Tetranuclear [Fe II 2 Fe III 2] 2+ molecular switches [Fe II (bik) 2 (N–) 2]spin-crossover complexes containing [Fe III (Tp)(CN) 3] – metalloligands as N-donor

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Tetranuclear $[Fe^{II}_2Fe^{III}_2]^{2+}$ molecular switches: $[Fe^{II}(bik)_2(N\text{-})_2]$ spin-crossover complexes containing $[Fe^{III}(Tp)(CN)_3]^-$ metalloligands as *N*-donor [#]

[#]This work is dedicated to Professor Michel Verdaguer

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Abstract: Three novel mixed valence cyanide-bridged $\{Fe^{III}_2Fe^{II}_2\}$ square complexes were obtained through the self-assembling of $[Fe^{III}(Tp)(CN)_3]^-$ or $[Fe^{III}(Tp^*)(CN)_3]^-$ cyanido building blocks with the in-situ formed $[Fe^{II}(bik)_2(S)_2]$ complex (Tp = hydrotris(pyrazol-1-yl)borate, Tp* = hydrotris(3,5-dimethyl-pyrazol-1-yl)borate, bik = bis(1-methylimidazol-2-yl)ketone, S = solvent). The structure of these three complexes (**2**, **3** and **4**) are reminiscent of that of our previously published square complex $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}.[Fe^{III}(Tp)(CN)_3]_2\cdot 18H_2O\cdot 4CH_3OH$ (**1**). They consist of cyanide-bridged square dicationic complexes, ClO_4^- (**2** and **3**) or BF_4^- (**4**) counter ions and solvate molecules. The FT-IR cyanide stretching vibrations observed at $\nu_{\square\square} \approx 2145\text{-}60$ cm^{-1} are typical of $\{Fe^{III}\text{-}CN\text{-}Fe^{II}\}$ moieties. The investigation of the magnetic properties of **2** reveals the occurrence of spin-crossover centred at $T_{1/2}$ = 227 K. The $\chi_M T$ variation, *ca.* 7 cm^3 mol^{-1} K, reflects the complete spin-state change occurring on both $\{Fe^{II}(bik)(\text{-}NC)_2\}$ moieties (-NC represent the cyanido building blocks). The Slichter-Drickammer model leads to a weak cooperativity factor, Γ = 1.6 kJ.mol^{-1} (with Γ □ □□□□□□), which reflects the gradual spin-state change. This is in agreement with the molecular structure of **2**, which does not present significant intermolecular interactions. The calculated enthalpy and entropy variation associated to the spin-state equilibrium are ΔH = 24 kJ.mol^{-1} and ΔS = 105 J. K^{-1}mol^{-1}. In contrast **3** and **4** show only partial spin crossover in the accessible temperature range (2 - 400 K) as the $T_{1/2}$ are shifted toward higher temperatures (*ca.* $T_{1/2}$ > 400 K). While no photomagnetic effect is observed for **3**, compound **4** shows a moderate increase of the magnetization upon irradiation at low temperature. This phenomenon is ascribed to the Light-Induced Excited Spin-State Trapping (LIESST) effect. Interestingly, the complex **2** also

shows a remarkable LIESST effect, which is observed with different laser lights covering the visible and NIR range. The resulting $\chi_M T$ value obtained in the photoinduced state suggests the occurrence of a ferromagnetic interaction inside the {Fe^{III}-CN-Fe^{II}} units.

Mots clés: photomagnétisme, complexe à transition de spin, Fe^{II}, ligand cyanure

Résumé: Trois nouveaux complexes carrés {$Fe^{III}_2Fe^{II}_2$} à pont cyanure et à valence mixte ont été obtenus par auto-assemblage des précurseurs cyanurés $[Fe^{III}(Tp)(CN)_3]^-$ ou $[Fe^{III}(Tp^*)(CN)_3]^-$ et du complexe cationique $[Fe^{II}(bik)_2(S)_2]$ formé in-situ (Tp = hydrotris(pyrazol-1-yl)borate, Tp* = hydrotris(3,5-dimethyl-pyrazol-1-yl)borate, bik = bis(1-methylimidazol-2-yl)cétone, S = solvant). Les structures de ces trois complexes (**2**, **3** et **4**) sont similaires à celle du complexe $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}.[Fe^{III}(Tp)(CN)_3]_2 \cdot 18H_2O \cdot 4CH_3OH$ (**1**) préalablement publié. Elles sont constituées de complexes carrés dicationiques à pont cyanure, de contre-ions ClO_4^- (**2** and **3**) ou BF_4^- (**4**) et de molécules de solvant. Les vibrations d'élongation des cyanures, observée en spectroscopie IR à $\nu_{\square\square} \approx 2145\text{-}60$ cm^{-1}, sont caractéristiques d'unités {Fe^{III}-CN-Fe^{II}}. L'étude des propriétés magnétiques de **2** révèle un équilibre de spin centré à $T_{1/2}$ = 227 K. La variation du produit $\chi_M T$, *ca.* 7 cm^3 mol^{-1} K, traduit une conversion de spin complète sur chacune des unités {$Fe^{II}(bik)(\text{-NC})_2$} du carré (-NC représente le complexe précurseur cyanuré). L'analyse des données par le modèle de Slichter-Drickammer conduit à un faible facteur de coopérativité, Γ = 1.6 $kJ.mol^{-1}K$ (with Γ □ □□□$_{□□□}$), en accord avec un changement d'état de spin graduel. Ces données sont en accord avec la structure de **2** qui ne montre pas d'interactions intermoléculaires notables. Les valeurs des variations d'enthalpie et d'entropie associées à la conversion de spin sont ΔH = 24 $kJ.mol^{-1}$ et ΔS = 105 $J.K^{-1}.mol^{-1}$. Au contraire de **2**, les composés **3** and **4** présentent seulement une conversion de spin partielle dans le domaine de température exploré (2 - 400 K) avec des valeurs $T_{1/2}$ déplacées vers les hautes températures (*ca.* $T_{1/2}$ > 400 K). Tandis qu'on n'observe pas d'effet photomagnétique pour **3** et seulement un faible effet dans **4**, le composé **2** présente une forte augmentation de son aimantation sous irradiation à basse température. Cet effet est dû au piégeage photo-induit d'un état excité de spin (effet « LIESST », Light-Induced Excited Spin-State Trapping) Il est observé avec différentes sources laser couvrant le spectre visible et le proche IR. Les valeurs de $\chi_M T$ obtenues dans l'état photo-induit suggèrent la présence d'une interaction ferromagnétique au sein de la paire {Fe^{III}-CN-Fe^{II}}.

1 Introduction and background

Switchable molecular systems featuring electronic, magnetic, or optical bistability are attracting a strong research interest because of their potential use as molecular memories, switches, actuators, or sensors. **[1–3]** Although their integration into actual devices is challenging, various encouraging results show promise in emerging fields such as molecular electronics.**[4–7]** The cyanide coordination chemistry has proven successful in providing access to a variety of responsive systems, whose optical and magnetic properties can be reversibly switched.**[8–10]** One of the emblematic example are the FeCo Prussian blue analogues, which show an increase of magnetization under light irradiation.[11] In these materials, the so-called photomagnetic effect was ascribed to a metal-metal electron transfer coupled to a spin-state change induced by light irradiation.[12,13] The system can then relax from the photo-induced metastable state to the ground state upon heating. Since few years various groups have been investigating polymetallic complexes showing photomagnetism.**[14,15]** In order to reduce the dimensionality of the FeCo PBAs, a common strategy is based on the use of $[Fe^{III}(L)(CN)_x]^-$ metalloligand toward partially blocked Co precursors.**[16]** More specifically, many of these systems were obtained with *fac*-$[Fe^{III}(Tp)(CN)_3]^-$ building blocks, where the Tp is a derivative of the hydrotris(pyrazol-1-yl)borate scorpionate ligand. These low dimensional systems can be used as model to better understand the photo-magnetic properties**[17,18]** or they can also be integrated using soft chemistry approaches into molecule based devices[19]. In the last years we have been interested in the study of $\{Fe_2Co_2\}$ squares obtained by self-assembling the $[Fe^{III}(Tp)(CN)_3]^-$ with partially blocked $[Co(bik)_2(S)_2]$ complexes (S = solvent, bik = bis(1-methylimidazol-2-yl)ketone).**[16,20,21]** Additionally, we also explored the use of $[Fe^{III}(Tp)(CN)_3]^-$ complexes as ligand toward other metal ions. The use of partially blocked Fe(II) complexes is of particular interest as the resulting material exhibits $Fe^{II}N_6$ coordination sites, which provide a suitable surrounding for observing spin-crossover phenomenon. As for the FeCo switchable materials, the spin-state change from the Fe^{II} low spin (LS) to the Fe^{II} high spin (HS) can be thermally induced and photo-induced at low temperature. In that case the photomagnetic properties arise from the so-called "Light-Induced Excited Spin State Trapping" (LIESST) effect. We first reported in 2012 the tetrametallic mixed-valence complex, $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}.[Fe^{III}(Tp)(CN)_3]_2\cdot 18H_2O\cdot 4CH_3OH$ (**1**), which exhibits a square core structure $\{Fe^{III}_2Fe^{II}_2\}$, while two $[Fe^{III}(Tp)CN)_3]^-$ act as counter ions.**[22]** The spin-state changes occurring in this complex are very dependent on the solvatation state of the material. The solvated phase shows a transition temperature located slightly above room temperature, $T_{1/2}$ = 330 K, while in the desolvated phase the transition is shifted toward low temperatures, with $T_{1/2}$ = 240 K. Moreover, the fresh solvated compound is not photomagnetic whereas the desolvated one shows an

efficient LIESST effect at 750 nm, with T_{relax} approx. 45 K (the relaxation temperature is measured by heating the sample at 0.3 K/min).

In the last years, we have decided to extend this work to study the influence of the *N*- donor on the switchable properties of the $[Fe^{II}(bik)_2(N\text{-})_2]$ spin-crossover subunit, where *N*- stands for N donor organic/inorganic ligand or metaloligands**[23–25]**. We observed that the spin transition temperature changed with the nature of the *N*- donor (metallo)ligand. Different *N*- donors can lead to different ligand field on the Fe(II) SCO units and thus a change in the LS / HS energy gap. However it is difficult to draw correlation in solid-state studies as other factors, such as the intermolecular interactions**[26,27]** or the inter-ligand interactions also affect the $T_{1/2}$.**[28,29]** For those cases where the compared spin-crossover complexes exhibit a similar structure such as in the case of the neutral $[Fe^{II}(bik)_2(NCX)_2]$ spin crossover complexes (X= S, Se) the rationalization of the properties is easier. Our experimental and theoretical study showed that the $NCSe^-$ ligands induce a stronger ligand field than the NCS^- one, and thus a higher spin-crossover transition temperature, $T_{1/2}$.[24] Comparison between the transition temperatures of similar complexes are also easier to rationalize in solution studies, as the influence of the intermolecular interaction vanishes. For example, we have shown that it was possible to rationalize the transition temperature in the family $[Fe(^{R}bik)_3]^{2+}$ and $[Fe(bim)_3]^{2+}$ complexes (R are methyl , ethyl or vinyl groups grafted on the imidazolyl donors of the bik ligand; bim is similar to bik but the ketone is replace by a methane group). Our experimental multitechnique approach led to results that are coherent with the theoretical prediction. Additionaly, we showed in this study that among the different techniques used to follow the spin-crossover equilibrium (UV-vis, Evans NMR, magnetic measurments in solution), the monitoring of an adequatly chosen ^1H NMR signal could be a very straigthforward and accurate approach.

More interestingly, we have systematically studied all these $[Fe(bik)_2(N\text{-})_2]$ systems by comparing the photomagnetic effect at 20 K by using different wavelengths in the visible and close infra-red regions (from 405 nm to 1313 nm). When replacing the $[Fe^{III}(Tp)(CN)_3]^-$ metalloligand with $[Mo^{V}(CN)_8]^{3-}$, we obtained another square complex of formula $\{[Mo(CN)_8]_2[Fe(bik)_2]_2\}(HMeIm)_2\cdot 5H_2O\cdot CH_3CN$, which exhibits similar $[Fe^{II}(bik)_2(N\text{-})_2]$ chomophores.[23] Here the thermally-induced spin transition is very gradual. It occurs near room temperature and it is accompanied by a solvent loss, that makes the spin-state change irreversible. The dehydrated compound is still photomagnetic with a close relaxation temperature (T_{relax} *ca.* 48 K). However the most efficient irradiation wavelength is located at 405 nm, which strikingly contrast with the situation observed in the previously reported $\{Fe^{III}_2Fe^{II}_2\}$ square, whose photomagnetic effect is higher near 750 nm. These results are to be compared with the model spin-crossover compound $[Fe^{II}(bik)_3](BF_4)_2$, which shows a spin-crossover centered at $T_{1/2}$ = 316 K and a photomagnetic effect, which is more efficient (but uncomplete) at 635 nm. Finally, in the

$[Fe^{II}(bik)_2(NCX)_2]$ complexes, the photomagnetic effect is still observed although it is incomplete. The most efficient wavelength to promote a LIESST effect is 900 nm (while the induced magnetization is lower at 808 and 635 nm). However these system relax at 20 K (the temperature set of all our photomagnetic measurments) when switching off the light, indicating thus the existence of an efficient relaxation pathway. It is worth noticing that no significant difference are observed in the visible absorption spectra of the all these spin crossover compounds. They all exhibit a broad absorption centered at *ca*. 630 nm and ascribed to a Metal-Ligand Charge Transfer (MLCT). No MMCT are observed for the polymetallic square complexes. Continuing these experimental investigations, we looked here at the properties of three novel $\{Fe_2Fe_2\}^{2+}$ square complexes $\{[Fe^{III}(Tp^*)(CN)_3]_2[Fe^{II}(bik)_2]\}_2(ClO_4)_2\cdot 2H_2O$ (**2**), $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}(ClO_4)_2\cdot 2CH_3CN$ (**3**), $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]\}_2(BF_4)_2\cdot 2CH_3OH$ (**4**) (Tp = hydrotris(pyrazol-1-yl)borate, Tp* = hydrotris(3,5-dimethyl-pyrazol-1-yl)borate, bik = bis(1-methylimidazol-2-yl)ketone). On the one hand two different $[Fe^{III}(L)(CN)_3]^-$ metaloligands were used, and on the other hand the counter ions are different so as the intermolecular interactions. Our purpose was to check in which extent the photomagnetic properties can be influenced by the metalloligand, the solid state structure and intermolecular interactions.

2 Results and discussion

2.1. Syntheses and general characterization

The $PPh_4[Fe^{III}(Tp)(CN)_3]\cdot H_2O$ complex was prepared in three steps by following our previously reported procedure [30]. First we isolated a $[Fe^{II}(Tp)_2]$ complex by direct reaction of the scorpionate ligand, KTp, and a Fe(II) salt. In a second step, the reaction of three equivalents of KCN with the recrystallized dried $[Fe^{II}(Tp)_2]$ allows the formation of the tricyanido complex $[Fe^{II}(Tp)(CN)_3]^{2-}$. This one is then carefully oxidized by hydrogen peroxide and recrystallized from acetonitrile solution after a salt metathesis reaction. The ammonium salt of the $[Fe^{III}(Tp^*)(CN)_3]^-$ building block can be obtained similarly in three steps, as reported by Holmes *et coll.* [31], however the result is more uncertain. The instability of the first intermediate product, $[Fe^{II}(Tp^*)_2]$, (that need to be immediately used), the potential hazards related to the partial solvent removal from a hydrogen peroxide contaminated acetonitrile solution, and its layering by diethyl ether are major drawbacks. We thus developed a new one-step synthesis whose yield is lower (*ca*. 32%) but which allows a straightforward access to single crystals of $PPh_4[Fe^{III}(Tp^*)(CN)_3]\cdot CH_3CN$ (crystallographic data and a view of the crystal structure, Fig. S1, are given in supplementary material). In this synthesis, the degased solution containing equivalent amount of Fe(II) salt and KTp* is to be added to the cyanide solution shortly after its preparation as $[Fe^{II}(Tp^*)_2]$ tends to precipitate over time. Then the oxidation is carried out by opening the flask to the air. As shown in supplementary material (Fig. S2), the redox potential of the obtained $PPh_4[Fe^{III}(Tp^*)(CN)_3]\cdot CH_3CN$ complex (-1.00 V to be compared with -0.82 V for $[Fe^{III}(Tp)(CN)_3]^-$; *versus* $[Fe^{III/II}(Cp)_2]$, 10^{-3} M in CH_3CN) allows a clean oxidation by air's oxygen in acetonitrile, avoiding so the critical oxidation step by hydrogen peroxide.

The square complexes were obtained by self-assembling the $[Fe^{III}(Tp)(CN)_3]^-$ or $[Fe^{III}(Tp^*)(CN)_3]^-$ building blocks with the *in-situ* formed $[Fe^{II}(bik)_2(S)_2]^{2+}$ complexes in alcoholic or acetonitrile solutions. While the use of $FeCl_2$ salt leads to the crystallization of the previously reported $\{Fe^{III}_2Fe^{II}_2\}^{2+}$ square complex with the $[Fe^{III}(Tp)(CN)_3]^-$ complex acting as counter ion,[22] the presence of fluoroborate or perchlorate anions in the solution (coming from the use of $Fe^{II}(BF_4)_2$ or $Fe^{II}(ClO_4)_2$ salts) leads to the fast crystallization of the dicationic $\{Fe^{III}_2Fe^{II}_2\}X_2$ square complexes (X = ClO_4, BF_4) in relatively good yield (**2**: 40%; **3**: 60%, **4**: 60%).

The compounds were all characterized by FT-IR spectroscopy. This technique is very useful to characterize the electronic states of cyanido-bridged complexes as the stretching cyanide vibration are sensitive to the oxidation state of the coordinated metal ions and to the bridging mode of the cyanide (bridging and non bridging). Typically, non bridging Fe^{III}-CN moieties are observed in the range 2130-2110 cm^{-1}, as observed for the precursors: 2123 cm^{-1} for $PPh_4[Fe^{III}(Tp^*)(CN)_3]\cdot CH_3CN$ 2121 cm^{-1} for $PPh_4[Fe^{III}(Tp)(CN)_3]\cdot H_2O$. In contrast stretching vibrations due to bridging cyanide

moieties are expected at higher wavenumber (*ca.* > 2145cm^{-1}). Here the observed cyanide stretching vibrations of **2-4** are reported in the table 1 together with other characteristic FT-IR peaks. They are very similar for all compounds and reveal the presence of one non-bridging Fe^{III}-CN unit and two different Fe^{III}-CN-Fe^{II} bridges in agreement with the crystal structure of the compounds (see below). The elemental analyses carried out on microcrystalline samples agree with the above formula for **4**, while a higher amount of water molecules (5 H_2O) is found in **2** in comparison to the crystal structure (2 H_2O), and a lower amount of CH_3CN is found for **3**, in comparison to the crystal structure. This can be associated to the presence of disordered water molecules in the crystal lattice of **2** and a partial loss of solvent molecule in **3**.

Table 1.
*Analytical and selected infra-red data of **2**, **3** and **4***

compounds	*Selected FT-IR vibration (cm^{-1})*					*Analyses (%)*[a]		
	ν_{CN} *bridging*	ν_{CN} *terminal*	*counter ions*	ν_{BH}	$\nu_{C=O}$	*C*	*H*	*N*
2	2160	2130	1093, 1064, 1054	2538	1634	42.88	4.43	23.49
	2147		(ClO_4^-)			(42.73)	(4.68)	(23.53)
3	2160	2113	1076, 1064, 1045	2503	1632	41.45	3.60	27.39
	2148		(ClO_4^-)			(41.23)	(3.52)	(27.14)
4	2161	2113	1109, 1094, 1045, 995	2510	1633	39.47	3.80	26.78
	2148		(BF_4^-)			(41.27)	(3.80)	(26.40)

[a] *calculated data are given in parentheses for the formula {[Fe^{III}(Tp*)(CN)$_3$]$_2$[Fe^{II}(bik)$_2$]}$_2$(ClO_4)$_2$ ·**5**H_2O (**2**), {[Fe^{III}(Tp)(CN)$_3$]$_2$[Fe^{II}(bik)$_2$]$_2$}(ClO_4)$_2$ **1**CH_3CN (**3**),{[Fe^{III}(Tp)(CN)$_3$]$_2$[Fe^{II}(bik)$_2$]}$_2$(BF_4)$_2$ 2CH_3OH (**4**)*

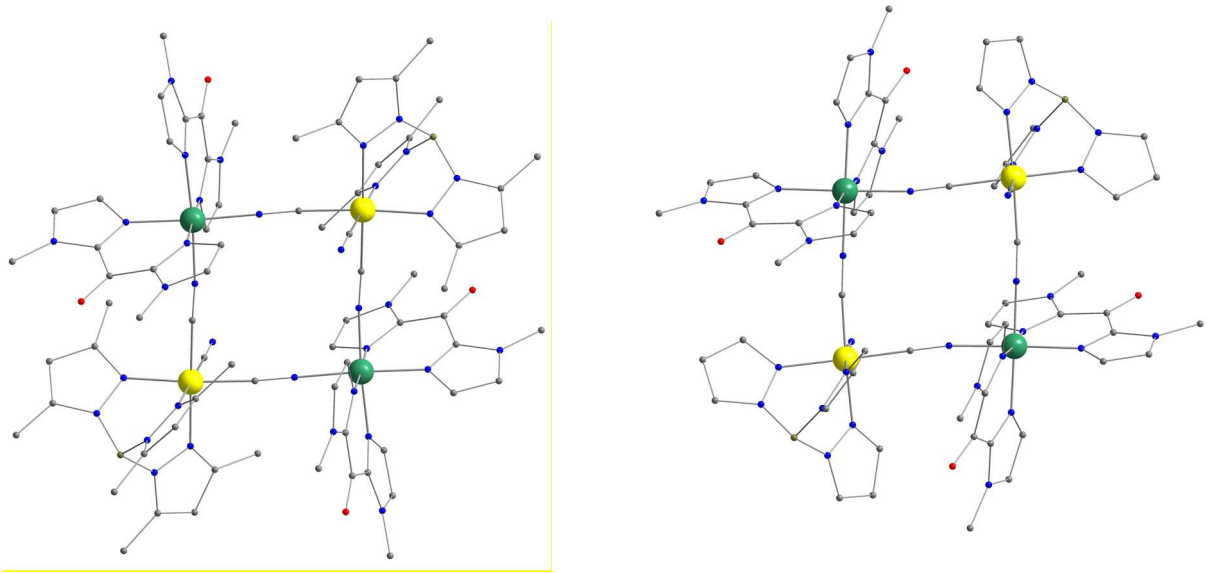

Fig. 1. Perspective views of the {$Fe^{III}_2Fe^{II}_2$} square complexes **2** (left) and **3** (right). C: grey, N: blue, O : red, B: pink-grey, Fe^{III}: yellow, Fe^{II}: green. Hydrogen atoms are omitted for clarity.

2.2. Description of the structures

Table 2
Crystal data and structure refinement parameters for **2-4**

Compounds	**2**	**3**	**4**
Formula	$C_{72}H_{88}B_2Fe_4N_{34}O_{14}Cl_2$	$C_{64}H_{66}B_2Fe_4N_{36}O_{12}Cl_2$	$C_{62}H_{68}B_4Fe_4N_{34}O_6F_8$
Fw (g/mol)	1969.62	1847.34	1804.04
Crystal system	triclinic	monoclinic	monoclinic
a (Å)	13.430(3)	15.2373(4)	14.9451(5)
b (Å)	13.470(3)	18.5182(4)	18.3777(6)
c (Å)	13.907(3)	15.1215(4)	15.5415(5)
α (°)	102.34(3)	90	90
β (°)	108.67(3)	110.4755(3)	101.927(2)
γ (°)	107.01(3)	90	90
V (Å^3)	2143.9(12)	3997.24(17)	4176.4(2)
Z	1	2	2
Space group	*P*-1	*P*2$_1$/*c*	*P*2$_1$/*c*
μ (cm^{-1})	0.808	0.861	0.771
ρ (g cm^3)	1.525	1.535	1.485
Merging R	0.241	0.0925	0.038
R [a]	0.0877	0.0507	0.0601
Rw [b]	0.1582 (0.1819)[c]	0.1196 (0.1324) [c]	0.1789(0.2049) [c]
Gof-fit	0.997	0.963	0.988
$\Delta\rho_{min}$ (e Å^{-3})	-2.82	-0.574	-1.53
$\Delta\rho_{max}$ (e Å^{-3})	1.42	0.657	1.88
T (K)	200	140	200

[a] $R = \Sigma||F_o| - |F_c|| / \Sigma|F_o|$
[b] $[\Sigma w(||F_o| - |F_c||)^2 / \Sigma w F_o^2]^{1/2}$
[c] *For $I > 2\sigma$(for all data in parentheses)*

Red prismatic crystals of **1** and orange-red square-like crystals of **3-4** suitable for single crystal X-ray diffraction experiment were obtained by slow evaporation of the mother solutions. The crystal data were all collected at low temperature (200 or 140 K). Compound **2** and **3-4** crystallize in the *P-1* and *P*2$_1$/*c* space groups, respectively. The structures of **2-4** consist of centrosymmetric dicationic cyanide-bridged tetrametallic units $\{Fe^{III}_2Fe^{II}_2\}$, perchlorate (**2-3**) or tetrafluoroborate (**4**) counter ions and lattice solvent molecules. The mixed valence tetrametallic units are made of $[Fe^{III}(L)(CN)_3]^-$ complex units (L = Tp* for **2** and L = Tp for **1, 3** and **4**) acting as metalloligand (trough *cis*-coordinated cyanide) toward divalent Fe(II) ions. The coordination sphere of the Fe(II) ions is completed by two bik α-ketone-β–diimine ligands. The tetrametallic core of **2-4** exhibits a [2+2] type diamond-like geometry whose main angles are reported in table 3. The edge distances in **2** are slightly longer than those found in **3-4,** in agreement with the slightly longer Fe^{II}-ligand bond distances. This is likely due to the spin-state of the Fe(II) ions, which is expected to be purely low-spin at 200 K in **3-4** while it may contain some HS component in **2**, as shown by the magnetic measurements (see below). In all complexes, only moderate deviations from linearity are observed on the both side of the cyanide bridges as shown in Table 3.

Both Fe^{II} and Fe^{III} ions exhibit slightly distorted octahedral coordination sphere. The Fe^{II} ions are in a N_6 coordination sphere formed by four N atoms of the bik ligands and two N atoms of the cyanide located in *cis* position. The distortion of the coordination sphere can be evaluated by the octahedral distortion, Σ, defined as the sum of the deviation to 90° of the twelve angles around the metal atom. The octahedral distortion, Σ, measured at 200 K amount to 23.2°, 20.8° and 25.0° for **2-4**, respectively. These moderate values agree with the essentially LS states of the complexes **2-4** at 200 K. This is confirmed by the averaged Fe^{II}-N bond lengths (1.981, 1.952, 1953 Å for 1,2 and 3), which compare well with those found in similar Fe(II) low-spin complexes.[32,22,33]

The Fe^{III} coordination environment is formed by three imine moieties of the pyrazolyl groups and by the carbon atoms of three cyanides. The Fe-C_{CN} bond lengths (average of 1.920, 1.911, 1.912 Å for **2-4**) are in agreement with those previously reported for similar Fe^{III}-LS complexes.**[34,31]** The Fe^{III}-$N_{(Tp/Tp^*)}$ bond lengths are longer with mean values of 2.007, 1.977 and 1.975 Å for **2-4**. The octahedral distortion, Σ, increases upon coordination of the *N*- donor atom. They amount to 26.8°, 22.4°, 29.9° in **2-4**, respectively, to be compare with 17.9° and 18.5 ° in the $PPh_4[Fe^{III}(Tp^*)(CN)_3]$ CH_3CN and $PPh_4[Fe^{III}(Tp)(CN)_3]$ H_2O precursors.**[34,31]**

Finally, the square units are well separated from each other by the counter ions and the solvent lattice molecules. The shortest metal-metal distances are over 7.5 Å in **3-4** (see table 3). In the case of **2**, a water molecule establishes a hydrogen bridge between the non-bridging cyanide and an oxygen atom of the perchlorate counter ion, a feature commonly observed in these materials.

Table 3
Selected M-M distances,[a] bond lengths (Å) and angles (degrees) for **2-4**

Compounds	**2**	**3**	**4**
Fe^{III}-Fe^{II}-Fe^{III}	96.0°	94.5°	91.9°
Fe^{II}-Fe^{III}-Fe^{II}	84.0°	85.4°	88.1°
Fe^{III}-CN-Fe^{II}	5.038-5.038	4.954-4.964	4.947-4.971
Fe^{II}-$N_{(CN)}$	1.961(6)-1.976(7)	1.918(3)-1.924(3)	1.918(2)-1.928(2)
Fe^{II}-N_{bik}	1.973(7)-1.996(6)	1.966(3)-1.986(3)	1.964(2)-1.982(2)
Fe^{III}-$C_{(CN)}$ (average)	1.910(9)-1.930(8) (1.920)	1.901(4)-1.930(5) (1.911)	1.897(3)-1.934(3)(1.912)
Fe^{III}-$N_{(Tp/Tp^*)}$ (average)	1.985(6)-2.024(6) (2.007)	1.963(3)-1.986(3) (1.977)	1.972(2)-1.978(2) (1.975)
Fe^{II}-NC	174.3 -179.1°	171.4-175.9°	173.3-171.7°
Fe^{III}-CN	174.4(6)-177.6(7)	174.2-178.2°	177.4-178.4°
Σ Fe^{II}	23.2°	20.8°	25.0°
Σ Fe^{III}	26.8°	22.4°	29.9°
Fe···Fe (shortest) [b]	9.46	7.617	7.540

[a] all distances are given with an error of 6 to 9 on the last digit | [b] shortest intermolecular Fe-Fe distance

2.3. Magnetic Properties

The magnetic properties of **2-4** are shown in Figure 2 as the $\chi_M T$ products *versus* temperature (χ_M is the molar magnetic susceptibility per $\{Fe_2^{II}Fe^{III}_2\}$ formula unit). The $\chi_M T$ product of compound **2** shows a sigmoidal curve whose shape is typical of spin-crossover complexes. The $\chi_M T$ variation between low and high temperature, *ca.* 7.1 cm^3 mol^{-1} K, suggests an almost complete LS ⇔ HS conversion for the two Fe(II) ions, the expected value being 2 x 3.6 = 7.2 cm^3 mol^{-1} K, with S = 2, g ≈ 2.2. The transition temperature, $T_{1/2}$ = 227 K, is lower than those measured in our previous studies on SCO molecule based on the $\{Fe(bik)_2(N\text{-})_2\}$ unit. The $\chi_M T$ value measured at low temperature, *ca*. 1.4 cm^3 mol^{-1} K at 20 K, is ascribed to the presence of two Fe(III) LS ions ($\chi_M T$ ≈ 0.4 cm^3 mol^{-1} K at low temperature for each Fe^{III} ion) and possibly to a small amount of residual Fe(II) HS. Once heated at 400 K in the magnetometer, the spin transition is only slightly shifted toward lower temperature with $T_{1/2}$ ≈ 223 K (Fig. S3). This shift is likely due to the loss of crystallization solvent. In order to check this hypothesis, an *ex-situ* desolvated sample of **2** was studied. The sample was prepared by heating the crystals in a TGA under a N_2 gas flow up to 80°C. The resulting $\chi_M T$ curve matches well with that obtained for the sample desolvated inside the magnetometer (Fig. S3).

In contrast with **2**, the compounds **3** and **4** do not show a full sigmoidal curve in the 10-400 K temperature range, but only an increase of the $\chi_M T$ products above room temperature that suggests the occurrence of gradual spin-state conversion at high temperatures. The $\chi_M T$ values measured at 400 K, *ca*. 3.2 and 3.7 cm^3 mol^{-1} K for **3** and **4,** account for roughly 25% and 32% of high-spin Fe(II) and indicate transition temperatures that are above 400 K. The $\chi_M T$ products measured at low temperature for **3** and **4**, $\chi_M T$ ≈ 1.10 cm^3 mol^{-1} K at 20 K, are also coherent with the occurrence of two low-spin Fe^{II} and two low-spin Fe^{III} ions. Here the non-Curie behaviour observed between 20 and 350 K (a slight slope is observed on the $\chi_M T$ *vs T* curves) is due to the spin-orbit coupling occurring in the $^3T_{1g}$ ground state term of the low-spin $[Fe^{III}(Tp)(CN)_3]^-$ subunit.**[35]** It is also worth noticing that the spin transition in the compounds **3** and **4** are reversible. Upon cooling from 400 K, the $\chi_M T$ curves superimpose with that of the fresh compounds (see Fig. S4).

The magnetic curve of **2**, which shows a complete and gradual spin-crossover can be simulated using the Slichter-Drickammer mean-field model (equation 1) in order to extract the thermodynamic parameters associated to the spin-state equilibrium and to evaluate the cooperativity.**[36]**

$$\boldsymbol{ln}[(\mathbf{1}-\boldsymbol{n_{HS}})/(\mathbf{1}-\boldsymbol{f_{HS}})] = [\Delta \boldsymbol{H} + \boldsymbol{\Gamma}(\boldsymbol{f_{HS}} + \mathbf{1} - \mathbf{2}\boldsymbol{n_{HS}})]/\boldsymbol{RT} - \boldsymbol{\Delta S}/\boldsymbol{R} \quad (1)$$

In equation 1, ΔH and ΔS are the enthalpy and entropy variations associated to the spin transition, n_{HS} is the HS molar fraction and f_{HS} is the residual HS molar fraction at low temperature. Γ is the

parameter accounting for the cooperativity: the higher the gamma is, the steeper the spin-crossover is. In the present case, good fits were obtained on the $\chi_M T$ curve of the fresh sample and on the *in-situ* dehydrated sample for the following values ΔH = 24.0 (23.5) $kJ.mol^{-1}$ and ΔS = 105 (105) $J.K^{-1}.mol^{-1}$. The extracted Γ values, 1.6 (**1**) $kJ.mol^{-1}$, is moderate and lead to a $\mathbf{\Gamma/2RT_{1/2}}$ [illegible], as expected for gradual transition. This agrees well with the weak intermolecular interactions observed in the crystal structure. The enthalpy and entropy values are important, but in the range of the values that are usually obtained for SCO.**[37–39]** There are also close to those reported in our recent studies on the $[Fe(^{R}bik)_3]^{2+}$ family (ΔH = 29-33 $kJ.mol^{-1}.K$ and ΔS = 100-107 $J.mol^{-1}$).[24]

Finally, it is interesting to compare the magnetic behaviour of these three compounds with that of the previously published square complex **1**, $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}.[Fe^{III}(Tp)(CN)_3]_2\cdot 18H_2O\cdot 4CH_3OH$. While the $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}X_2$ square complexes (X = $[Fe(Tp)(CN)_3]^-$ (**1**), ClO_4^-, (**3**) and BF_4^- (**4**)) exhibit spin-transition temperature above room temperature (330 K for **1**, and over 400 K for **3** and **4**), the $T_{1/2}$ value of **2** is much lower, near 230 K. This suggests at first glance that $[Fe^{III}(Tp)(CN)_3]^-$ metalloligand is a better donor inducing a stronger ligand field on the $[Fe(bik)_2(N\text{-})_2]$ SCO units, as compared to the $[Fe^{III}(Tp^*)(CN)_3]^-$. However the situation is more complicated as the desolvation of **1** leads to a $T_{1/2}$ of *ca*. 240 K, a temperature close to the transition temperature of **2**. Actually intermolecular interactions can play a drastic role on the spin transition behaviour. In **1**, a complex hydrogen-bond network is observed, including a quite strong hydrogen bond between the non-bridging cyanide and a water molecule (N···O approx. 2.8 Å). The impact of the desolvation is much lower for **3** and **4**. This could be correlated to the nature of the intermolecular interactions. In **3** and **4**, only weak intermolecular interactions with CH_3CN (**3**) and MeOH (**4**) are observed, and the molecular squares remain well isolated from each other. The impact of the metalloligand could be compared by using solution studies, as we previously did for the $[Fe(bik)_3]^{2+}$ complexes, however we observed that the square complexes partially disassembled in CH_3CN or MeOH solution preventing such a study.

The Light-Induced Excited Spin-State Trapping (LIESST) was probed on fresh samples of **2-4** by measuring the magnetization *versus* time upon laser light irradiation at low temperature. The experimental conditions for **2-4** are similar to those previously used for the $\{Fe^{III}_2Fe^{II}_2\}$ (**1**) and $\{Mo^{V}_2Fe^{II}_2\}$ squares complexes (which include the $\{Fe(bik)_2(NC\text{-})_2\}$ subunits).**[22,23]** Among the three new compounds, compound **2** shows the most important photomagnetic effect. In fact compound **3** does not show any photomagnetic effect while only a weak increase of the magnetization is observed on **4** upon irradiation with a visible light (see Figure S6). As shown in

Figure 3, all the irradiations lead to a rapid and efficient increase of the $\chi_M T$ product of **2** for the six wavelengths that were probed. As for **1**, a plateau is reached after only *ca.* 20 min.**[22]** The irradiation at 635, 808 and 900 nm are the most efficient ones. These wavelengths correspond to MLCT band of the $[Fe^{II}(bik)_2(N\text{-})_2]$ units.**[22]** The $\chi_M T$ saturation value obtained in **2**, *ca.* 12 cm^3 mol^{-1} K, is notably higher than that found in **1**. It is also higher than that obtained at room temperature. This suggests the occurrence of ferromagnetic interactions in the photo-excited state. Actually both ferromagnetic and antiferromagnetic exchange pathways can be expected in a $\{Fe^{III}_{LS}\text{-}CN\text{-}Fe^{II}_{HS}\}$ bridge. However it has already been observed that the ferromagnetic interaction can dominate in similar low-dimensional systems.**[40,41]** The drop of the $\chi_M T$ product at low temperature (below 10 K) is likely due to the occurrence of antiferromagnetic intermolecular interaction or/and magnetic anisotropy effects. The stability of the metastable state of **2** was then probed by measuring the $\chi_M T$ *versus* T curve after irradiation (at 808 nm or 900nm), from 2 to 100 K (Figure 4). Upon heating the magnetization relaxes back to the diamagnetic state and becomes close to initial value at *ca.* 35 K. This behaviour is similar to that observed for **1**. Overall, the LIESST effect is much more efficient in **2** than in **3** and **4**. This is coherent with previous studies, which reported that the higher the thermal transition is (the higher the energy difference between the LS ground state and the HS state is), the smaller the stabilization of the metastable state is, and the lower the relaxation temperature is.**[42–44]**

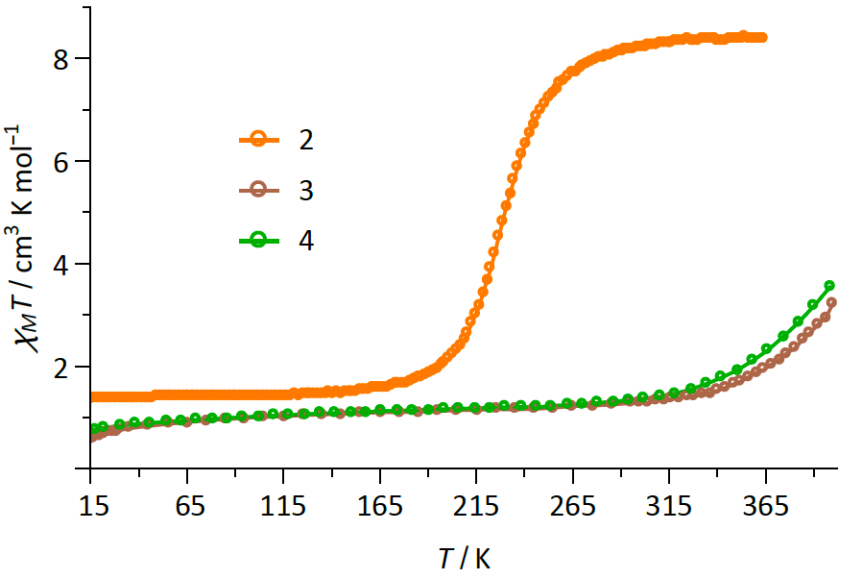

Fig. 2. Magnetic properties of **2-4** depicted as the $\chi_M T$ *versus* T curve.

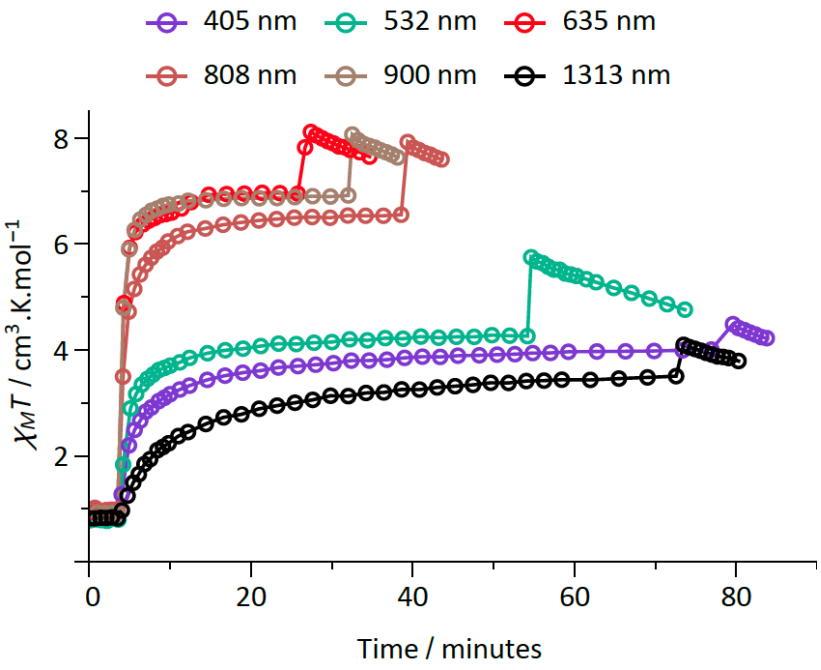

Fig. 3. Measurement of the $\chi_M T$ product *versus* time under laser light irradiation (ca. 5 mW/cm^2) at 20 K for the compound **2**. The first jump corresponds to the moment the light is switched on (photomagnetic effect) while the second jump corresponds to the moment the light is switched off and the temperature of the sample decreases.

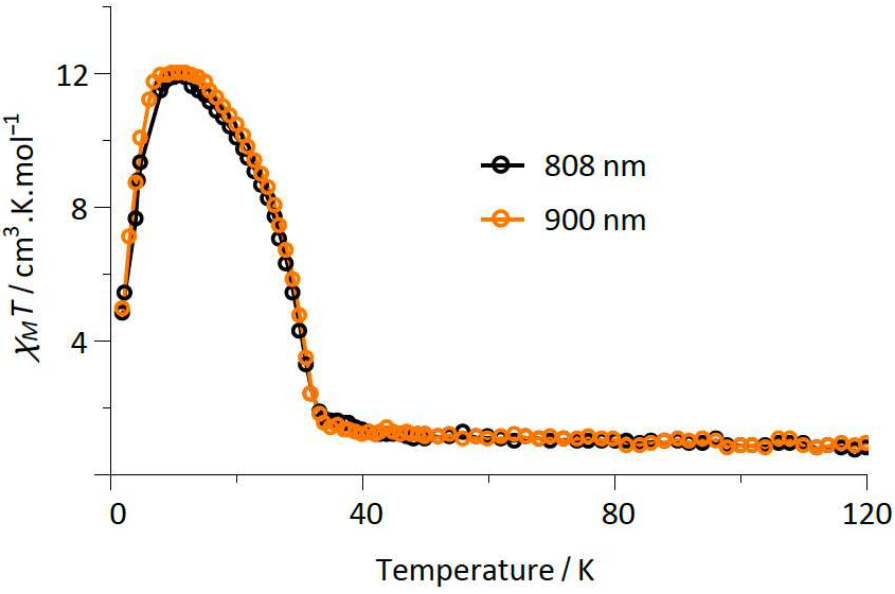

Fig. 4. Measurement of the $\chi_M T$ product after laser light irradiation at 20 K (H = 1T). The sample is measured from 2 to 100 K at 0.4 K /min.

3 Conclusions

Continuing our investigation on the $[Fe^{II}(bik)_2(N\text{-})_2]$ SCO units, we reported here three novel mixed valence cyanide-bridged $\{Fe^{III}_2Fe^{II}_2\}X_2$ square complexes (X = ClO_4, BF_4) which are reminiscent of our previously published $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}.[Fe^{III}(Tp)(CN)_3]_2$ square complex. These squares were obtained by reacting $[Fe^{II}(bik)_2(S)_2]^{2+}$ with the $[Fe^{III}(Tp)(CN)_3]^-$ or $[Fe^{III}(Tp^*)(CN)_3]^-$ *N*-donor metalloligands. The compounds **3** and **4** based on the $[Fe^{III}(Tp)(CN)_3]^-$ metalloligand exhibit spin-crossover at high temperature ($T_{1/2}$ > 400 K) and no (**3**) or moderate (**4**) photomagnetic effect. In contrast, the use of the $[Fe^{III}(Tp^*)(CN)_3]^-$ as *N*-donor metalloligand leads to a shift of the spin-crossover of the $[Fe^{II}(bik)_2(N\text{-})_2]$ units toward lower temperature, with $T_{1/2}$ = 233 K. The spin-state change in **2** is gradual and the thermodynamic parameter extracted from the magnetic measurements leads to weak value of cooperative factor ($\Gamma < \square\square\square\square_{1/2}$) in agreement with the absence of significant intermolecular interactions in **2**. The decrease of the $T_{1/2}$ temperature is associated to a stabilization of the high-spin metastable state and thus a more efficient LIESST effect. Remarkably, the irradiation with different laser lights covering the visible and NIR range leads to significant photomagnetic effect. However, the strongest increase of magnetization is observed when irradiating the sample at 635, 808 or 900 nm, in the region corresponding to LMCT band of the $[Fe^{II}(bik)_2(N\text{-})_2]$ units. The resulting photoinduced $\chi_M T$ values are high and indicate the occurrence of ferromagnetic interactions in the $\{Fe^{III}_{LS}\text{-}CN\text{-}Fe^{II}_{HS}\}$ metastable state.

Finally, this experimental study nicely complete our previous investigations on the photomagnetic properties of the $[Fe^{II}(bik)_2(N\text{-})_2]$ SCO units. When confronting all these studies, it appears that the nature of the *N*-donor metalloligand strongly influences the energy of the most efficient laser light inducing the LIESST effect. While the most efficient irradiation is located at 405 nm when using the $[Mo(CN)_8]^{3-}$ as N-donor ligands, it becomes 635 nm for the bik *N*-donor ligand, 750 nm for the $[Fe^{III}(Tp)(CN)_3]^-$ metalloligand, and 900 nm for the NCS^- N-donor ligand. The present study leads to the best results as a remarkable magnetization increase is observed with the same efficiency at 635, 808 and 900 nm when using the $[Fe(Tp^*)(CN)_3]^-$ N-donor metalloligand. The correlation with the absorption spectra of these compounds do not lead to clear conclusion as all of them are dominated by a intense LMCT bands of the $[Fe^{II}(bik)_2(N\text{-})_2]$ units. We believe that the rationalization of this behaviour would require advanced physical measurements to better investigate the excitation states involved in the LIESST effect.

4 Experimental Section

4.1. General

All chemicals used were of reagent grade quality. They were purchased from commercial sources and used as received. The bik ligand was prepared as reported Braussaud *et al.*[45]. The

$PPh_4[Fe^{III}(Tp)(CN)_3]\cdot H_2O$ complex was prepared in three steps by following our previously reported procedure [30].

4.2. Preparations

4.2.1. $PPh_4[Fe^{III}(Tp^*)(CN)_3]\ CH_3CN$

A degased solution of KTp* (319 mg, 1.0 mmol) in 5 mL methanol was added dropwise to a methanolic solution of $FeCl_2\cdot 4H_2O$ (199 mg, 1.0 mmol in 15 mL). The violet suspension was stirred for one hour, then added dropwise to a methanolic solution of KCN (214 mg, 3.3 mmol). The resulting red suspension was stirred at room temperature overnight and the methanol was evacuated to dryness. The red resulting solid was redissolved in acetonitrile and filtered. Crystals of $K[Fe^{III}(Tp^*)(CN)_3]$ were produced by slow evaporation of the acetonitrile solution. Crystals of $PPh_4[Fe^{III}(Tp^*)(CN)_3]\ H_2O$ suitable for X-ray diffraction analysis were produced in 1-2 weeks by slow evaporation of an acetonitrile solution of $K[Fe^{III}(Tp^*)(CN)_3]$ in which was added one equivalent of tetraphenylphosphonium chloride and a small amount of water. Yield: 32%. Elemental analysis (%): calculated for $C_{42}H_{42}BFeN_9P\cdot CH_3CN\cdot 0.5H_2O$: C64.41, H 5.65, N 17.07; found: C 64.71, H 5.43, N 16.79. ESI-MS *m/z* (%): 431.14 (100) $[Fe^{III}(Tp^*)(CN)_3]^-$, 339.13 (100), $[PPh_4]^+$.

4.2.2. $\{[Fe^{III}(Tp^*)(CN)_3]_2[Fe^{II}(bik)_2]\}_2(ClO_4)_2\cdot 2H_2O$ (**2**)

0.05 mmol of $Fe^{II}(ClO_4)_2\cdot 6H_2O$ (19 mg,) and bik ligand (19 mg, 0.1 mmol) were dissolved in 15 mL of a methanol/water (5/1) mixture. The resulting deep dark blue solution was added to a solution of $K[Fe(Tp^*)(CN)_3]$ (25 mg) in 15 mL of the same mixture of solvents. The purple solution was further stirred about 10 minutes before being filtered. Slow evaporation of the reaction mixture produced carmine red crystals suitable for X-ray diffraction analysis. Yield 40 %.

4.2.3. $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}(ClO_4)_2\ 2CH_3CN$ (**3**)

0.01 mmol (70.6 mg) of $PPh_4[Fe^{III}(Tp)(CN)_3]\cdot H_2O$ was dissolved in *ca.* 10 mL of acetonitrile and added dropwise to a solution containing $[Fe(bik)_2(S)_2](ClO_4)_2$ (S = solvent). The air stable $[Fe^{II}(bik)_2(S)_2]Cl_2$ complex was prepared *in situ* by reacting approximately 0.01 mmol (36.2 mg) of $Fe(ClO_4)_2\cdot 6H_2O$ and 0.02 mmol (38 mg) of the bik ligand in *ca.* 5 mL of acetonitrile. The resulting bluish-red solution was stirred at room temperature for 30 min and then filtered. Slow evaporation of the filtrate under ambient conditions afforded orange-red plate-like crystal of $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]_2\}(ClO_4)_2\ 2CH_3CN$ (**3**). Yield 60 %.

4.2.4. $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]\}_2(BF_4)_2\ 2CH_3OH$ (**4**)

0.01 mmol (70.6 mg) of $PPh_4[Fe^{III}(Tp)(CN)_3]\cdot H_2O$ was dissolved in ca. 10 mL of a methanolic/acetonitrile mixture and added dropwise to a solution containing $[Fe(bik)_2(S)_2]Cl_2$. The air stable $[Fe(bik)_2(S)_2]Cl_2$ complex was prepared *in situ* by reacting of 0.01 mmol (33.8 mg) of

$Fe(BF_4)_2 \cdot 6H_2O$ and 0.02 mmol (38 mg) of the bik ligand in *ca.* 5 mL of methanol. The resulting bluish-red solution was stirred at room temperature for 30 min and then filtered. Slow evaporation of the filtrate under ambient conditions afforded orange-red plate-like crystal of $\{[Fe^{III}(Tp)(CN)_3]_2[Fe^{II}(bik)_2]\}_2(BF_4)_2\ 2CH_3OH$ (**4**). Yield 60 %.

4.3. X-ray data collection and refinement

A single crystal of each compound was selected, mounted onto a cryoloop, and transferred in a cold nitrogen gas stream. Intensity data were collected with graphite-monochromated Mo-K□ radiation (λ = 0.71073 Å) either with a BRUKER Kappa-APEXII or a RIGAKU-OXFORD Saphire 3 Xcalibur diffractometer. Data collection, unit-cell parameters refinement, integration and data reduction were performed with Crysalis (Oxford) or APEX2 (Bruker) softwares. Multi-scan absorption corrections were applied. The structure were solved with either ShelxS86,**[46]** Superflip**[47]** or Sir97**[48]** programs and refined by full-matrix least-squares methods using SHELXL-14**[49]** or CRYSTALS**[50]**.

4.4. Physical techniques

Elemental analyses (C, H, N, Co, Mn, and Ni) were performed by the Service Central d'Analyse du CNRS (France). IR spectra were recorded on a Perkin–Elmer 882 spectrophotometer as KBr pellets. Variable-temperature (2.0–400 K) magnetic susceptibility measurements in the direct-current (dc) mode were carried out with a Quantum Design SQUID magnetometer using applied magnetic fields of 0.5 T (T > 30 K) and 0.025 T (T < 30 K). The magnetic data were corrected for the diamagnetism of the constituent atoms and the sample holder. The samples of **2-4** were measured as crystalline powders, removed from their mother solution directly before the measurements and introduced into the magnetometer at 200 K to avoid the loss of solvent crystallization molecules.

5 Supplementary material

X-ray crystallographic files in CIF format (excluding structure factors) for **2-4** and **5** are available from the Cambridge Crystallographic Data Centre, 12 Union road, Cambridge CB2 1EZ, UK (fax: +44-1223-336033; email: deposit@ccdc.cam.ac.uk or www: http://www.ccdc.cam.ac.uk) on request, on quoting the deposition number CCDC 1891890 (**2**), 1891892 (**3**), 1891891 (**4**) and 1891893 (**5**).

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