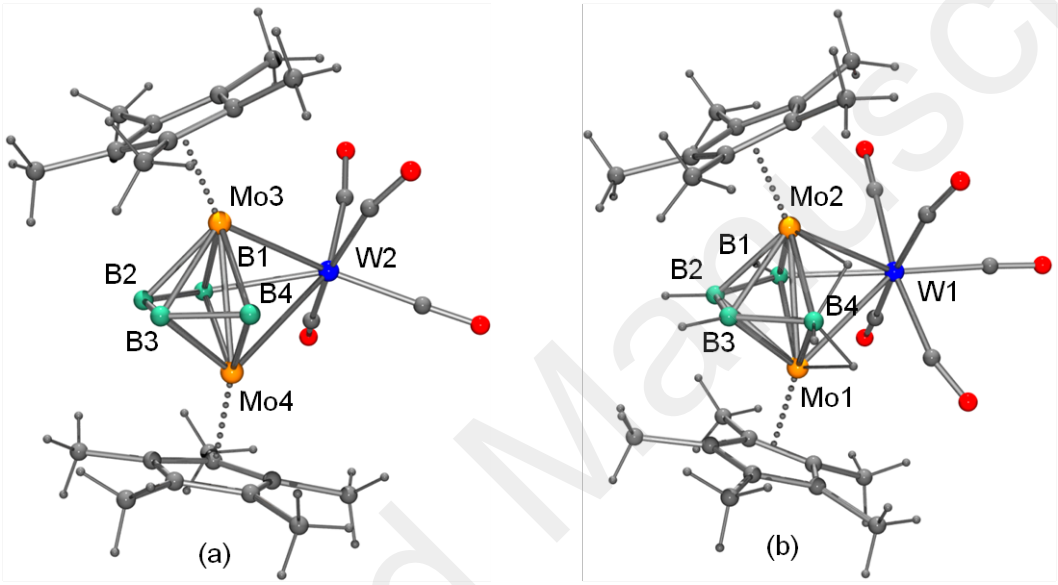


Figure 6. Molecular structure of (a) **7** and (b) **8**. Bridging hydride ligands and the hydrogen atoms attached to the boron atoms of **7** could not be located. Selected bond lengths (Å) and bond angles (^{o}). **7**: Mo3-B2 2.1821(15), Mo3-B3 2.2380(16), Mo3-B4 2.2671(16), Mo3-B1 2.2841(16), Mo3-Mo4 2.8275(18), Mo3-W2 2.9908(15), W2-B4 2.3955(7), B1-B2 1.7399(3), B2-B3 1.7185(6), B3-B4 1.7003(7). Mo4-W2-Mo3 55.95(4), Mo3-B1-Mo4 75.87(5), Mo3-B4-Mo4 75.73(5), **8**: B1-B2 1.706(5), B1-Mo1 2.283(3), B1-W1 2.315(3), B2-B3 1.739(5), B2-Mo1 2.175(3), B3-B4 1.744(5), B3-Mo2 2.180(4), B4-Mo2 2.309(3), Mo1-W1 2.9801(3), Mo1-Mo2 2.8308(3), Mo2-B1-Mo1 77.17(10), B1-W1-Mo1 49.13(8), Mo1-W1-Mo2 56.588(7).

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Single crystals of **7** and **8** suitable for X-ray diffraction were grown from hexane at -10 °C. The solid-state X-ray structures of **7** and **8**, shown in Figures 6 (a) and 7(b) respectively, are consistent with the spectroscopic data. The core geometries are analogous to $[(Cp^*Cr)_2B_4H_8Fe(CO)_3]$ and $[(Cp^*Cr)_2Co\text{-}(CO)_3B_4H_7]$.[42] Note that, the hydrogen atoms attached to boron could not be located by solid state X-ray diffraction study of **7**[43], however their positions were fixed based on ^{1}H NMR as shown in Scheme 3. As $\{W(CO)_4\}$ and $\{W(CO)_5\}$ are isolobal to (BH_3) and (BH), compounds **7** and **8** are formally related to 46 cve (cluster valance electron) cluster **I**. Thus, the core geometries of **7** and **8** can be viewed as a 6 *sep* oblato-*nido* hexagonal bipyramid or $\{Mo_2B_3H_3\}$ trigonal-bipyramidal core face capped by (BH_3) and $\{W(CO)_4\}$ or $\{W(CO)_5\}$ fragments. The Mo-Mo bond distances in **7**, **8** and **I** are comparable (**7**: 2.8275(18), **8**: 2.8308(3), $[(Cp'Mo)_2B_5H_5(\mu\text{-}H)_4]$: 2.80 Å; Table 2), however they are shorter than those observed in $[CpMo(CO)_3]_2$ (3.22 Å) and $[Mo_2(CO)_{10}]^{2-}$ (3.123 Å).[44]

The HOMO and HOMO-1 of **8'** (Figures 7b and c) shows electron localization over the metal centres. The HOMO, in particular, shows that the $\{W(CO)_5\}$ fragment interacts through its half occupied π-symmetry orbital. The empty σ orbital of the $\{W(CO)_5\}$ fragment interacts with a low lying filled Cp*Mo orbital leading to donor-acceptor interactions (Figures S26-S27). This was further supported by the electron density distribution analysis where the Laplacian plot showed more electron density along the Mo1-Mo2 than W1-Mo1/Mo2 (Figure 7f). As a result, the W centre attains the 18-electron rule. In addition, the MO analyses (HOMO and HOMO-1, Figures 7b and c, respectively) demonstrate week interactions between the Mo based d orbitals and π* of the CO ligands that established the week nature of Mo-CO interactions and more like terminal to W1 centre. This observation was also supported by the WBI values and the molecular graphs (Figures 7d-e). On the other hand, the $[W(CO)_4]$ moiety in **7** attached to $[(Cp^*Mo)_2B_4H_8]$

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through two face-bridging hydrogen and a pair of Mo-W interactions (Figure S25). An interesting statement that can be concluded from the NBO charge analysis of **8**' is the diverse oxidation states of the metal centres. The comparatively reduced oxidation state of W1 (natural charge -1.57, whereas that of Mo centres are -0.87 and -0.84) indicates the flow of electron density predominantly towards W1 centre. This further supports the accepter behaviour of $[W(CO)_5]$ moiety.

Table 2. Selected structural parameters of **7** and **8** and other related compounds.

Compounds	sep	*d*(M-M) (Å)	*d*(M-B) (Å)[a]	*d*(B-B) (Å)[a]
$[(Cp^*Cr)_2B_4H_8Fe(CO)_3]$	6	2.71[a]	2.15	1.72
$[(Cp^*Cr)_2Co(CO)_3B_4H_7]$	6	2.69[b]	2.13	1.69
$[(Cp'Mo)_2B_5H_5(\mu\text{-}H)_4]$	6	2.81	2.26	1.73
$[(Cp^*W)_2B_5H_5(\mu\text{-}H)_4]$	6	2.82	2.26	1.71
$[(Cp^*Mo)_2B_5H_9Fe(CO)_3]$	7	2.94[c]	1.75	1.76
$[(Cp^*W)_2B_5H_9Fe(CO)_3]$	7	2.93[d]	1.79	1.77
7	6	2.83	2.24	1.72
8	6	2.83	2.24	1.73

[a] average distance, Cp' = (η^5-MeC_5H_4)

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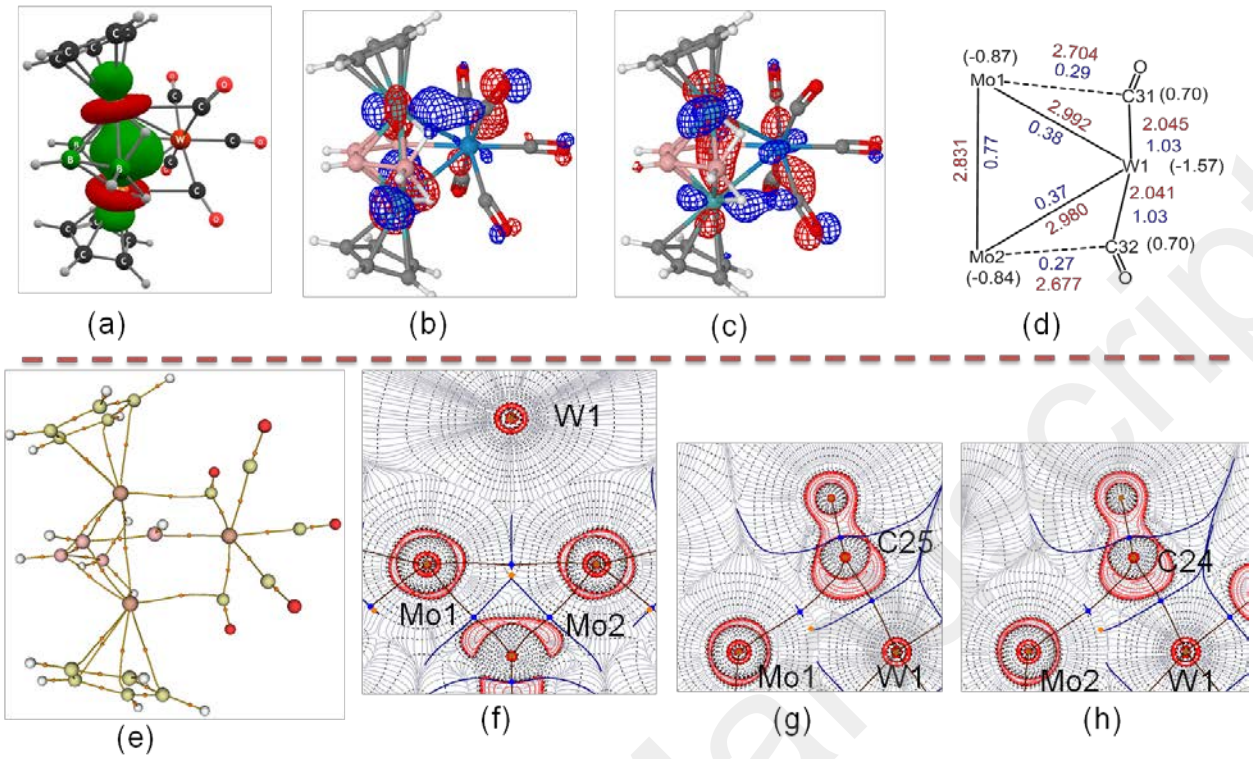


Figure 7. Mo-Mo bonding in **8'** (Cp analogue of **8**). (a) natural localized molecular orbital (NLMO) obtained from NBO analysis, (b) HOMO, (c) HOMO-1, (d) The Mo_3-core (WBIs are shown in blue, bond distances in red and natural charges in parenthesis), (e) molecular graphs for the model **8'**, contour line diagrams of the Laplacian of electron density in Mo_3 plane and (f) Mo-C-Mo planes (g-h) [solid red lines indicate areas of charge concentration ($\nabla^2\rho(r) < 0$), while dashed black lines show areas of charge depletion ($\nabla^2\rho(r) > 0$)].

CONCLUSION

In summary, we have described the synthesis of molybdenum diborane(4) complex, $[\{Cp^*Mo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}B_2H_4\}]$ (**2**) that not only shows the presence of unique B_2H_4 of C_s type, it also mimics Cotton's dimolybdenum-acetylene complex, $[\{CpMo(CO)_2\}_2C_2H_2]$. Our pursuit of growing the library of this type of complexes led us to isolate homoleptic molybdaboranes, $[(Cp^*Mo)_3(\mu\text{-}H)_2(\mu_3\text{-}H)(\mu\text{-}CO)_2B_4H_4]$ (**4**) and a vertex-fused cluster,

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$[(Cp^*W)_3WB_9H_{18}]$ (**5**). Further, we have described the synthesis of two mixed metallaboranes, $[(Cp^*Mo)_2B_4H_8W(CO)_4]$ (**7**) and $[(Cp^*Mo)_2B_4H_6W(CO)_5]$ (**8**) along with diborene(2) complex, $[\{Cp^*Mo(CO)_2\}_2B_2H_2\ W(CO)_4]$ (**6**). Electron counting as well as molecular orbital analyses suggest that metallaboranes **7** and **8** are the analogues of parent ($[(Cp^*Mo)_2B_5H_9]$) cluster. These results exemplify the presence of multimetallic molecular clusters with exciting geometry and bonding.

EXPERIMENTAL SECTION

General Procedures and Instrumentation. All the manipulations were conducted under an atmosphere of dry argon or in vacuo using standard Schlenk line or glove box techniques. Solvents (hexane, toluene, thf) were purified by distillation from appropriate drying agents-(sodium/benzophenone) under dry argon prior to use. $CDCl_3$ and C_6D_6 were degassed by three freeze-pump-thaw cycles and stored over molecular sieves. Compounds $[Cp^*MCl_4]$ (M = Mo and W), $[M'(CO)_5.thf]$, (M' = Cr, Mo and W) were prepared according to literature method,[45] while other chemicals such as, $[LiBH_4.thf]$ 2.0 M in THF, Cp*H, n-BuLi, $[Cr(CO)_6]$, $[Mo(CO)_6]$ and $[W(CO)_6]$ were obtained commercially (Aldrich) and used as received. MeI was purchased from Aldrich and freshly distilled prior to use. The external reference for the ^{11}B NMR, $[Bu_4N(B_3H_8)]$ was synthesized with the literature method.[46] Preparative thin layer chromatography was performed with Merck 105554 TLC silica gel 60 F_{254}, layer thickness 250 µm on aluminium sheets (20x20 cm). NMR spectra were recorded on 400 and 500 MHz Bruker FT-NMR spectrometers. The residual solvent protons were used as reference (δ, ppm, d_6-benzene, 7.16, $CDCl_3$, 7.26), while a sealed tube containing $[Bu_4N(B_3H_8)]$ in d_6-benzene (δ_B, ppm, –30.07) was used as an external reference for the ^{11}B NMR. The infrared spectra were recorded on a Nicolet iS10 spectrometer. Electrospray mass (ESI-MS) spectra were recorded on

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a Qtof Micro YA263. The photo-reactions described in this report were conducted in a Luzchem LZC-4 V photo reactor, with irradiation at 254−350 nm. MALDI-TOF mass spectra were recorded on a Bruker Ultraflextreme by using 2,5-dihydroxybenzoic acid as a matrix and a ground steel target plate and CH analyses were obtained on Perkin Elmer Instruments series II model 2400.

Synthesis of 4: A solution of **1** (0.140 g, 0.25 mmol) in toluene was treated with carbon monoxide over a period of 15 mins and the reaction mixture was stirred for 12 h at ambient temperature. The solvent was evaporated in vacuo and the crude was extracted with hexane: CH_2Cl_2 (80:20) solvent mixture and filtered through a frit using celite. After removal of solvent from the filtrate, the residue was subjected to chromatographic work-up using silica gel TLC plates. Elution with hexane:CH_2Cl_2 (75:25) solvent mixture afforded yellow **2** (0.052 g, 32 %),[18] brown **3** (0.035 g, 15.4 %) and brown **4** (0.020 g, 10%).

4: HRMS (ESI^+): *m/z* calculated for $[C_{32}H_{52}B_4Mo_3O_2+H]^+$: 807.1579; found 807.1588; $^{11}B\{^1H\}$ NMR (22 °C, 128 MHz, $CDCl_3$): δ = 23.5 (br, 1B), 46.3 (br, 1B), 83.2 (br, 1B), 88.8 (br, 1B), ^1H NMR (22 °C, 500 MHz, $CDCl_3$): δ = 4.82 (br, 4H, B$\underline{H}_t$), 1.95 (s, 15H, 1Cp*), 1.91 (s, 30H, 2Cp*), -8.85, -9.35, -13.17 (br, 3H, Mo–$\underline{H}$–B);$^{13}C\{^1H\}$ NMR (22 °C, 125 MHz, $CDCl_3$): δ = 187.2, 183. 9 (CO), 102.4, 100.2 ($\underline{C}_5Me_5$), 13.0, 10.8 ($C_5\underline{Me}_5$); IR (hexane) ν/cm^{-1}: 2420(BH_t),1787(CO), 1729 (CO).

Synthesis of 5: In a flame-dried Schlenk tube [Cp*WCl_4], (0.1 g, 0.22 mmol) in 15 mL of toluene was treated with 5-fold excess of [$LiBH_4$.THF] (0.55 mL, 1.1 mmol) at -40°C and allowed to stir at allowed to stir at 10 °C for 48 hour.In the reaction mixture CO(g) passed for 12 hour at room temperature. After removal of toluene, the residue was extracted into hexane and

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filtered through a frit using Celite. After removal of solvent from the filtrate, the residue was subjected to chromatographic workup using silica gel TLC plates. Elution with a hexane/CH_2Cl_2 (75:15 v/v) mixture yielded brown **5** (0.030 g, 10.9%) along with known [$(Cp^*W)_2B_5H_9$] (0.048 g, 31%).

5: HRMS (ESI^+): *m/z* calculated for $[C_{30}H_{63}B_9W_4+H]^+$: 1259.3883; found 1259.3915;. $^{11}B\{^1H\}$ NMR (22 °C, 128 MHz, $CDCl_3$): δ = 55.3 (br, 1B), 51.0 (br, 1B), 47.8 (br, 1B), 45.8 (br, 1B), 44.6 (br, 1B), 40.7 (br, 1B), 27.6 (br, 1B), 24.5 (br, 1B), 21.9 (br, 1B).^1H NMR (22 °C, 500 MHz, $CDCl_3$): δ= 6.19 (br, BH_t), 2.22 (s, 15H, 1Cp*), 2.20 (s, 15H, 1Cp*) 2.19 (s, 15H, 1Cp*), -1.79 (br, 1H, B–H–B), -3.61 (br, 1H, B–H–B), -7.58 (br, 1H, W–H–B), -8.56 (br, 2H, W–H–B), -9.58 (br, 1H, W–H–B), -10.64 (br, 1H, W–H–B), -13.08 (br, 1H, W–H–B). $^{13}C\{^1H\}$ NMR (22 °C, 100 MHz, $CDCl_3$): δ =108.8 ($\underline{C}_5Me_5$), 107.2 ($\underline{C}_5Me_5$), 12.6 ($C_5\underline{Me}_5$), 11.6 ($C_5\underline{Me}_5$). IR (hexane) ν/cm^{-}: 2473(BH_t).

Synthesis of 7 and 8: In a flame-dried Schlenk tube [Cp^*MoCl_4], (0.1 g, 0.27 mmol) in 10 mL of toluene was treated with 5-fold excess of [$LiBH_4$.THF] (0.7 mL, 1.4 mmol) at -40°C and allowed to stir at room temperature for one hour. After removal of toluene, the residue was extracted into hexane and filtered through a frit using Celite. The brownish-green hexane extract was dried in vacuo, and taken in 10 mL of THF and heated at 65°C with [$W(CO)_5$.THF] for 24 hours. The solvent was evaporated in vacuo and residue was extracted into hexane and passed through celite. After removal of solvent from the filtrate, the residue was subjected to chromatographic workup using silica gel TLC plates. Elution with a hexane/CH_2Cl_2 (80:10 v/v) mixture yielded dark brown **7** (0.025 g, 11.4%) and green **8** (0.025 g, 11%).

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7: HRMS (ESI$^+$): *m/z* calculated for $[C_{24}H_{38}B_4O_4Mo_2W_1+Na]^+$: 837.0657; found 837.0647; $^{11}B\{^1H\}$ NMR (22 °C, 128 MHz, $CDCl_3$): δ= 27.9 (br, 1B), 41.9 (br, 1B), 81.2 (br, 2B); ^1H NMR (22 °C, 400 MHz, $CDCl_3$): δ = 4.46, (br, BH_t), 2.03 (s, 30H, 2Cp*), -6.90 (br, 2H, Mo–H–B), -10.17 (br, 1H, W–H–B), -13.32 (br, 1H, Mo–H–W); $^{13}C\{^1H\}$ NMR (22 °C, 100 MHz, $CDCl_3$): δ =195.4, 189.4, 181.7 (CO), 98.9 ($\underline{C}_5Me_5$), 10.9 ($C_5\underline{Me}_5$); IR (hexane) ν/cm^{-1}: 2474 (BH_t), 2008(CO), 1937(CO), 1878 (CO)

8: HRMS (ESI$^+$): *m/z* calculated for $[C_{25}H_{36}B_4O_5Mo_2W_1+H]^+$: 841.0630; found 841.0623; $^{11}B\{^1H\}$ NMR (22 °C, 128 MHz, $CDCl_3$): δ= 103.0(br, 2B), 77.9 (br, 1B), 77.0 (br, 1B);^1H NMR (22 °C, 400 MHz, $CDCl_3$): δ = 5.87, (br, BH_t),1.93 (s, 30H, 2Cp*), -6.55 (br, 1H, Mo–H–B), -7.35 (br, 1H, Mo–H–B); $^{13}C\{^1H\}$ NMR (22 °C, 100 MHz, $[D_6]$-benzene): δ =198.6, 192.9 (CO), 108.5 (C_5Me_5), 12.9 (C_5Me_5); IR (hexane) ν/cm^{-1}: 2498 (BH_t), 2043 (CO), 2008(CO), 1969 (CO), 1890(CO).

Computational details. Geometry optimizations and electronic structure calculations were carried out on Gaussian09 (rev. C.01)[47] program package using BP86 functional[48] and def2-TZV[49] basis set from EMSL Basis Set Exchange Library. The 28 core electrons of molybdenum and tungsten were replaced by the quasi-relativistic effective core potential def2-ECP.[50] Vibrational analyses were carried out for all structures, and the absence of any imaginary frequency confirmed that all structures represent minima on the potential energy hypersurface. The NMR chemical shifts were calculated on the BP86/def2-TZVP optimized geometries using the hybrid Becke−Lee−Yang−Parr (B3LYP) functional and aforementioned basis set.[51] The NMR calculation utilized gauge-including atomic orbitals (GIAOs) method.[52] The ^{11}B NMR chemical shifts were calculated relative to B_2H_6 (B3LYP B shielding constant 84.23 ppm) and converted to the usual $[BF_3.OEt_2]$ scale using the experimental $\delta(^{11}B)$ value of B_2H_6, 16.6 ppm.

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Natural bonding analyses were performed with the natural bond orbital (NBO) partitioning scheme as implemented in the Gaussian 09 suite of programs.[53] Wiberg bond indexes (WBI) were obtained on natural bond orbital analysis.[54] In order to understand the nature of bonding of the synthesized molecules in greater detail, the topological properties of the resultant electron density, ρ, obtained from the wave functions of all the optimized structures were analyzed with the quantum theory of atoms in molecules (QTAIM).[55] The QTAIM analysis were carried out utilizing Multiwfn V.3.3.8 package[56] whereas the wave functions were generated with Gaussian09 at the same level of theory as was used for geometry.

X-ray Structure Determination. Crystal data for **4, 5** and **8** were collected and integrated using Bruker Kappa apexII CCD single crystal diffractometer and for **7** with OXFORD Diffraction SUPER NOVA CCD Diffractometer, equipped with graphite monochromated Mo-Kα (λ = 0.71073 Å) radiation. Data collection for **4** and **8** were carried out at 296 K and for **5** and **7** at 293 K, using ω-φ scan modes. Multi-scan absorption correction has been employed for the data using SADABS[57] program. The structures were solved by heavy atom methods using SHELXS-97 or SIR92[58] and refined using SHELXL-2014.[59] These data can be obtained free of charge from The Cambridge Crystallographic Data Centre via www.ccdc.cam.ac.uk/ data_request/cif

ASSOCIATED CONTENT

Supporting Information. The Supporting Information is available free of charge on the ACS Publications website.

Spectroscopic data for **4-5** and **7-8**; optimized geometries and MO diagrams (PDF).

CCDC 1878369 (**4**), 1878370 (**5**) 1878367 (**7**) and 1878368 (**8**) contain the supplementary crystallographic data for this paper. These data can be obtained free of charge via

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www.ccdc.cam.ac.uk/data_request/cif, or by emailing data_request@ccdc.cam.ac.uk, or by contacting The Cambridge Crystallographic Data Centre, 12 Union Road, Cambridge CB2 1EZ, UK; fax: +44 1223 336033

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Notes

The authors declare no competing financial interest.

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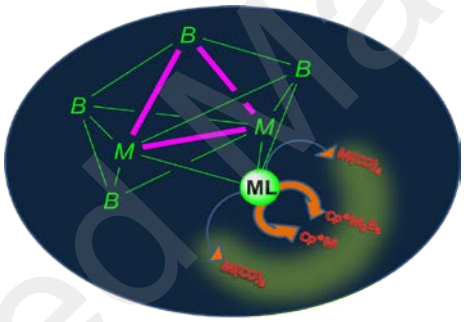
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Use of Single Metal Fragments for Cluster Building: Synthesis, Structure and Bonding of Heterometallaboranes

Using CO gas we have synthesized an extremely fluxional molybdenum diborane(4) species, $[\{Cp^*Mo(CO)_2\}_2\{\mu\text{-}\eta^2\text{:}\eta^2\text{-}B_2H_4\}]$ that mimic Cotton's dimolybdenum-acetylene complex, $[\{CpMo(CO)_2\}_2C_2H_2]$. In addition, we have isolated and structurally characterized a homoleptic molybdaborane, $[(Cp^*Mo)_3(\mu\text{-}H)_2(\mu_3\text{-}H)(\mu\text{-}CO)_2B_4H_4]$, and a vertex-fused cluster, $[(Cp^*W)_3WB_9H_{18}]$ (see picture).



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