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Rational synthesis and dimensionality tuning of MOFs from preorganized heterometallic molecular complexes

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Rational design of a series of new heterometallic MOFs was carried out by the judicious choice of corresponding pivalate complexes $[\text{Li}_2\text{M}_2(\text{piv})_6(\text{py})_2]$ ($\text{M} = \text{Zn}^{2+}, \text{Co}^{2+}$, $\text{piv}^- = \text{pivalate anion}$ and $\text{py} = \text{pyridine}$) as a source of secondary building units, $\{\text{LiM}(\text{O}_2\text{CR})_3\}$ and organic tricarboxylate linker as a node defining the dimensionality of the framework by the orientation of carboxylic group in or out of the central aromatic ring plane. Thus trimesate (btc^{3-}) linker results in 3D **srs** topology frameworks with intersecting system or isolated channels, 1,3,5-benzenetribenzoate (btb^{3-}) results in layered **hcb** isostructural compounds additionally stabilized with H- π interactions between layers. The layered compounds demonstrate a permanent porosity with BET surface area up to $688 \text{ m}^2 \text{ g}^{-1}$ with the possibility of selective gas adsorption (CO_2 over N_2 and CH_4). Zn-based coordination polymers show a notable color changes and drastic (up to 30 times) quenching of the luminescence upon inclusion of different nitroaromatics.

Introduction

Gas storage [1–6], separation [7–9], heterogeneous catalysis [10–11], chemical sensing [12–13], optoelectronics (ferroelectricity, non-linear optics, and LEDs) [14], drug delivery [15] – functional materials based on metal-organic frameworks (MOFs) find their places in each of these, as well as in many other areas. The most relevant and functionally important MOFs, such as HKUST-1 [16–17], IRMOFs [18–22], UiOs [23–24] and MILs [25–28], are based on polynuclear secondary building units (SBUs) self-assembled from metal ions at particular reaction conditions. In certain cases, these SBUs are known as discrete polynuclear complexes which provide possibility to facilitate the assembly of periodic coordination structures through a ligand exchange on those pre-organized complexes. This rationalizes the engineering of new compounds with predictable structures and suppresses the crystallization of impurities through competing self-assembly routes. More importantly, such approach could often be the only way to synthesize MOFs based on complex heterometallic SBUs where self-assembly route is very hard to

control and optimize [29]. Nevertheless, the ligand substitution on pre-organized SBUs is only demonstrated in a limited number of examples of synthesis of MOFs, therefore a further development of preparation methods of coordination polymers from pre-organized complexes is an important direction in the MOF area.

This work demonstrates the preparation of a series of porous coordination networks from tetranuclear complexes $[\text{Li}_2\text{M}_2(\text{piv})_6]$ ($\text{M} = \text{Co}^{2+}, \text{Zn}^{2+}$; $\text{piv}^- = \text{pivalate}$) which could be regarded as a dimer of two binuclear carboxylate units $\{\text{LiM}(\text{RCOO})_3\}$ with 3-connected nature. An extension of the $\{\text{LiM}(\text{RCOO})_3\}$ units through tricarboxylic ligands such as trimesate (btc^{3-}) or 1,3,5-benzenetribenzoate (btb^{3-}) predictably leads to the formation of MOFs with 3-connected topologies. The obtained MOFs feature selective gas adsorption properties, luminescence quenching upon an inclusion of different nitroaromatic molecules as well as the separation of benzene/cyclohexane mixture in a liquid phase.

Results and discussion

Material synthesis and structures.

Tetranuclear coordination complex $[\text{Li}_2\text{Co}_2(\text{piv})_6(\text{py})_2]$ (**1**) was synthesized by a reaction of lithium pivalate salt $\text{Li}(\text{piv})$, polymeric amorphous $[\text{Co}(\text{piv})_2]_n$ and pyridine in equimolar ratio in a mixture of CH_3CN and THF. This synthetic method is similar to that for previously reported $[\text{Li}_2\text{M}_2(\text{piv})_6]$ ($\text{M} = \text{Ni}, \text{Co}, \text{Zn}$) complexes [30–32].

The compound **1** crystallizes in triclinic space group $P\bar{1}$ (Fig. 1). The center of symmetry of the tetranuclear molecule is at the intersection of the diagonals of the almost square $[\text{Li}_2(\mu_2-$

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O₂) moiety (Figure 1), and the Li₂Co₂ core has a planar structure (Co...Li 3.098(5) Å, Li...Li 2.665(9) Å, Co–Li–Li 121.8(3)°).

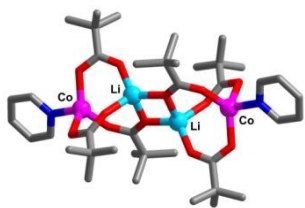
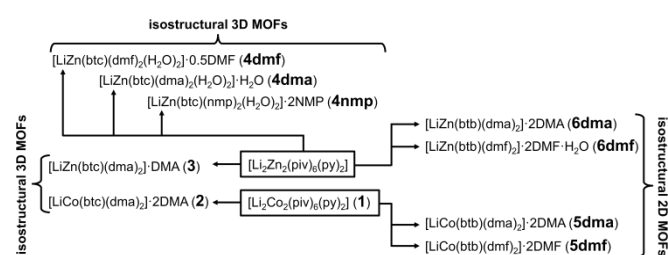


Figure 1. Molecular structure of [Li₂Co₂(piv)₆(py)₂].

Each Co²⁺ ion coordinates N atom of a pyridine ligand and three O atoms of three carboxylate groups (Co–N 2.067(2), Co–O 1.930(2)–1.952(2) Å), thus adopting a distorted tetrahedral coordination (CoNO₃). The coordination environment of the Li⁺ ions (LiO₄) is tetrahedral (Li–O 1.889(5)–1.985(5) Å). Three carboxylate bridges link the Co²⁺ and Li⁺ ions with two carboxylates acting as bridges between two heterometallic centers, and one acting as a μ₃-bidentate linker bridging two Li⁺ ions through one O atom. The six carboxylate anions are directed perpendicularly to the Co–Li axis of the node, while Co–N bonds are directed along this axis. In general complex is isostructural to previously reported [Li₂Zn₂(piv)₆(py)₂] [32]. The only difference is that one μ₂-bidentate oxygen atom in carboxylate group of pivalate anion connected to lithium cation. This atom is placed closer to Co cation forming weak interaction 2.590(7) Å unlike 2.749(2) Å in case of Li₂Zn₂ complex. Nevertheless such differences do not affect strongly on the overall similarity.

The tetranuclear complex [Li₂M₂(piv)₆] could be viewed as a dimer of 3-connecting dinuclear units [LiM(piv)₃] fused *via* lateral Li–O coordination interactions. The substitution of pivalate ligands to polycarboxylic ligands in such dinuclear units [LiM(piv)₃] will result in a formation of 3-connected coordination polymers based on the {LiM} nodes. It must be pointed out that the tricarboxylate dinuclear complexes [M₂(O₂CR)₃] are very rare, compared to the well-known and relatively abundant tetracarboxylate analogues [M₂(O₂CR)₄]. Therefore, the [Li₂M₂(piv)₆] complexes could be regarded as valuable source of heterometallic tricarboxylate dinuclear SBUs, previously observed only in a handful of MOF structures. The following nine MOFs are obtained from the [Li₂M₂(piv)₆] (M = Co²⁺, Zn²⁺) and are solely based on such rare [LiM(O₂CR)₃] SBUs (Scheme 1).



Scheme 1. General scheme of MOFs obtained from molecular complexes [Li₂M₂(piv)₆(py)₂] (M = Co, Zn).

Compound [LiCo(btcb)(dma)₂]₂·2DMA (**2**) was synthesized by a heating of **1** with 1,3,5-benzenetricarboxylic acid in DMA. **2** crystallizes in tetragonal chiral space group *P*4₃2₁2. In the structure Li and Co atoms are in the same positions with 50% occupancy. Each Li/Co ion is bounded to the oxygen of DMA molecule and to three O-donors of three carboxylate groups, adopting a distorted tetrahedral coordination (Li/Co–O 1.916(2)–1.961(2) Å). Li and Co cations are connected by three bridging carboxylate groups forming a rare example of a tricarboxylate bimetallic paddle wheel fragment (Fig. 2).

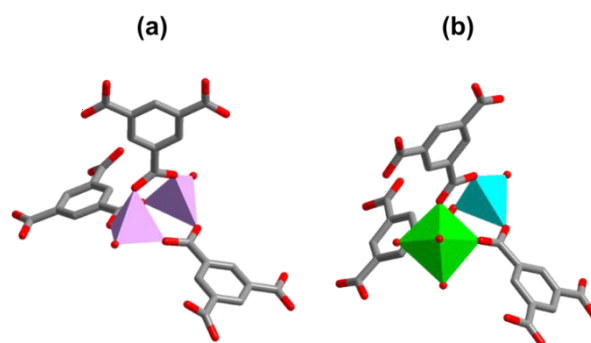


Figure 2. View of node in **2** and proposed for **4** (from left to the right). (hydrogen atoms are omitted; Li/Co cations are presented as lilac tetrahedra, Li – as cyan, Zn – as green).

These building blocks are connected with six others by benzenetricarboxylate linkers into 3D porous structure of chiral Si net in SrSi₂ topology (*srs*) (Fig. 3a, S6). In a projection of the crystal structure perpendicular to *b* axis one can identify a 1D channels about 7 Å in diameter which are occupied by two coordinated and two guest DMA solvent molecules per formula unit (Fig. 3b). All solvent molecules were identified from the single-crystal X-ray diffraction data. The phase purity and chemical composition of **2** was confirmed by a number of methods and analyses, including AAS which supported the equimolar ratio for Co²⁺ and Li⁺ ions for this heterometallic compound. It should be noted that the synthesis and structural investigation of the chiral antipode (enantiomer) of the compound **2** was published recently [33], however, the reported here X-ray crystallography data quality is much better, allowing the refinement of the DMA guest molecules in the channels.

The total accessible for solvent volume is 42% according to PLATON [34] calculated for structure without guest DMA molecules. The thermogravimetric (TG) data shows steady 54 % weight decrease upon heating to 400 °C corresponding to removal of all solvent molecules, followed by the framework decomposition (Fig. S16).

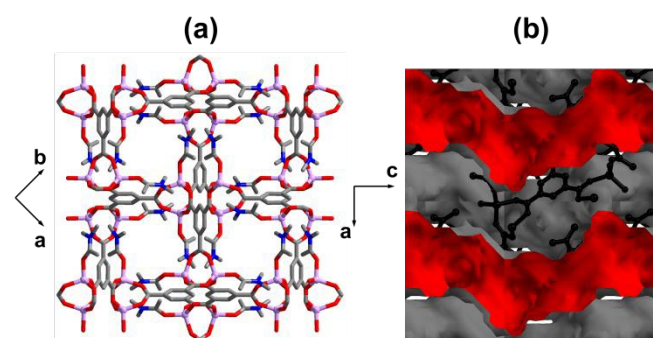


Figure 3. Structure projection of **2** perpendicular to *c* (*a*) and *b* (*b*) axis (the inner surface of channels is presented as red; Li/Co cations are presented as lilac balls).

Compound $[\text{LiZn}(\text{btc})(\text{dma})_2]\cdot\text{DMA}$ (**3**) could be obtained in a solvothermal reaction of $[\text{Li}_2\text{Zn}_2(\text{piv})_6(\text{py})_2]$ with H_3btc in DMA. Only a few crystals of the product were isolated as a minor impurity to the dominating phase **4dma** (see below). The compound **3** is totally isostructural to **2**. Like in previous case, both Li(I) and Zn(II) cations are disordered over two positions with 50% occupancy, thus they have similar coordination geometry. Each Li/Zn cation is bounded to the oxygen atom of DMA molecule and to three O-donors of three carboxylate groups, thus adopting a distorted tetrahedral coordination (Li/Zn–O 1.897(3)–1.989(2) Å). The total accessible for solvent volume is 42%.

The heating of $[\text{Li}_2\text{Zn}_2(\text{piv})_6(\text{py})_2]$ complex with benzenetricarboxylic acid in either DMF, DMA or NMP resulted in a precipitation of the compounds $[\text{LiZn}(\text{btc})(\text{dmf})_2(\text{H}_2\text{O})_2]\cdot 0.5\text{DMF}$ (**4dmf**), $[\text{LiZn}(\text{btc})(\text{dma})_2(\text{H}_2\text{O})_2]\cdot\text{H}_2\text{O}$ (**4dma**), and $[\text{LiZn}(\text{btc})(\text{nmp})_2(\text{H}_2\text{O})_2]\cdot 2\text{NMP}$ (**4nmp**). Note, that the heating of a very similar $[\text{Li}_2\text{Co}_2(\text{piv})_6(\text{py})_2]$ (**1**) produces **2** in DMA, while the other media (DMF, NMP) affords only amorphous phases. We were not able to identify structures **4** by the single-crystal X-ray diffraction method, but the powder X-ray data clearly suggested that all compounds **4** are isostructural to the known MOFs $[\text{Zn}_2(\text{btc})(\text{H}_2\text{O})(\text{C}_2\text{H}_5\text{OH})_2(\text{NO}_3)]$ [35] or $[\text{CoZn}(\text{btc})(\text{dmf})_3(\text{NO}_3)]$ [36]. The Le Bail fit was carried out to identify the space group and to calculate the unit cell parameters from the PXRD data (Fig. S13–15). For **4dmf** found unit cell is $P2_13$ / $a = 14.356(2)$ Å; **4dma** – $P2_13$ / $a = 14.338(3)$ Å; **4nmp** – $P2_13$ / $a = 14.494(4)$ Å which is very close to both $[\text{Zn}_2(\text{btc})(\text{H}_2\text{O})(\text{C}_2\text{H}_5\text{OH})_2(\text{NO}_3)]$ – $P2_13$ / $a = 14.728(4)$ Å and $[\text{CoZn}(\text{btc})(\text{dmf})_3(\text{NO}_3)]$ – $P2_13$ / $a = 14.1520(3)$ Å. In the known compound $[\text{CoZn}(\text{btc})(\text{dmf})_3(\text{NO}_3)]$ tetrahedral Zn(II) ion is filled by a nitrate to maintain charge neutrality, while the Co(II) is in octahedral environment due to the coordination of three solvent molecules. Taking into account the literature data as well as the chemical analyses including AAS spectrometry to determine Li and Zn concentration, the compounds **4** are supposedly built from the dinuclear heterometallic $\{\text{LiZn}(\text{RCOO})_3\}$ tricarboxylate units, already seen in **3**. The Li^+ cations are coordinated by four atoms: three carboxylate oxygens and one solvent molecule. The coordination environment of Zn^{2+} in **4** is believed to be octahedral with three carboxylate anions, one solvent molecule and two H_2O

molecules according to microelemental analysis. The dinuclear tricarboxylate SBUs $\{\text{LiZn}(\text{RCOO})_3\}$ are then supposedly connected by trimesate anions into 3-connected network with **srs** topology, similar to that in **2** or **3**. Assuming the structural similarity to the known compounds, the pore system in **4** should contain a 3-D structure with nearly 44% of the framework represented by an extended channel system having a 14 Å cross-section filled with coordinated and free solvent molecules [35]. The final formulas for the series of isostructural compounds (Fig. 4c) **4** were established on the basis of chemical analysis as well as TG data. According to TG analyses, the **4dmf** loses 44% of weight with removal of all solvent molecules upon heating up to 400°C (43.9% calculated from empirical formula), **4dma** TG weight loss is 43.2% (45% calculated from empirical formula) and **4nmp** TG weight loss is 60.4% (60.8% from empirical formula) (Fig. S16). The obtained crystalline MOFs **2–4** with demonstrated **srs** network topology are isoreticular to previously known series of homo- and heterometallic trimesates $[\text{MM}'(\text{btc})](\text{An})$, where $\text{M}, \text{M}' = \text{Zn}^{2+}, \text{Co}^{2+}, \text{Ni}^{2+}$, $\text{An} = \text{anion}$ [35–38]. The main difference is that the frameworks **2–4** are charge-neutral in contrast with the cationic ones reported in the literature. This should be considered as a potential advantage since the absence of any counter-ions expands the guest accessible volume for a porous structure.

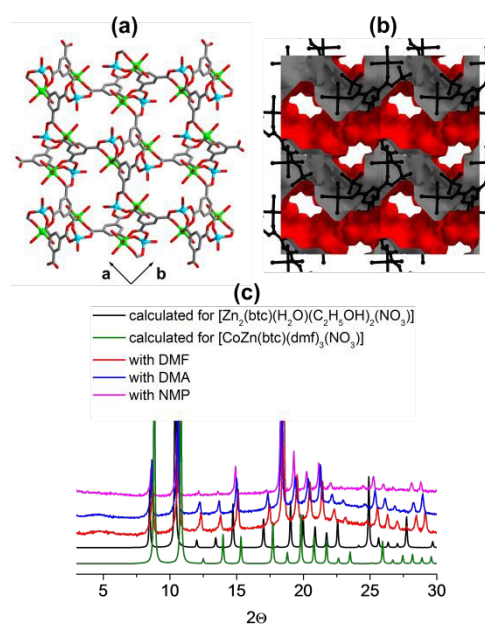


Figure 4. Proposed structure projection of **4** perpendicular to *c* axis (*a, b*) (Li cations are presented as blue balls, Zn – as green; calculated pore surface is presented as red - *b*). Comparison of PXRD data for known trimesates $[\text{Zn}_2(\text{btc})(\text{H}_2\text{O})(\text{C}_2\text{H}_5\text{OH})_2(\text{NO}_3)]$ [35] or $[\text{CoZn}(\text{btc})(\text{dmf})_3(\text{NO}_3)]$ [36] and for **4** compounds (*c*).

Single crystals of $[\text{LiCo}(\text{btb})(\text{dma})_2]\cdot\text{DMA}$ (**5dma**) and $[\text{LiZn}(\text{btb})(\text{dmf})_2]\cdot 2\text{DMF}\cdot\text{H}_2\text{O}$ (**6dmf**) were synthesized by heating of $[\text{Li}_2\text{M}_2(\text{piv})_6(\text{py})_2]$ ($\text{M} = \text{Co}, \text{Zn}$) complexes, respectively, with 1,3,5-tris(4-carboxyphenyl)benzene (H_3btb) in either DMA/ CH_3CN or DMF/ CH_3CN . All these compounds appear to be isostructural as evidenced by powder X-ray diffraction data (Figs. S9, S10). **5dma** crystallizes in monoclinic

space group $I2/a$. In this structure Li and Co cations are disordering over two positions with 50% occupancy, similarly to **2** and **3**. Each Li/Co ion coordinates the O atom of DMA molecule and three O atoms of three carboxylate groups, thus adopting a distorted tetrahedral coordination (Li/Co–O 1.900(2)–2.016(3) Å) (Fig. 5).

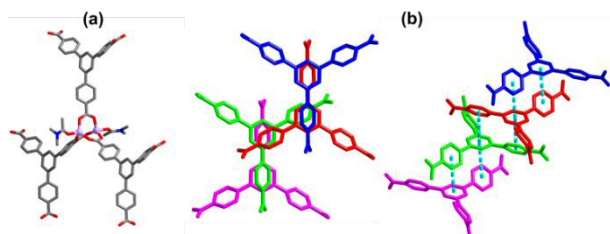


Figure 5. View of node in **5dma** (a); hydrogen atoms are omitted; Li/Co cations are presented as lilac balls. Stabilizing interactions between benzene rings of the btb-linker of four independent layers of **5dma** in different projections (b). Short contacts are shown as cyan dashed lines.

The {CoLi} dinuclear tricarboxylate moieties are connected by benzenetribenzoate (btb^{3-}) linkers into a 3-connected layered network with honeycomb topology **hcb**. Four independent layers are stacked atop each other with slight shift forming a 4-periodic ABCD packing (Fig. 6). The packing of the layer is stabilized by weak C–H $\cdots\pi$ interactions between benzene rings of btb^{3-} linkers of alternating layers (Fig. 5b). In a projection of the crystal packing perpendicular to the c axis, a 1D channels about 6 Å in diameter, occupied by highly disordered solvent molecules could be identified. The PLATON/SQUEEZE procedure allowed us to calculate the contribution to the diffraction from the solvent region and produced a set of solvent-free diffraction intensities [39]. The final formula of **5dma**, $[\text{LiCo}(\text{btb})(\text{dma})_2] \cdot 2\text{DMA}$, was derived from the SQUEEZE results (188e per unit cell – one DMA guest molecule per formula unit) which is in acceptable agreement with elemental and thermogravimetric analyses (which gives 2 guest DMA molecules per formula unit). The solvent volume of the channels was calculated to be 25.8%. The TGA data shows steady 39% weight decrease upon heating to 400 °C corresponding to removal of all solvent molecules (40.9% calculated from empirical formula), followed by the framework decomposition (Fig. S16).

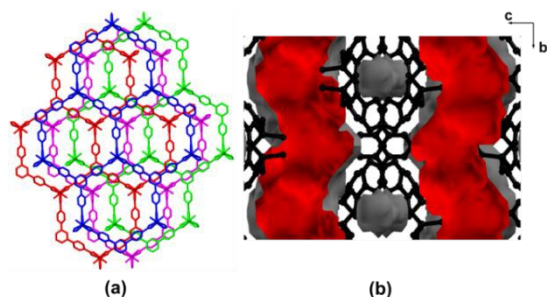


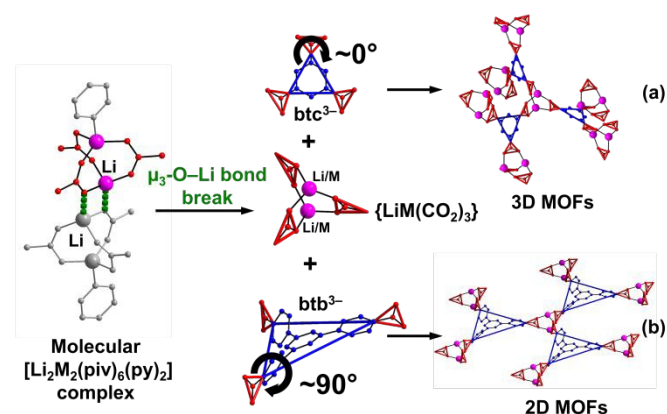
Figure 6. Independent layers in **5** (a), and channels view perpendicular to a axis (b).

Except for the difference in the solvent molecules, **6dmf** is isostructural to **5dma**. Consequently, **6dmf** is based on the bimetallic tricarboxylate paddlewheel {LiZn} unit both Li and Zn

cations merged between two positions. The Li/Zn–O metal-ligand distances are in the range 1.898(3)–2.016(3) Å. The overall crystal structure **6** is formed by stacking of honeycomb layers (**hcb**) with the system of channels almost identical to that in **5** (Fig. S5). The final formula of **6dmf** was derived from the SQUEEZE results (168e per unit cell) with elemental analysis and TGA (2 guest DMF and one guest water molecule per formula unit). The thermogravimetric data shows steady 41% weight decrease upon heating to 400 °C corresponding to removal of both guest and coordinated solvent molecules (37.9% calculated from empirical formula), followed by the framework decomposition (Fig. S16). Similar compounds **5dmf** and **6dma** could be also precipitated from the corresponding solvents. According to the PXRD data these are fully isostructural to **5dma** and/or **6dmf** (Fig. S9, S10).

It should be noted that all the reported compounds **2–6** were prepared from the complexes $[\text{Li}_2\text{M}_2(\text{piv})_6]$, which tetranuclear core is composed of two fused dinuclear units $\{\text{LiM}(\text{piv})_3\}$ via lateral Li–O coordination interactions. In previous examples we showed the preservation of the tetranuclear core during the MOF assembly through linear ditopic ligands [32,40]. Apparently, the polycondensation in the presence of 3-connected carboxylic ligands facilitates the dichotomy of the tetranuclear complex to dinuclear units and the formation of either 2D or 3D 3-connected coordination networks, or only presence of N-donor ditopic ligands can stabilize the tetranuclear core [40]. Remarkably, the honeycomb (**hcb**) and SrSi_2 (**srs**) nets are the most regular topologies among the 2D or 3D 3-connected nets, respectively [41]. Such regular nets are especially interesting and important for the crystal engineering since self-assembly of periodic structures usually follows the most symmetrical and simple way. For example, diamond net and primitive cubic net are the overwhelmingly dominating topologies among known 3D structures, being the most regular network types for 4- and 6-connecting nets, respectively [42].

Interestingly, the key factor in dimensionality of the coordination networks in the obtained products lies in the geometry and flexibility of the organic ligand (Scheme 2).



Scheme 2. General representation of {LiM} building block formation and construction of coordination polymers with different dimensionality depending on the position of carboxylic group in plane of the ligand (a) forming a 3D framework or out of the plane (b) forming **hcb** layered structure.

Indeed, a rather rigid structure of the tricarboxylate unit $\{\text{LiM}(\text{O}_2\text{CR})_3\}$ fixes the planes of the carboxylate groups along the LiM axis, which is perpendicular to the plane of this 3-connected node. The planar 2D honeycomb net could only be assembled if the organic ligand allows the same orientation of the carboxylate groups, perpendicular to the plane of the organic ligand itself. The aromatic rings of the btb^{3-} anion are known to be tilted due to $\text{H}\cdots\text{H}$ repulsion while the trimesate (btc^{3-}) anion is almost perfectly planar with all three carboxylate groups lying in the same plane. Therefore, the connection of $\{\text{LiM}(\text{O}_2\text{CR})_3\}$ units through btc^{3-} necessarily directs the alignment of the 3-connected nodes out of the plane forming a 3D structure. Contrary, more flexible btb^{3-} anion allows the same orientation of $\{\text{LiM}(\text{O}_2\text{CR})_3\}$ units within the plane; thus, the formation of honeycomb 2D layers becomes possible. Therefore, the crystallization of 2D **hcb** or 3D **srs** 3-connected nets by the reaction of $[\text{Li}_2\text{M}_2(\text{piv})_6]$ complexes with btb^{3-} or btc^{3-} ligands, respectively, is to be expected. Even though we did not anticipate the outcome, the above-mentioned retro-synthetic analysis of the results rationalizes the underlying relations between the reaction conditions and the crystal structure of the products. The synthesis of the compounds **2-6** represents an example of a successful design of coordination polymers with regular 3-connected topologies based on rare tricarboxylate secondary building blocks.

Gas adsorption and selectivity

As far as **2** demonstrates low stability upon the guest removal (Figure S7) and no gas adsorption the compound $[\text{LiZn}(\text{btc})(\text{dmf})_2(\text{H}_2\text{O})_2] \cdot 0.5\text{DMF}$ (**4dmf**) was selected to study gas adsorption properties of isorecticular 4 3D nets with **srs** topology. The liquid-phase replacement of DMF solvent molecules with volatile acetone, followed by a low-pressure activation at 80 °C was carried out. Despite relatively mild activation conditions, the powder XRD data (Fig. S8) indicate amorphization of the porous material. Accordingly, the remaining solid demonstrates poor CO_2 adsorption *ca.* 40 $\text{mL}\cdot\text{g}^{-1}$ at 1 bar, 195 K, which is an order of magnitude lower than the expected value.

The gas adsorption properties of isostructural 2D MOFs with **hcb** topology were also investigated. Despite their layered nature these compounds turned out to be stable upon the guest molecules removal (Figure S12). The characterization of texture and porosity of the activated material $[\text{LiCo}(\text{btb})(\text{dmf})_2]$ **5dmf** was carried out by measurements of N_2 and CO_2 adsorption isotherms at 77 K and 195 K, respectively. Although the observed nitrogen adsorption was quite low (19.4 mL/g) the CO_2 capacity was rather reasonable reaching 106 mL/g (Figure 7a). The measured pore volume is 0.153 $\text{cm}^3\cdot\text{g}^{-1}$, which is in a good agreement with the calculated porosity of the framework **5dmf** based on the structural data, assuming the complete removal of guest DMF molecules. Similarly, the porosity of the activated **5dma** and **6dmf** was also characterized (Table 2, Figure S19). We tried to find out the nature of such selective CO_2 adsorption. Some indirect conclusions could be drawn from the careful analysis of the

pore surfaces in the structures **5dma** and **6dmf** (all compounds **5** and **6** are isostructural but have different coordinated solvent molecules). As could be seen in the projection of the structures along the channels (Figure S5), a coordinated DMF (left) or DMA (right) molecules could clearly be identified on the surface of the channel. Such geometry potentially allows the intermolecular interactions between the CO_2 molecules and the coordinated solvent (DMF, DMA) molecules.

Table 2 Calculated parameters of porous structure for activated **5** and **6** compounds for different gases and temperatures.

No	Gas@T	Specific surface area $\text{m}^2\cdot\text{g}^{-1}$			$V_{\text{pore}} / \text{cm}^3\cdot\text{g}^{-1}$		$V_{\text{ads}}^c / \text{cm}^3(\text{STP})\cdot\text{g}^{-1}$
		Langmuir	BET	DFT	Total ^a	DFT ^b	
5dma	N_2 @77 K	—	4.8	—	—	—	4.5
	CO_2 @195 K	509.9	426.3	224.6	0.230	0.157	108.6
5dmf	N_2 @77 K	69.5	63.4	62.2	0.030	0.028	19.4
	CO_2 @195 K	509.1	419.4	224.3	0.224	0.153	106.0
5dmf^d	CO_2 @195 K	318.4	265.7	143.2	0.157	0.107	74.0
	N_2 @77 K	125.3	116.9	133.3	0.071	0.075	45.6
6dmf	CO_2 @195 K	768.5	687.7	820.3	0.394	0.251	185.1

^a measured at $P/P_0 = 0.95$;

^b micropore volume;

^c not applicable to the isotherm or have not been calculated.

^d sample was additionally vacuumated at 120°C for 4 hours

The calculated BET surface areas were found to be 419, 426 and 688 m^2/g for the activated compounds **5dmf**, **5dma** and **6dmf**, respectively. Attempts of the activation of **5dmf** at higher temperature in order to remove the coordinated DMF solvent only led to the gradual decomposition of the porous structure as evidenced by the decrease of the gas adsorption properties and PXRD data (Figure 7a, Figure S12).

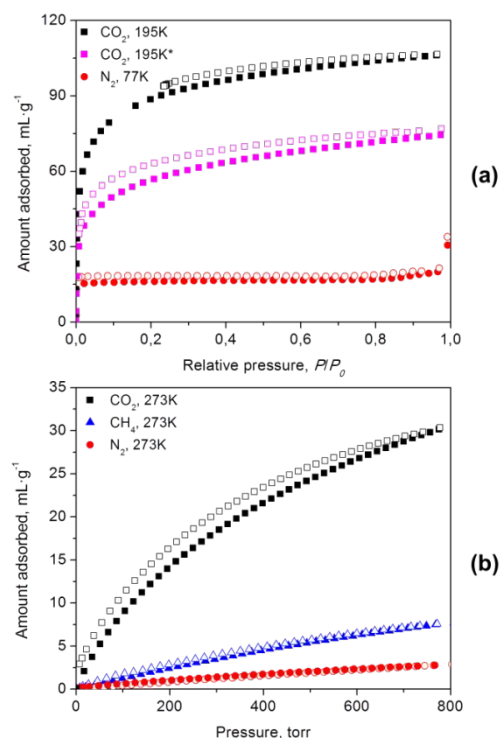


Figure 7. Adsorption–desorption isotherms of N_2 at 77 K and CO_2 at 195 K by activated **5dmf** (a), * – sample was additionally evacuated at 393 K for 4 hours. Adsorption–desorption isotherms of CO_2 , CH_4 and N_2 at 273 K by activated **6dmf** (b).

The activated **5dmf** was selected as prototype of the title layered compounds for the detailed gas adsorption measurements and selectivity calculations. The adsorption

isotherms of N₂, CH₄ and CO₂ was measured at 273 K (Figure 7b) and 298 K (Figure S20, Table S1). The selectivity factors (*SF*) of the adsorption of CO₂ over either N₂ or CH₄ were estimated by three commonly used methods (see ESI): i) as a ratio of adsorbed volumes; ii) as a ratio of Henry's constants; iii) by ideal adsorbed solution theory (IAST). Table 3 shows the results of the selectivity factors calculations for **5dmf** at 273 K for equimolar gas mixture. The compound demonstrates moderate selectivity in case of CO₂/CH₄ (Table 3 at 273 K, Table S5 at 298 K) and rather high selectivity in case of CO₂/N₂ gas mixtures. According to IAST, the highest *SF* values were *SF* = 8.3 and 70.5 for CO₂/CH₄ and CO₂/N₂ equimolar gas mixtures, respectively. Also, the IAST calculations (see ESI) allowed us to obtain the selectivity factors as a function of the gas mixture compositions or the absolute pressure for CO₂/N₂ gas mixture at 273 K (Figure S21) as well as for CO₂/CH₄ mixture at both 273 K and 298 K (Figure S22). These data demonstrate a good potential of the obtained layered MOF compounds for the gas separation applications.

Table 3 Calculated adsorption selectivity factors on **5dmf** for equimolar binary gas mixtures at 273 K.

Gas mixture	V ₂ /V ₁	K _{h2} /K _{h1}	IAST
CO ₂ /N ₂	10.9	31.4	70.5
CO ₂ /CH ₄	4.0	9.5	8.3

Benzene/cyclohexane separation

Cyclohexane is a valuable chemical utilized in synthesis of polymers. It is obtained exclusively by a hydrogenation of benzene. The effective cyclohexane/benzene separation is one of the most important industrial processes to be optimized since a conventional vapor fractional separation is hampered by a similarity of the boiling points of these compounds [43]. A great potential of porous MOFs for the cyclohexane/benzene separation has been already demonstrated a number of times [44–46]. The small aperture of channels in **6**, lined with aromatic functionalities can facilitate benzene extraction from the benzene/cyclohexane mixtures due to possible π - π stacking interactions. The as synthesized **6dmf** was immersed into the equimolar benzene/cyclohexane mixture at room temperature for 24 h. After that the crystalline phase was filtered and treated with DMF for 24 h to extract the solvent molecules. The ¹H NMR data showed that the composition of desorbed benzene/cyclohexane mixture was 3:1 in favor of the aromatic benzene, supporting our idea (Figure S23). Total amount of adsorbed benzene corresponds to nearly 1 benzene molecule per formula unit.

Guest-Exchange and Photoluminescence Studies.

Due to the benzenetricarboxylate and benzenetribenzoate aromatic ligands, the title frameworks **4dmf** and **6dmf** exhibit photoluminescent (PL) properties. The dependence of such properties upon the presence of a certain guest molecules in the channels of porous MOFs have been exploited in sensor applications a number of times already [12,47]. In this work porous **4dmf** and **6dmf** were selected to investigate their guest-dependent luminescence towards various nitroaromatic molecules, such as nitrobenzene (NB), 4-nitro-*m*-xylene (NX)

and 2,6-dimethylnitrobenzene (DMNB). The guest (solvent) molecules were substituted by immersing of the as-synthesized crystals into a nitroaromatic compound for 2 days. According to powder XRD data, the crystals **4dmf** appeared to be unstable in NX and DMNB and moderately stable in NB (Figure S8) while **6dmf** is stable in all three aromatic solvents (Figure S11). The superior stability of layered **hcb** compound **6dmf** over 3D **srs** framework **4dmf** supports the observations during the gas adsorption measurements.

The PL spectra of the as-synthesized **4dmf** has maximum emission at $\lambda_{em} = 370$ nm ($\lambda_{ex} = 300$ nm, Fig. S24) with rather low quantum yield $\varphi \approx 1\%$. Because of complete electron shell configurations of both Zn²⁺ and Li⁺, the observed emission should apparently be attributed to the intraligand $\pi^* \rightarrow \pi$ electron transfer of the btc³⁻ anion, similar to the other reported examples [48-51]. The inclusion of the nitrobenzene shifts the emission maximum to $\lambda_{em} = 490$ nm and significantly reduces the luminescence of **4dmf-NB** about 18 times. Such luminescence quenching is usually attributed to a charge-transfer mechanism from the excited btc³⁻ ligands to the electron-deficient aromatic system of guest NB [12, 40, 52].

The PL spectra of the as-synthesized **6dmf** has maximum emission at $\lambda_{em} = 350$ nm ($\lambda_{ex} = 300$ nm, Fig. 9), similarly to known examples [53]. Like in previous case, the PL mechanism should be assigned solely to the intraligand $\pi^* \rightarrow \pi$ transition in the btb³⁻ anion. Due to the polyaromatic nature of the organic ligand, the emission of **6dmf** is far more intense and efficient than in **4dmf**, with $\varphi \approx 30\%$. The substitution of solvent DMF to either NB, NX or DMNB results in a reduction of the luminescence intensity. The spectra of the PL emission of the guest-exchanged samples **6dmf-G** feature bathochromic shift of the maxima ($\lambda_{em} = 515$ – 540 nm). Such UV-induced color changes clearly visualize the presence of the nitroaromatic species (Fig. 8). The guest-dependent luminescence demonstrated here for **4dmf** and **6dmf** is an important demonstration of the principle of a detection of different nitroaromatic species.

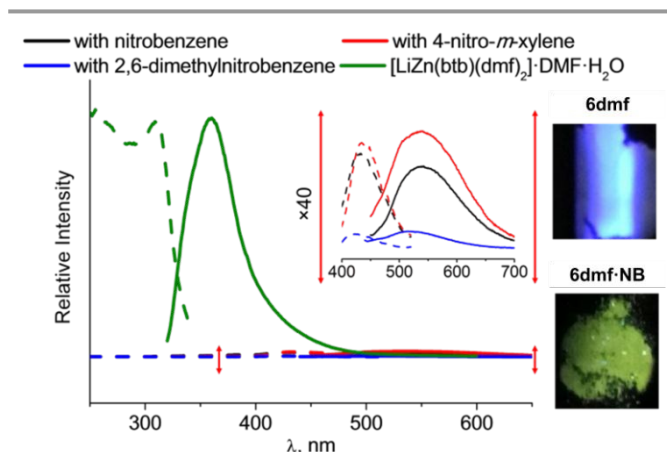


Figure 8. Spectra of emission (solid) and excitation (dashed) of **6dmf** and inclusion compounds ($\lambda_{ex} = 300$ nm, spectral slit width = 1.8 nm for **6dmf**; 420 nm, 5 nm for guest-exchanged samples). Microphotographs of luminescent species of **6dmf** as synthesized and **6dmf-NB**.

Experimental

Materials and Methods.

Commercial reagents (pyridine (>99%)) and solvents (acetonitrile (99%), tetrahydrofuran (THF, 99.9%)) were used without additional purification. The starting complexes $[\text{Co}(\text{piv})_2]_n$ and $\text{Li}(\text{piv})$ were synthesized according to procedures described previously [54,55]. The attenuated total reflection infrared (ATR-IR) spectra of **1** were recorded in the range 4000–400 cm^{-1} on a Perkin Elmer Spectrum 65 instrument. Elemental analysis of **1** was carried out with a Carlo Erba automatic C,H,N,S-analyzer. Unless specified, all reagents for synthesis of **2–6** were obtained from commercial sources and were used as received. Solvents acetone (99.99%), *N,N*-dimethylacetamide (DMA, 99%) and *N*-methyl-2-pyrrolidone (NMP, 99%) were used without additional purification. *N,N*-dimethylformamide (DMF, 99%) was treated over activated molecular sieves 3 Å. Powder X-ray diffraction data were obtained on a Shimadzu XRD 7000S diffractometer (Cu K α radiation). The following parameters were used: 2θ step = 0.03°, counting time = 1.0–2.5 s, 2θ scan range = 3–30°. For compounds **2–6**, the powder X-ray diffraction patterns are present in Figures S7–11, ESI. Fourier-transform infrared (FT-IR) spectra in the range from 4000–300 cm^{-1} were measured on a Vertex 80 spectrometer. The IR spectra are present in Figure S17, Supplementary Information. Thermogravimetric analyses (TGA) were obtained on a NETZSCH TG 209 F1. The sample quantity ranged from 2 to 10 mg. All samples were heated under helium atmosphere from room temperature up to 550 °C at a 10 °C·min⁻¹ heating rate. Elemental C, H, N analysis data for **2–6** were obtained on a Eurovector 600 analyzer. The excitation and emission spectra of the solid samples were recorded on a fluorescence spectrophotometer Cary Eclipse (Varian). The emission spectra of the complexes were recorded at room temperature under the following experimental conditions: λ_{ex} = 300 and 420 nm, V = 500 V, spectral slit width = 1.8 – 5 mm. The photoluminescent (PL) spectra were registered for the samples prepared by grinding compounds to powder between two quartz glasses. The thin layer of powder between glasses was placed at 45° to the excitation light beam. The xenon flash lamp was used as a light source to excite the steady-state PL spectra. ¹H NMR spectra were recorded on a Bruker Advance 500 spectrometer and referenced to residual protons of DMF-d7 solvent. The atomic absorption spectroscopy data were collected using AA spectrophotometer Z-8000 (AAS analysis) for **2**, **4**, **5** and **6** compounds to define Li:Zn, Li:Co ratios. Samples were dissolved in acidified water.

Synthetic procedures.

Synthesis of $[\text{Li}_2\text{Co}_2(\text{piv})_6(\text{py})_2]$ (1**).** A mixture of $[\text{Co}(\text{piv})_2]_n$ (0.30 g, 1.15 mmol) and $\text{Li}(\text{piv})$ (0.124 g, 1.15 mmol) was dissolved in a mixture of CH_3CN (30 mL) and THF (10 mL). The reaction mixture after addition of pyridine (0.111 mL, 1.38 mmol) was refluxed and heated at 50°C for 60 min. The resulting blue colour solution was filtered and evaporated up to 20 mL volume and cooled to room temperature. THF (5 mL) was added to a blue solution and kept at room temperature for 24

hours. Formed blue crystals (suitable for X-ray diffraction) were separated from the solution by decantation, washed with cold CH_3CN (5°C), and dried in air. The yield of compound **1** was 0.310 g (60%). Anal (%). Found: C, 53.5; H, 7.25; N, 3.15. Calculated for $\text{C}_{40}\text{H}_{64}\text{Co}_2\text{Li}_2\text{N}_2\text{O}_{12}$: C, 53.6; H, 7.2; N, 3.1. IR (cm^{-1}): 2960 m, 2924 w, 2902 w, 2969 w, 1603 s, 1577 s, 1563 s, 1482 s, 1457 m, 1447 s, 1419 s, 1403 s, 1378 s, 1358 s, 1218 s, 1153 w, 1070 m, 1043 m, 1030 w, 1015 v.w, 938 v.w, 893 m, 793 m, 754 m, 695 s, 678 v.w, 639 m, 613 s, 574 m, 551 w, 434 s.

Synthesis of $[\text{Li}_2\text{Zn}_2(\text{piv})_6(\text{py})_2]$. Heterometallic complex was obtained according to previously published technique [32].

Synthesis of $[\text{LiCo}(\text{btc})(\text{dma})_2]\cdot 2\text{DMA}$ (2**).** $[\text{Li}_2\text{Co}_2(\text{piv})_6(\text{py})_2]$ (25 mg, 0.03 mmol) was dissolved in 2 mL of DMA. Then 2 mL of H_3btc (25 mg, 0.12 mmol) DMA solution was added. The resultant mixture was sealed in a glass tube and heated up to 100°C (1°C/min) and kept at 100°C for 24 hours. Violet transparent crystals were filtered off and dried in air (yield 25 mg, 0.04 mmol, 69% based on the starting complex). Anal (%). Found: C, 48.3; H, 6.6; N, 9.0. Calculated for $\text{C}_{25}\text{H}_{39}\text{CoLiN}_4\text{O}_{10}$: C, 48.3; H, 6.3; N, 9.0. IR (KBr, cm^{-1}): 1629 s, 1598 s, 1509 m, 1442 s, 1411 s, 1349 s, 1260 m, 1185 m, 1091 w, 1038 m, 1016 m, 941 w, 807 w, 771 s, 749 w, 709 s, 594 m, 563 w, 482 m. Li:Co ratio found 1.08 : 1.

Synthesis of $[\text{LiZn}(\text{btc})(\text{dma})_2]\cdot \text{DMA}$ (3**).** $[\text{Li}_2\text{Zn}_2(\text{piv})_6(\text{py})_2]$ (2 mg, 0.002 mmol) was dissolved in 1 mL of DMA. Then 0.5 mL of H_3btc (2 mg, 0.008 mmol) DMA solution was added. The resultant mixture was sealed in a glass tube and heated up to 100°C (1°C/min) and kept at 100°C for 24 hours. A couple of colourless transparent crystals suitable for X-ray diffraction were obtained.

Synthesis of $[\text{LiZn}(\text{btc})(\text{dmf})_2(\text{H}_2\text{O})_2]\cdot 0.5\text{DMF}$ (4dmf**).** $[\text{Li}_2\text{Zn}_2(\text{piv})_6(\text{py})_2]$ (25 mg, 0.03 mmol) was dissolved in 2 mL of DMF. Then 2 mL of H_3btc (25 mg, 0.12 mmol) DMF solution was added. The resultant mixture was sealed in a glass tube and heated up to 100°C (1°C/min) and kept at 100°C for 24 hours. Colourless transparent crystals were filtered off and dried in air (yield 18 mg, 0.037 mmol, 62% based on the starting complex). Anal (%). Found: C, 39.5; H, 4.9; N, 7.3. Calculated for $\text{C}_{16.5}\text{H}_{24.5}\text{LiN}_{2.5}\text{O}_{10.5}\text{Zn}$: C, 39.8; H, 4.9; N, 7.0. IR (KBr, cm^{-1}): 1664 s, 1620 s, 1576 m, 1491 w, 1433 s, 1354 s, 1247 w, 1091 m, 1056 w, 931 w, 776 s, 722 s, 665 m, 554 w, 474 w. Li:Zn ratio found 1.09 : 1.

Synthesis of $[\text{LiZn}(\text{btc})(\text{dma})_2(\text{H}_2\text{O})_2]\cdot \text{H}_2\text{O}$ (4dma**).** The method of synthesis was similar to that for **4dmf** but in DMA solvent instead of DMF. Colourless transparent crystals were filtered off and dried in air (yield 21 mg, 0.042 mmol, 70% based on the starting complex). Anal (%). Found: C, 40.4; H, 5.2; N, 5.5. Calculated for $\text{C}_{17}\text{H}_{27}\text{LiN}_2\text{O}_{11}\text{Zn}$: C, 40.2; H, 5.3; N, 5.5. IR (KBr, cm^{-1}): 1642 s, 1575 m, 1500 w, 1438 s, 1358 s, 1251 w, 1096 m, 1047 w, 1011 w, 936 w, 771 s, 718 s, 665 m, 558 w, 465 w. Li:Zn ratio found 1.12 : 1.

Synthesis of $[\text{LiZn}(\text{btc})(\text{nmp})_2(\text{H}_2\text{O})_2]\cdot 2\text{NMP}$ (4nmp**).** The method of synthesis was similar to that for **4dmf** but in NMP solvent instead of DMF. Colourless transparent crystals were filtered

off and dried in air (yield 25 mg, 0.036 mmol, 60% based on the starting complex). Anal (%). Found: C, 48.3; H, 6.3; N, 7.8. Calculated for $C_{29}H_{43}LiN_4O_{12}Zn$: C, 48.9; H, 6.1; N, 7.8. IR (KBr, cm^{-1}): 1642 s, 1589 m, 1504 m, 1438 s, 1353 s, 1300 m, 1260 w, 1167 w, 1100 m, 980 w, 936 w, 776 s, 722 s, 656 m, 576 w, 469 m. Li:Zn ratio found 1.10 : 1.

Synthesis of $[LiCo(btbb)(dmf)_2] \cdot 2DMF$ (5dmf**).** $[Li_2Co_2(piv)_6(py)_2]$ (25 mg, 0.03 mmol) was dissolved in 2 mL of DMF and 4 mL of CH_3CN were added. Then 2 mL of H_3btb (53 mg, 0.12 mmol) DMF solution was added. The resultant mixture was sealed in a glass tube and heated up to 100°C (1°C/min) and kept at 100°C for 24 hours. Big violet transparent crystals were filtered off and dried in air (yield 30 mg, 0.038 mmol, 65% based on the starting complex). Anal (%). Found: C, 58.3; H, 5.9; N, 6.9. Calculated for $C_{39}H_{43}CoLiN_4O_{10}$: C, 59.0; H, 5.5; N, 7.0. IR (KBr, cm^{-1}): 1664 s, 1602 s, 1522 s, 1389 s, 1251 w, 1180 w, 1100 w, 1003 m, 851 m, 785 s, 709 w, 660 w, 469 m. Li:Co ratio found 1.14 : 1.

Synthesis of $[LiCo(btbb)(dma)_2] \cdot 2DMA$ (5dma**).** The method of synthesis was similar to that for **5dmf** but in DMA solvent instead of DMF. Big violet transparent crystals were filtered off and dried in air (yield 36 mg, 0.042 mmol, 72% based on the starting complex). Anal (%). Found: C, 60.3; H, 5.9; N, 6.6. Calculated for $C_{43}H_{51}CoLiN_4O_{10}$: C, 60.8; H, 6.0; N, 6.6. IR (KBr, cm^{-1}): 1620 s, 1593 s, 1553 w, 1504 w, 1380 s, 1261 w, 1176 w, 1140 w, 1096 w, 1012 m, 865 m, 803 w, 785 s, 701 w, 665 w, 580w, 482 m. Li:Co ratio found 1.12 : 1.

Synthesis of $[LiZn(btbb)(dmf)_2] \cdot 2DMF \cdot H_2O$ (6dmf**).** The method of synthesis was similar to that for **5dmf** but with $[Li_2Zn_2(piv)_6(py)_2]$ as starting complex. Colourless transparent crystals were filtered off and dried in air (yield 36 mg, 0.044 mmol, 75% based on the starting complex). Anal (%). Found: C, 56.8; H, 5.3; N, 6.4. Calculated for $C_{39}H_{45}LiN_4O_{11}Zn$: C, 57.2; H, 5.5; N, 6.8. IR (KBr, cm^{-1}): 1651 s, 1593 s, 1549 m, 1496 w, 1442 m, 1385 s, 1251 m, 1180 m, 1135 w, 1096 m, 1056 w, 1011 w, 860 m, 807 w, 780 s, 713 w, 665 m, 563 w, 496 m. Li:Zn ratio found 1.11 : 1.

Synthesis of $[LiZn(btbb)(dma)_2] \cdot 2DMA$ (6dma**).** The method of synthesis was similar to that for **5dma** but in DMA solvent instead of DMF and with $[Li_2Zn_2(piv)_6(py)_2]$ as starting complex. Colourless transparent crystals were filtered off and dried in air (yield 34 mg, 0.041 mmol, 68% based on the starting complex). Anal (%). Found: C, 60.0; H, 5.9; N, 6.6. Calculated for $C_{43}H_{51}LiN_4O_{10}Zn$: C, 60.3; H, 6.0; N, 6.5. IR (KBr, cm^{-1}): 1660 s, 1584 s, 1558 m, 1500 w, 1438 m, 1380 s, 1251 m, 1176 w, 1140 w, 1100 m, 1060 w, 1016 m, 856 m, 807 w, 785 s, 709 w, 665 m, 482 m. Li:Zn ratio found 1.08 : 1.

Guest exchange and synthesis of $[LiZn(btbb)(H_2O)_4] \cdot C_6H_5NO_2$ (4dmf-NB**), $[LiZn(btbb)(H_2O)_4] \cdot C_8H_9NO_2$ (**4dmf-NX**) and $[LiZn(btbb)(H_2O)_4] \cdot C_8H_9NO_2$ (**4dmf-DMNB**).** Fresh crystals of **4dmf** were soaked in the correspondent guest solution (NB - nitrobenzene, NX - 4-nitro-m-xylene, DMNB - 2,6-dimethylnitrobenzene) (1 mL) for 48 hours, then filtered and dried in air. Anal (%). Found for **4dmf-NB**: C, 37.1; H, 3.3; N, 2.5. Calculated for $C_{15}H_{16}LiNO_{12}Zn$: C, 37.9; H, 3.4; N, 2.9.

Found for **4dmf-NX**: C, 38.3; H, 3.5; N, 3.2. Calculated for $C_{17}H_{20}LiNO_{12}Zn$: C, 40.6; H, 4.0; N, 2.8. Found for **4dmf-DMNB**: C, 40.2; H, 3.5; N, 3.7. Calculated for $C_{17}H_{20}LiNO_{12}Zn$: C, 40.6; H, 4.0; N, 2.8.

Guest exchange and synthesis of $[LiZn(btbb)(dmf)_2] \cdot 2C_6H_5NO_2$ (6dmf-NB**), $[LiZn(btbb)(dmf)_2] \cdot 2C_8H_9NO_2 \cdot H_2O$ (**6dmf-NX**) and $[LiZn(btbb)(dmf)_2] \cdot 2C_8H_9NO_2$ (**6dmf-DMNB**).** Fresh crystals of **6dmf** were soaked in the correspondent guest solution (NB - nitrobenzene, NX - 4-nitro-m-xylene, DMNB - 2,6-dimethylnitrobenzene) (1 mL) for 48 hours, then filtered and dried in air. Anal (%). Found for **6dmf-NB**: C, 59.6; H, 4.5; N, 6.5. Calculated for $C_{45}H_{39}LiN_4O_{12}Zn$: C, 60.0; H, 4.4; N, 6.2. Found for **6dmf-NX**: C, 61.7; H, 5.1; N, 5.8. Calculated for $C_{49}H_{47}LiN_4O_{12}Zn$: C, 61.5; H, 5.0; N, 5.8. Found for **6dmf-DMNB**: C, 59.5; H, 5.2; N, 5.7. Calculated for $C_{49}H_{49}LiN_4O_{13}Zn$: C, 60.4; H, 5.1; N, 5.7.

X-ray crystallography.

The X-ray diffraction study of complex **1** was carried out on an Enraf Nonius CAD-4 diffractometer (graphite-monochromatized Mo-K α radiation, $\lambda = 0.71073 \text{ \AA}$) at 293 K according to a standard technique [56]. Semi-empirical absorption corrections were applied [57]. The structure was solved by direct methods and by Fourier techniques and was refined by the full-matrix least squares method against F^2 with anisotropic thermal parameters for all non-hydrogen atoms. The hydrogen atoms of the carbon-containing ligands in the compounds were positioned geometrically and refined using the riding model. All calculations were carried out with the use of the SHELX97 and SHELXL-2013 software packages [58]. Diffraction data for single crystals of **2**, **3**, **5dma** and **6dmf** were obtained at 130 K on an automated Agilent Xcalibur diffractometer equipped with a two-dimensional AtlasS2 detector (graphite monochromator, $\lambda(\text{MoK}\alpha) = 0.71073 \text{ \AA}$). Integration, absorption correction, and determination of unit cell parameters were performed using the CrysAlisPro program package [59]. The structures of all compounds were solved by a direct method and refined by the full-matrix least squares technique in the anisotropic approximation (except hydrogen atoms) using the SHELX-2014 software [57]. Positions of hydrogen atoms of organic ligands were calculated geometrically and refined in the riding model. The crystallographic data and details of the structure refinements are summarized in Table 1. CCDC 1883246-1883250 contain the supplementary crystallographic data for this paper. These data can be obtained free of charge from The Cambridge Crystallographic Data Center at http://www.ccdc.cam.ac.uk/data_request/cif.

Surface Area and Porous Structure.

An analysis of the porous structure was performed by a nitrogen adsorption technique using Quantochrome's Autosorb iQ at 77 K. Initially the compounds **5** and **6** were first activated in dynamic vacuum (10^{-8} bar) at 70°C during 6 h after guest exchange with acetone for 48 h. The nitrogen sorption on activated sample **2** was studied on a dynamic SORBTOMETER-M instrument by 5-point method in relative

pressures range 0.05–0.25. The sample was activated under N₂ stream at 100°C for 30 min.

N₂ adsorption–desorption isotherms were measured within the range of relative pressures of 10⁻⁶ to 0.995. The specific surface area was calculated from the data obtained on the basis of the conventional BET, Langmuir and DFT models. Pore size distributions were calculated using DFT.

N₂, CO₂ and CH₄ Sorption Experiments at 273 and 298 K. Gases adsorption isotherm measurements were carried out volumetrically on Quantochrome's Autosorb iQ equipped with thermostat TERMEX M01 to adjust temperature with 0.1 K accuracy. Adsorption–desorption isotherms were measured within the range of pressures of 1 to 800 torr. The database of the National Institute of Standards and Technology Thermophysical Properties of Fluid Systems, Database of National Institute of Standards and Technology, NIST. <http://webbook.nist.gov/chemistry/fluid/> was used as a source of p–V–T relations at experimental pressures and temperatures.

Conclusions

Substitution of pivalate ligands in [Li₂M₂(piv)₆(py)₂] (where M=Zn²⁺, Co²⁺) with tricarboxylate bridges (btc³⁻, btb³⁻) affords two general structural types of coordination polymers: 3D frameworks with chiral SrSi₂ topology when the carboxylate group plane towards the plane of benzene ring in btc³⁻ lies in the same plane or 2D honeycomb structures in case of btb³⁻ linker when the angle between the central benzene ring and the plane of COO⁻ group is close to 90°. In both cases binuclear {LiM(RCOO)₃} 3-connected nodes preserve their structure during the ligand substitution demonstrating a rare example of rational step-by-step MOF structure design from preorganized heterometallic complexes. The obtained layered Zn-containing coordination polymers are capable of a guest-dependent luminescent behavior with drastic changes in luminescence intensity and quantum yields on nitroaromatics inclusion. Moreover, a rare example of permanent porosity and fine selectivity in both cases of CO₂/N₂ and CO₂/CH₄ separation for layered compounds was found.

Conflicts of interest

There are no conflicts to declare.

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Table 1. Crystal data and structure refinement for 1, 2, 3, 5 and 6

Identification code	1	2	3	5dma	6dmf
Empirical formula	C ₄₀ H ₆₄ Li ₂ N ₂ O ₁₂ CO ₂	C ₂₅ H ₃₉ CoLiN ₄ O ₁₀	C ₂₁ H ₃₀ LiN ₃ O ₉ Zn	C ₃₉ H ₄₂ CoLiN ₃ O ₉	C ₃₆ H ₃₆ LiN ₃ O ₉ Zn
<i>M</i> , g/mol	896.67	621.47	540.79	762.62	726.99
<i>T</i> , K	293	130	130	130	130
Crystal system	Triclinic	Tetragonal	Tetragonal	Monoclinic	Monoclinic
Space group	<i>P</i> -1	<i>P</i> 4 ₃ 2 ₁ 2	<i>P</i> 4 ₁ 2 ₁ 2	<i>I</i> 2/ <i>a</i>	<i>I</i> 2/ <i>a</i>
<i>a</i> , Å	9.379(3)	13.1202(2)	13.2635(3)	11.2469(4)	11.8491(5)
<i>b</i> , Å	12.106(5)	13.1202(2)	13.2635(3)	29.1732(9)	29.0049(10)
<i>c</i> , Å	12.656(5)	17.6987(4)	17.6166(4)	12.3612(4)	12.2371(4)
α , deg.	117.147(10)	90	90	90	90
β , deg.	96.610(10)	90	90	92.248(3)	93.249(3)
γ , deg.	89.912(10)	90	90	90	90
<i>V</i> , Å ³	1268.0(8)	3046.65(12)	3099.12(17)	4052.7(2)	4198.9(3)
<i>Z</i>	1	4	4	4	4
<i>D</i> (calcd), g/cm ³	1.174	1.355	1.159	1.250	1.150
μ , mm ⁻¹	0.706	0.621	0.836	0.478	0.634
<i>F</i> (000)	474	1308	1128	1596	1512
Crystal size, mm	0.43 × 0.23 × 0.12	0.20 × 0.20 × 0.18	0.33 × 0.28 × 0.22	0.38 × 0.17 × 0.11	0.23 × 0.21 × 0.15
θ range for data collection, deg.	1.89–24.41	3.31–26.37	3.44–29.45	3.58–29.50	3.44–25.35
Index ranges	0 ≤ <i>h</i> ≤ 9, -14 ≤ <i>k</i> ≤ 14, -14 ≤ <i>l</i> ≤ 14	-14 ≤ <i>h</i> ≤ 12, -16 ≤ <i>k</i> ≤ 11, -22 ≤ <i>l</i> ≤ 13	-17 ≤ <i>h</i> ≤ 17, -11 ≤ <i>k</i> ≤ 16, -24 ≤ <i>l</i> ≤ 22	-14 ≤ <i>h</i> ≤ 15, -34 ≤ <i>k</i> ≤ 36, -14 ≤ <i>l</i> ≤ 15	-14 ≤ <i>h</i> ≤ 14, -34 ≤ <i>k</i> ≤ 28, -14 ≤ <i>l</i> ≤ 11
Reflections collected / independent	4744 / 4015	8033 / 3093	10902 / 3696	10599 / 4784	9055 / 3838
<i>R</i> _{int}	0.0539	0.0198	0.0272	0.0186	0.0220
Reflections with <i>I</i> > 2 σ (<i>I</i>)	3027	2907	3266	3780	3040
Goodness-of-fit on <i>F</i> ²	1.110	1.092	1.133	1.053	1.061
Final <i>R</i> indices [<i>I</i> > 2 σ (<i>I</i>)]	<i>R</i> ₁ = 0.0718, <i>wR</i> ₂ = 0.1910	<i>R</i> ₁ = 0.0359, <i>wR</i> ₂ = 0.0824	<i>R</i> ₁ = 0.0637, <i>wR</i> ₂ = 0.1849	<i>R</i> ₁ = 0.0782, <i>wR</i> ₂ = 0.2430	<i>R</i> ₁ = 0.0657, <i>wR</i> ₂ = 0.2063
<i>R</i> indices (all data)	<i>R</i> ₁ = 0.0944, <i>wR</i> ₂ = 0.2109	<i>R</i> ₁ = 0.0398, <i>wR</i> ₂ = 0.0842	<i>R</i> ₁ = 0.0712, <i>wR</i> ₂ = 0.1951	<i>R</i> ₁ = 0.0934, <i>wR</i> ₂ = 0.2578	<i>R</i> ₁ = 0.0795, <i>wR</i> ₂ = 0.2180
Largest diff. peak / hole, e/Å ³	0.500 / -0.798	0.375 / -0.156	0.851 / -0.405	0.875 / -0.804	0.645 / -0.509